ICHÖ
General instructions
Cover sheet

Please return this cover sheet together with all the related question sheets.
International Chemistry Olympiad 2021 Japan
53rd IChO2021 Japan
25th July - 2nd August, 2021
https://www.icho2021.org
NZL English Version
General Instruction

- You are only allowed to use pen to write answers.
- Your calculator must be non-programmable.
- This examination has 9 problems.
- You can solve the problems in any order.
- You will have 5 hours to solve all problems.
- You can begin working only after the START command is given.
- All results must be written in the appropriate answer boxes with pen on the answer sheets. Use the back of the question sheets if you need paper for working. Remember that answers written outside the answer boxes will not be graded.
- You must write relevant calculations in the appropriate boxes when necessary. Full marks will only be given for correct answers when your work is shown.
- The invigilator will announce a 30-minute warning before the STOP command.
- You must stop working when the STOP command is given. Failure to stop writing will lead to the nullification of your examination.
- The official English version of this examination is available on request for clarification.
- You are not allowed to leave your working place without permission. If you need any assistance (e.g., broken calculator, need to visit a restroom), raise your hand and wait until an invigilator arrives.

GOOD LUCK!

Problems and Grading Information

<table>
<thead>
<tr>
<th>Title</th>
<th>Total Score</th>
<th>Percentage</th>
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<tbody>
<tr>
<td>1 Hydrogen at a Metal Surface</td>
<td>24</td>
<td>11</td>
</tr>
<tr>
<td>2 Isotope Time Capsule</td>
<td>35</td>
<td>11</td>
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<tr>
<td>3 Lambert–Beer Law?</td>
<td>22</td>
<td>8</td>
</tr>
<tr>
<td>4 The Redox Chemistry of Zinc</td>
<td>32</td>
<td>11</td>
</tr>
<tr>
<td>5 Mysterious Silicon</td>
<td>60</td>
<td>12</td>
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<tr>
<td>6 The Solid-State Chemistry of Transition Metals</td>
<td>45</td>
<td>13</td>
</tr>
<tr>
<td>7 Playing with Non-benzenoid Aromaticity</td>
<td>36</td>
<td>13</td>
</tr>
<tr>
<td>8 Dynamic Organic Molecules and Their Chirality</td>
<td>26</td>
<td>11</td>
</tr>
<tr>
<td>9 Likes and Dislikes of Capsules</td>
<td>23</td>
<td>10</td>
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<tr>
<td>Total</td>
<td>100</td>
<td></td>
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</table>
## Physical Constants and Equations

### Constants

<table>
<thead>
<tr>
<th>Constant</th>
<th>Value</th>
</tr>
</thead>
<tbody>
<tr>
<td>Speed of light in vacuum</td>
<td>$c = 2.99792458 \times 10^8$ m s$^{-1}$</td>
</tr>
<tr>
<td>Planck constant</td>
<td>$h = 6.62607015 \times 10^{-34}$ J s</td>
</tr>
<tr>
<td>Elementary charge</td>
<td>$e = 1.602176634 \times 10^{-19}$ C</td>
</tr>
<tr>
<td>Electron mass</td>
<td>$m_e = 9.10938370 \times 10^{-31}$ kg</td>
</tr>
<tr>
<td>Electric constant (permittivity of vacuum)</td>
<td>$\varepsilon_0 = 8.85418781 \times 10^{-12}$ F m$^{-1}$</td>
</tr>
<tr>
<td>Avogadro constant</td>
<td>$N_A = 6.02214076 \times 10^{23}$ mol$^{-1}$</td>
</tr>
<tr>
<td>Boltzmann constant</td>
<td>$k_B = 1.380649 \times 10^{-23}$ J K$^{-1}$</td>
</tr>
<tr>
<td>Faraday constant</td>
<td>$F = N_A \times e = 9.6485321233100184 \times 10^4$ C mol$^{-1}$</td>
</tr>
<tr>
<td>Gas constant</td>
<td>$R = N_A \times k_B = 8.31446261815324$ J K$^{-1}$ mol$^{-1}$</td>
</tr>
<tr>
<td>Unified atomic mass unit</td>
<td>$u = 1$ Da $= 1.66053907 \times 10^{-27}$ kg</td>
</tr>
<tr>
<td>Standard pressure</td>
<td>$p = 1$ bar $= 10^5$ Pa</td>
</tr>
<tr>
<td>Atmospheric pressure</td>
<td>$p_{atm} = 1.01325 \times 10^5$ Pa</td>
</tr>
<tr>
<td>Zero degree Celsius</td>
<td>$0^\circ$ C $= 273.15$ K</td>
</tr>
<tr>
<td>Ångstrom</td>
<td>$1$ Å $= 10^{-10}$ m</td>
</tr>
<tr>
<td>Picometer</td>
<td>$1$ pm $= 10^{-12}$ m</td>
</tr>
<tr>
<td>Electronvolt</td>
<td>$1$ eV $= 1.602176634 \times 10^{-19}$ J</td>
</tr>
<tr>
<td>Part-per-million</td>
<td>$1$ ppm $= 10^{-6}$</td>
</tr>
<tr>
<td>Part-per-billion</td>
<td>$1$ ppb $= 10^{-9}$</td>
</tr>
<tr>
<td>Part-per-trillion</td>
<td>$1$ ppt $= 10^{-12}$</td>
</tr>
<tr>
<td>pi</td>
<td>$\pi = 3.141592653589793$</td>
</tr>
<tr>
<td>The base of the natural logarithm (Euler’s number)</td>
<td>$e = 2.718281828459045$</td>
</tr>
</tbody>
</table>
# Equations

- **The ideal gas law**
  \[ PV = nRT \]
  where \( P \) is the pressure, \( V \) is the volume, \( n \) is the amount of substance, \( T \) is the absolute temperature of ideal gas.

- **Coulomb's law**
  \[ F = k_e \frac{q_1 q_2}{r^2} \]
  where \( F \) is the electrostatic force, \( k_e \approx 9.0 \times 10^9 \text{ N m}^2\text{C}^{-2} \) is Coulomb's constant, \( q_1 \) and \( q_2 \) are the magnitudes of the charges, and \( r \) is the distance between the charges.

- **The first law of thermodynamics**
  \[ \Delta U = q + w \]
  where \( \Delta U \) is the change in the internal energy, \( q \) is the heat supplied to the system, \( w \) is the work done on the system.

- **Enthalpy**
  \[ H = U + PV \]

- **Entropy based on Boltzmann's principle**
  \[ S = k_B \ln W \]
  where \( W \) is the number of microstates.

- **The change of entropy**
  \[ \Delta S = \frac{q_{rev}}{T} \]
  where \( q_{rev} \) is the heat for the reversible process.

- **Gibbs free energy**
  \[ G = H - TS \]
  \[ \Delta_r G^\circ = -RT \ln K = -zFE^\circ \]
  where \( K \) is the equilibrium constant, \( z \) is the number of electrons, \( E^\circ \) is the standard electrode potential.

- **Reaction quotient**
  \[ Q = \frac{[C]^c[D]^d}{[A]^a[B]^b} \]
  where \([A]\) is the concentration of A.
<table>
<thead>
<tr>
<th>Equation</th>
<th>Description</th>
</tr>
</thead>
<tbody>
<tr>
<td>Heat change ( \Delta q )</td>
<td>( \Delta q = n c_m \Delta T ), where ( c_m ) is the temperature-independent molar heat capacity.</td>
</tr>
<tr>
<td>Nernst equation for redox reaction</td>
<td>( E = E^\circ + \frac{RT}{ZF} \ln \frac{C_{\text{ox}}}{C_{\text{red}}} ), where ( C_{\text{ox}} ) is the concentration of oxidized substance, ( C_{\text{red}} ) is the concentration of reduced substance.</td>
</tr>
<tr>
<td>Arrhenius equation</td>
<td>( k = A \exp \left( -\frac{E_a}{RT} \right) ), where ( k ) is the rate constant, ( A ) is the pre-exponential factor, ( E_a ) is the activation energy.</td>
</tr>
<tr>
<td>Lambert–Beer equation</td>
<td>( A = \varepsilon l c ), where ( A ) is the absorbance, ( \varepsilon ) is the molar absorption coefficient, ( l ) is the optical path length, ( c ) is the concentration of the solution.</td>
</tr>
<tr>
<td>Henderson–Hasselbalch equation</td>
<td>For an equilibrium ( \text{HA} \rightleftharpoons \text{H}^+ + \text{A}^- ), where equilibrium constant is ( K_a ), ( \text{pH} = pK_a + \log \left( \frac{[\text{A}^-]}{[\text{HA}]} \right) ).</td>
</tr>
<tr>
<td>Energy of a photon</td>
<td>( E = h\nu = \frac{h}{\lambda} ), where ( \nu ) is the frequency, ( \lambda ) is the wavelength of the light.</td>
</tr>
<tr>
<td>The sum of a geometric series</td>
<td>When ( x \neq 1 ), ( 1 + x + x^2 + \cdots + x^n = \sum_{i=0}^{n} x^i = \frac{1 - x^{n+1}}{1 - x} ).</td>
</tr>
<tr>
<td>Approximation equation that can be used to solve problems</td>
<td>When ( x \ll 1 ), ( \frac{1}{1 - x} \approx 1 + x ).</td>
</tr>
<tr>
<td>1</td>
<td>2</td>
</tr>
<tr>
<td>----</td>
<td>----</td>
</tr>
<tr>
<td>H</td>
<td>He</td>
</tr>
<tr>
<td>K</td>
<td>Ca</td>
</tr>
<tr>
<td>39.098</td>
<td>40.078</td>
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<tr>
<td>Rb</td>
<td>Sr</td>
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<td>85.468</td>
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</tr>
<tr>
<td>Cs</td>
<td>Ba</td>
</tr>
<tr>
<td>132.905</td>
<td>137.327</td>
</tr>
<tr>
<td>Fr</td>
<td>Ra</td>
</tr>
<tr>
<td>223</td>
<td>226</td>
</tr>
</tbody>
</table>

Key:
- **Symbol**: Atomic symbol
- **Name**: Element name
- **Atomic Number**: Atomic number
- **Atomic Weight**: Atomic weight (in parentheses for the radioactive element)

**Notation**:
- **La-Lu**: Lanthanides
- **Ac-Lr**: Actinides

**Note**: The image shows a periodic table with atomic numbers and symbols for elements 1 to 18, with the additional notations for the lanthanides and actinides. The table includes elements from hydrogen (H) to oganesson (Og).
$\Delta \delta$ for one alkyl group-substitution: ca. +0.4 ppm
Please return this cover sheet together with all the related question sheets.
Hydrogen is expected to be a future energy source that does not depend on fossil fuels. Here, we will consider the hydrogen-storage process in a metal, which is related to hydrogen-transport and -storage technology.

**Part A**

As hydrogen is absorbed into the bulk of a metal via its surface, let us first consider the adsorption process of hydrogen at the metal surface, \( H_2(g) \rightarrow 2H(ad) \), where the gaseous and adsorbed states of hydrogen are represented as (g) and (ad), respectively.

Hydrogen molecules (\( H_2 \)) that reach the metal surface (M) dissociate at the surface and are adsorbed as H atoms (Fig. 1). Here, the potential energy of \( H_2 \) is represented by two variables: the distance between the hydrogen atoms, \( d \), and the height relative to the surface metal atom, \( z \). It is assumed that the axis along the two H atoms is parallel to the surface and that the centre of mass is always on the vertical dotted line in Fig. 1.

Fig. 2 shows the potential energy contour plot for the dissociation at the surface. The numerical values on the graph represent the potential energy in units of \( \text{kJ per mole of } H_2 \). The solid line spacing is 20 \( \text{kJ mol}^{-1} \), the dashed line spacing is 100 \( \text{kJ mol}^{-1} \), and the spacing between solid and dashed lines is 80 \( \text{kJ mol}^{-1} \). The zero-point vibration energy is ignored.
Fig. 1 Definition of variables. Drawing is not in scale.

Fig. 2 Potential energy contour plot for the dissociation at the surface.
**A.1** For each of the following items (i)–(iii), select the closest value from A–G.  
(i) The interatomic distance for a gaseous H$_2$ molecule 
(ii) The interatomic distance between metal atoms ($d_M$ in Fig. 1) 
(iii) The distance of adsorbed H atoms from the surface ($h_{ad}$ in Fig. 1) 

<table>
<thead>
<tr>
<th></th>
<th>A. 0.03 nm</th>
<th>B. 0.07 nm</th>
<th>C. 0.11 nm</th>
<th>D. 0.15 nm</th>
<th>E. 0.19 nm</th>
<th>F. 0.23 nm</th>
<th>G. 0.27 nm</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
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<td></td>
</tr>
</tbody>
</table>

**A.2** For each of the following items (i)–(ii), select the closest value from A–H.  
(i) the energy required for the dissociation of gaseous H$_2$ to gaseous H 
   [H$_2$(g) $\rightarrow$ 2H(g)] 
   (ii) the energy released during the adsorption of a gaseous H$_2$ [H$_2$(g) $\rightarrow$ 2H(ad)] 

<table>
<thead>
<tr>
<th></th>
<th>A. 20 kJ mol$^{-1}$</th>
<th>B. 40 kJ mol$^{-1}$</th>
<th>C. 60 kJ mol$^{-1}$</th>
<th>D. 100 kJ mol$^{-1}$</th>
<th>E. 150 kJ mol$^{-1}$</th>
<th>F. 200 kJ mol$^{-1}$</th>
<th>G. 300 kJ mol$^{-1}$</th>
<th>H. 400 kJ mol$^{-1}$</th>
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<tr>
<td></td>
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<td></td>
<td></td>
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</tr>
</tbody>
</table>
Part B

The adsorbed hydrogen atoms are then either absorbed into the bulk, or recombine and desorb back into the gas phase, as shown in the reactions (1a) and (1b). H(ab) represents a hydrogen atom absorbed in the bulk.

\[ \text{H}_2(\text{g}) \xrightarrow{k_1/k_2} 2\text{H(ad)} \] (1a)

\[ \text{H(ad)} \xrightarrow{k_3} \text{H(ab)} \] (1b)

The reaction rates per surface site for adsorption, desorption, and absorption are \( r_1 [s^{-1}] \), \( r_2 [s^{-1}] \) and \( r_3 [s^{-1}] \), respectively. They are expressed as:

\[ r_1 = k_1 P_{\text{H}_2} (1 - \theta)^2 \] (2)

\[ r_2 = k_2 \theta^2 \] (3)

\[ r_3 = k_3 \theta \] (4)

where \( k_1 [s^{-1} \text{Pa}^{-1}] \), \( k_2 [s^{-1}] \) and \( k_3 [s^{-1}] \) are the reaction rate constants and \( P_{\text{H}_2} \) is the pressure of H\(_2\).

Among the sites available on the surface, \( \theta (0 \leq \theta \leq 1) \) is the fraction occupied by H atoms.

It is assumed that adsorption and desorption are fast compared to absorption (\( r_1, r_2 \gg r_3 \)) and that \( \theta \) remains constant.

B.1 \( r_3 \) can be expressed as:

\[ r_3 = \frac{k_3}{1 + \sqrt{\frac{1}{P_{\text{H}_2} C}}} \] (5)

Express \( C \) using \( k_1 \) and \( k_2 \).
A metal sample with a surface area of $S = 1.0 \times 10^{-3} \text{ m}^2$ was placed in a container ($1 \text{ L} = 1.0 \times 10^{-3} \text{ m}^3$) with $\text{H}_2 (P_{\text{H}_2} = 1.0 \times 10^2 \text{ Pa})$.

The number of hydrogen-atom adsorption sites per unit area on the surface was $N = 1.3 \times 10^{18} \text{ m}^{-2}$.

The surface temperature was kept at $T = 400 \text{ K}$.

As the reaction (1) proceeded, $P_{\text{H}_2}$ decreased at a constant rate of $v = 4.0 \times 10^{-4} \text{ Pa s}^{-1}$.

Assume that $\text{H}_2$ is an ideal gas and that the volume of the metal sample is negligible.

**B.2** Calculate the amount of H atoms in moles absorbed per unit area of the surface per unit time, $A \left[ \text{mol s}^{-1} \text{ m}^{-2} \right]$.

**B.3** At $T = 400 \text{ K}$, $C$ equals $1.0 \times 10^2 \text{ Pa}^{-1}$. Calculate the value of $k_3$ at 400 K. If you did not obtain the answer to B.2, use $A = 3.6 \times 10^{-7} \text{ mol s}^{-1} \text{ m}^{-2}$.

**B.4** At a different $T$, $C = 2.5 \times 10^3 \text{ Pa}^{-1}$ and $k_3 = 4.8 \times 10^{-2} \text{ s}^{-1}$ are given. For $r_3$ as a function of $P_{\text{H}_2}$ at this temperature, select the correct plot from (a)-(h).
Hydrogen at a Metal Surface

Part A

A.1 (6 pt)

(i)  (ii)  (iii)

A.2 (4 pt)

(i)  (ii)
Part B

B.1 (5 pt)

\[ C = \]
B.2 (3 pt)

\[ A = \text{mol s}^{-1} \text{m}^{-2} \]

B.3 (3 pt)

\[ k_3 = \text{s}^{-1} \]

B.4 (3 pt)

____________
Please return this cover sheet together with all the related question sheets.
Isotope Time Capsule

<table>
<thead>
<tr>
<th>Question</th>
<th>A.1</th>
<th>A.2</th>
<th>A.3</th>
<th>A.4</th>
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<td>Score</td>
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<td></td>
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</table>

Molecular entities that differ only in isotopic composition, such as CH₄ and CH₃D, are called isotopologues. Isotopologues are considered to have the same chemical characteristics. In nature, however, there exists a slight difference.

Assume that all of the substances shown in this Question are in a gas phase.

Let us consider the following equilibrium:

\[ ^{12}\text{C}^{16}\text{O}_2 + ^{12}\text{C}^{18}\text{O}_2 \rightleftharpoons 2^{12}\text{C}^{16}\text{O}^{18}\text{O} \]

The entropy, \( S \), increases with increasing the number of possible microscopic states of a system, \( W \):

\[ S = k_B \ln W \]  

\( W = 1 \) for \(^{12}\text{C}^{16}\text{O}_2\) and \(^{12}\text{C}^{18}\text{O}_2\).

In contrast, \( W = 2 \) for a \(^{12}\text{C}^{16}\text{O}^{18}\text{O}\) molecule because the oxygen atoms are distinguishable in this molecule. As the right-hand side of the equilibrium shown in eq. 1 has two \(^{12}\text{C}^{16}\text{O}^{18}\text{O}\) molecules, \( W = 2^2 = 4 \).
A.1 The enthalpy change, $\Delta H$, of eq. 3 is positive regardless of the temperature.

\[ H_2 + DI \rightleftharpoons HD + HI \]  

(3)

**Calculate** the equilibrium constants, $K$, for eq. 3 at very low (think of $T \to 0$) and very high (think of $T \to +\infty$) temperatures. Assume that the reaction remains unchanged at these temperatures and that $\Delta H$ converges to a constant value for high temperatures.

The $\Delta H$ of the following process can be explained by molecular vibrations.

\[ 2HD \rightleftharpoons H_2 + D_2 \quad K = \frac{[H_2][D_2]}{[HD]^2} \]  

(4)

At $T = 0$ K, the vibrational energy of a diatomic molecule whose vibration frequency is $\nu$ [s$^{-1}$] is expressed as:

\[ E = \frac{1}{2} h \nu \]  

(5)

\[ \nu = \frac{1}{2\pi} \sqrt{\frac{k}{\mu}} \]  

(6)

Wherein $k$ is the force constant and $\mu$ the reduced mass, which is expressed in terms of the mass of the two atoms in the diatomic molecule, $m_1$ and $m_2$, according to:

\[ \mu = \frac{m_1 m_2}{m_1 + m_2} \]  

(7)

A.2 The vibration of $H_2$ is at 4161.0 cm$^{-1}$ when reported as a wavenumber.  

**Calculate** the $\Delta H$ of the following equation at $T = 0$ K in units of J mol$^{-1}$.

\[ 2HD \rightarrow H_2 + D_2 \]  

(8)

Assume that:

- only the vibrational energy contributes to the $\Delta H$.
- the $k$ values for $H_2$, HD, and $D_2$ are identical.
- the mass of H to be 1 Da and the mass of D to be 2 Da.
The molar ratio of $\text{H}_2$, HD, and $\text{D}_2$ depends on the temperature in a system in equilibrium. Here, $\Delta_{\text{D}_2}$ is defined as the change of the molar ratio of $\text{D}_2$.

$$\Delta_{\text{D}_2} = \frac{R_{\text{D}_2}}{R_{\text{D}_2}^*} - 1$$  \hspace{1cm} (9)

Here, $R_{\text{D}_2}$ refers to $\frac{[\text{D}_2]}{[\text{H}_2]}$ in the sample and $R_{\text{D}_2}^*$ to $\frac{[\text{D}_2]}{[\text{H}_2]}$ at $T \rightarrow +\infty$. It should be noted here that the distribution of isotopes becomes random at $T \rightarrow +\infty$.

**A.3 Calculate** $\Delta_{\text{D}_2}$ with natural D abundance when the isotopic exchange is in equilibrium at the temperature where $K$ in eq. 4 is 0.300. Assume that the natural abundance ratios of D and H are $1.5576 \times 10^{-4}$ and $1 - 1.5576 \times 10^{-4}$, respectively.
In general, the molar ratio of a doubly substituted isotopologue, which contains two heavy isotope atoms in one molecule, increases with decreasing temperature.

Let us consider the molar ratio of CO$_2$ molecules with molecular weights of 44 and 47, which are described as CO$_2$[44] and CO$_2$[47] below. The quantity $\Delta_{47}$ is defined as:

$$\Delta_{47} = \frac{R_{47}}{R_{47}^*} - 1$$

(R$_{47}$ refers to $\left[\frac{\text{CO}_2[47]}{\text{CO}_2[44]}\right]$ in the sample and $R_{47}^*$ to $\left[\frac{\text{CO}_2[47]}{\text{CO}_2[44]}\right]$ at $T \to +\infty$. The natural abundances of carbon and oxygen atoms are shown below; ignore isotopes that are not shown here.

<table>
<thead>
<tr>
<th></th>
<th>$^{12}\text{C}$</th>
<th>$^{13}\text{C}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>natural abundance</td>
<td>0.988888</td>
<td>0.011112</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th></th>
<th>$^{16}\text{O}$</th>
<th>$^{17}\text{O}$</th>
<th>$^{18}\text{O}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>natural abundance</td>
<td>0.997621</td>
<td>0.0003790</td>
<td>0.0020000</td>
</tr>
</tbody>
</table>

The temperature dependence of $\Delta_{47}$ is determined as follows, where $T$ is given as the absolute temperature in units of K:

$$\Delta_{47} = \frac{36.2}{T^2} + 2.920 \times 10^{-4}$$

The $R_{47}$ of fossil plankton obtained from the Antarctic seabed was $4.50865 \times 10^{-5}$. Estimate the temperature using this $R_{47}$. This temperature is interpreted as the air temperature during the era in which the plankton lived. Consider only the most common isotopologue of CO$_2$[47] for the calculation.
Isotope Time Capsule

A.1 (8 pt)

\[ T \to 0 : K = \quad , \quad T \to +\infty : K = \]
Δ\(H\) = \[\text{J mol}^{-1}\]
Δ_{D_2} = 

A.3 (10 pt)
\[ T = \text{K} \]
Please return this cover sheet together with all the related question sheets.
Lambert–Beer Law?

<table>
<thead>
<tr>
<th>Question</th>
<th>A.1</th>
<th>B.1</th>
<th>B.2</th>
<th>Total</th>
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</thead>
<tbody>
<tr>
<td>Points</td>
<td>10</td>
<td>6</td>
<td>6</td>
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</tr>
<tr>
<td>Score</td>
<td></td>
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<td></td>
<td></td>
</tr>
</tbody>
</table>

In this problem, ignore the absorption of the cell and the solvent. The temperatures of all solutions and gases are kept constant at 25 °C.

**Part A**

An aqueous solution X was prepared using HA and NaA.

The concentrations \([A^-], [HA], \text{ and } [H^+]\) in solution X are \(1.00 \times 10^{-2} \text{ mol L}^{-1}, 1.00 \times 10^{-3} \text{ mol L}^{-1}, \text{ and } 1.00 \times 10^{-4} \text{ mol L}^{-1}\), respectively, which are correlated via the following acid-base equilibrium:

\[
HA \rightleftharpoons A^- + H^+ \quad K = \frac{[A^-][H^+]}{[HA]} \quad (1)
\]

The optical path length is \(l\) in Part A. Ignore the density change upon dilution. Assume that no chemical reactions other than eq 1 occur.

**A.1**  
The absorbance of X was \(A_1\) at a wavelength of \(\lambda_1\). Then, solution X was diluted to twice its initial volume using hydrochloric acid with pH = 2.500. After the dilution, the absorbance was still \(A_1\) at \(\lambda_1\). **Determine** the ratio \(\varepsilon_{HA}/\varepsilon_{A^-}\), where \(\varepsilon_{HA}\) and \(\varepsilon_{A^-}\) represent the absorption coefficients of HA and of A\(^{-}\), respectively, at \(\lambda_1\).
Part B

Let us consider the following equilibrium in the gas phase.

\[ D \rightleftharpoons 2M \quad (2) \]

Pure gas D is filled into a cuboid container that has a transparent movable wall with a cross-section of \( S \) (see the figure below) at a pressure \( P \), and equilibrium is established while the total pressure is kept at \( P \).

The absorbance of the gas is \( A = \varepsilon (n/V)l \), where \( \varepsilon \), \( n \), \( V \), and \( l \) are the absorption coefficient, amount of the gas in moles, volume of the gas, and optical path length, respectively.

Assume that all components of the gas mixture behave as ideal gases.

Use the following definitions if necessary.

<table>
<thead>
<tr>
<th></th>
<th>Initial state</th>
<th>After equilibrium</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>D</td>
<td>M</td>
</tr>
<tr>
<td>Partial pressure</td>
<td>( P )</td>
<td>0</td>
</tr>
<tr>
<td>Amount in moles</td>
<td>( n_0 )</td>
<td>0</td>
</tr>
<tr>
<td>Volume</td>
<td>( V_0 )</td>
<td>( V )</td>
</tr>
</tbody>
</table>

B.1 The absorbance of the gas at \( \lambda_{B1} \) measured from direction \( x \) (\( l = l_x \)) was \( A_{B1} \) both at the initial state and after the equilibrium. **Determine** the ratio \( \varepsilon_D/\varepsilon_M \) at \( \lambda_{B1} \), where \( \varepsilon_D \) and \( \varepsilon_M \) represent the absorption coefficients of D and of M, respectively.

B.2 The absorbance of the gas at \( \lambda_{B2} \) measured from direction \( y \) was \( A_{B2} \) both at the initial state (\( l = l_{y0} \)) and after the equilibrium (\( l = l_y \)). **Determine** the ratio \( \varepsilon_D/\varepsilon_M \) at \( \lambda_{B2} \).
Lambert–Beer Law?

Part A

A.1 (10 pt)

(Continued on the next page)
A.1 (cont.)

\[ \frac{\varepsilon_{HA}}{\varepsilon_{A^-}} = \]
Part B

B.1 (6 pt)

\[ \varepsilon_D/\varepsilon_M = \]
\[ \frac{\varepsilon_D}{\varepsilon_M} = \]
Please return this cover sheet together with all the related question sheets.
The Redox Chemistry of Zinc

<table>
<thead>
<tr>
<th>Question</th>
<th>A.1</th>
<th>A.2</th>
<th>B.1</th>
<th>B.2</th>
<th>B.3</th>
<th>B.4</th>
<th>Total</th>
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<tbody>
<tr>
<td>Points</td>
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<td>5</td>
<td>4</td>
<td>3</td>
<td>5</td>
<td>9</td>
<td>32</td>
</tr>
</tbody>
</table>

Zinc has long been used as alloys for brass and steel materials. The zinc contained in industrial wastewater is separated by precipitation to detoxify the water, and the precipitate is reduced to re-use it as metallic zinc.

**Part A**

The dissolution equilibrium of zinc hydroxide Zn(OH)$_2$(s) at 25°C and the relevant equilibrium constants are given in eq. 1–4.

\[
\text{Zn(OH)}_2(\text{s}) \rightleftharpoons \text{Zn}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq}) \quad K_{sp} = 1.74 \times 10^{-17} \tag{1}
\]

\[
\text{Zn(OH)}_2(\text{s}) \rightleftharpoons \text{Zn(OH)}_2(\text{aq}) \quad K_1 = 2.62 \times 10^{-6} \tag{2}
\]

\[
\text{Zn(OH)}_2(\text{s}) + 2\text{OH}^-(\text{aq}) \rightleftharpoons \text{Zn(OH)}^2_4(\text{aq}) \quad K_2 = 6.47 \times 10^{-2} \tag{3}
\]

\[
\text{H}_2\text{O}(l) \rightleftharpoons \text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \quad K_w = 1.00 \times 10^{-14} \tag{4}
\]
The solubility, \( S \), of zinc (total concentration of zinc in a saturated aqueous solution) is given in eq. 5.

\[
S = [\text{Zn}^{2+}(aq)] + [\text{Zn(OH)}_2(aq)] + [\text{Zn(OH)}_4^{2-}(aq)]
\]  

**A.1** When the equilibria in eq. 1–4 are established, **calculate** the pH range in which \([\text{Zn(OH)}_2(aq)]\) is the greatest among \([\text{Zn}^{2+}(aq)]\), \([\text{Zn(OH)}_2(aq)]\) and \([\text{Zn(OH)}_4^{2-}(aq)]\).

**A.2** A saturated aqueous solution of \(\text{Zn(OH)}_2(s)\) with pH = 7.00 was prepared and filtered. NaOH was added to this filtrate to increase its pH to 12.00. **Calculate** the molar percentage of zinc that precipitates when increasing the pH from 7.00 to 12.00. Ignore any volume or temperature changes.

**Part B**

Next, the recovered zinc hydroxide is heated to obtain zinc oxide according to the reaction below:

\[
\text{Zn(OH)}_2(s) \rightarrow \text{ZnO(s)} + \text{H}_2\text{O(l)}
\]  

The zinc oxide is then reduced to metallic zinc by reaction with hydrogen:

\[
\text{ZnO(s)} + \text{H}_2(g) \rightarrow \text{Zn(s)} + \text{H}_2\text{O(g)}
\]  

**B.1** In order for reaction (7) to proceed at a hydrogen pressure kept at 1 bar, it is necessary to reduce the partial pressure of the generated water vapor. **Calculate** the upper limit for the partial pressure of water vapor to allow reaction (7) to proceed at 300 °C.

Here, the Gibbs formation energies of zinc oxide and water vapor at 300 °C and 1 bar for all gaseous species are

\[
\Delta G_{\text{ZnO}}^{\circ}(300\degree C) = -2.90 \times 10^2 \text{kJ mol}^{-1}
\]

\[
\Delta G_{\text{H}_2\text{O}}^{\circ}(300\degree C) = -2.20 \times 10^2 \text{kJ mol}^{-1}
\]

Metallic zinc is used as a negative electrode (anode) material for metal-air batteries. The electrode consists of Zn and ZnO. It uses the following redox reaction to generate electricity with the electromotive force (e.m.f.) at 25 °C and pressure of 1 bar, \( E' \).

\[
\text{Zn(s)} + \frac{1}{2}\text{O}_2(g) \rightarrow \text{ZnO(s)} \quad \quad E' = 1.65 \text{V}
\]  

**B.2** A zinc-air battery was discharged at 20 mA for 24 hours. **Calculate** the change in mass of the negative electrode (anode) of the battery.
B.3 Consider the change of e.m.f. of a zinc–air battery depending on the environment.  

Calculate the e.m.f. of a zinc–air battery at the summit of Mt. Fuji, where the temperature and altitude are $-38 \, ^\circ C$ (February) and $3776$ m, respectively. The pressure of the atmosphere at a given altitude is represented by

$$P \text{[bar]} = 1.013 \times \left(1 - \frac{0.0065h}{T + 0.0065h + 273.15}\right)^{5.257}$$

(9)

at altitude $h$ [m] and temperature $T$ [$^\circ C$].

The molar ratio of oxygen in the atmosphere is 21%.

The Gibbs energy change of reaction (8) is $\Delta G_{\text{Znno}}(-38 \, ^\circ C) = -3.26 \times 10^2$ kJ mol$^{-1}$ at $-38 \, ^\circ C$ and 1 bar.

B.4 Calculate the Gibbs energy change for reaction (6) at 25 $^\circ C$.

Note that the standard reduction potentials, $E^\circ (\text{Zn}^{2+}/\text{Zn})$ and $E^\circ (\text{O}_2/\text{H}_2\text{O})$ at 25 $^\circ C$ and 1 bar are given in (10) and (11), respectively.

$$\text{Zn}^{2+} + 2e^- \rightarrow \text{Zn} \quad E^\circ (\text{Zn}^{2+}/\text{Zn}) = -0.77 \text{V} \quad (10)$$

$$\text{O}_2 + 4\text{H}^+ + 4e^- \rightarrow 2\text{H}_2\text{O} \quad E^\circ (\text{O}_2/\text{H}_2\text{O}) = 1.23 \text{V} \quad (11)$$
The Redox Chemistry of Zinc

Part A

A.1 (6 pt)
Part B

B.1 (4 pt)

\[ p_{H_2O} = \text{bar} \]

B.2 (3 pt)

\[ \text{g} \]
$$\Delta G^\circ = \text{ J mol}^{-1}$$
Please return this cover sheet together with all the related question sheets.
Mysterious Silicon

Although silicon is also a group 14 element like carbon, their properties differ significantly.

Part A

Unlike the carbon–carbon triple bond, the silicon–silicon triple bond in a compound formulated as \( R^1\text{–Si} \equiv \text{Si} \equiv R^1 \) (R: organic substituent) is extremely reactive.

For example, it reacts with ethylene to form a cyclic product that contains a four-membered ring.

\[
R^1\text{–Si} \equiv \text{Si} \equiv R^1 + H_2C\equiv CH_2 \rightarrow R^1\text{Si} = \text{Si} = R^1
\]

When \( R^1\text{–Si} \equiv \text{Si} \equiv R^1 \) is treated with an alkyne (\( R^2\text{–C} \equiv C \equiv R^2 \)), the four-membered-ring compound \( \text{A} \) is formed as an initial intermediate.

Further reaction of another molecule of \( R^2\text{–C} \equiv C \equiv R^2 \) with \( \text{A} \) affords isomers \( \text{B} \) and \( \text{C} \), both of which have benzene-like cyclic conjugated structures, so-called ‘disilabzenes’ that contain a six-membered ring.
and can be formulated as \((R^1 \cdot Si)_2(R^2 \cdot C)_4\).

\[
R^1Si≡Si≡R^1 + R^2≡C≡C≡R^2 \rightarrow A \quad \xrightarrow{R^2\cdot C≡C≡R^2} \quad B + C
\]

The \(^{13}\text{C}\) NMR analysis of the corresponding six-membered ring skeletons \(Si_2C_4\) shows two signals for \(B\) and one signal for \(C\).

**A.1** Draw the structural formulae of \(A\), \(B\), and \(C\) using \(R^1\), \(R^2\), \(Si\), and \(C\), with one of the possible resonance structures.

**A.2** Calculate the aromatic stabilization energy (ASE) for benzene and \(C\) (in the case of \(R^1 = R^2 = H\)) as positive values, considering the enthalpy change in some hydrogenation reactions of unsaturated systems shown below (Fig. 1).

\[
\begin{align*}
\text{H}_2\text{C}≡\text{CH}_2 + \text{H}_2 & \rightarrow \text{H}_3\text{C}≡\text{CH}_3 \quad \Delta H = -135 \text{ kJ mol}^{-1} \\
\text{H}_2\text{Si}≡\text{CH}_2 + \text{H}_2 & \rightarrow \text{H}_3\text{Si}≡\text{CH}_3 \quad \Delta H = -213 \text{ kJ mol}^{-1} \\
\text{H}_2\text{Si}≡\text{SiH}_2 + \text{H}_2 & \rightarrow \text{H}_3\text{Si}≡\text{SiH}_3 \quad \Delta H = -206 \text{ kJ mol}^{-1} \\
\text{H}_2\text{C} & \rightarrow 3 \text{H}_2 \quad \Delta H = -173 \text{ kJ mol}^{-1} \\
\text{HSi}≡\text{SiH} + 3 \text{H}_2 & \rightarrow \text{H}_2\text{Si}≡\text{SiH}_2 \quad \Delta H = -326 \text{ kJ mol}^{-1} \\
\text{HSi}≡\text{SiH} + 3 \text{H}_2 & \rightarrow \text{H}_2\text{Si}≡\text{SiH}_2 \quad \Delta H = -368 \text{ kJ mol}^{-1} \\
\text{HSi}≡\text{SiH} + 3 \text{H}_2 & \rightarrow \text{H}_2\text{Si}≡\text{SiH}_2 \quad \Delta H = -389 \text{ kJ mol}^{-1}
\end{align*}
\]

Fig. 1
When a xylene solution of C is heated, it completely undergoes isomerisation to give a mixture of compounds D and E in equilibrium.

The molar ratio is $D : E = 1 : 40.0$ at 50.0 °C and $D : E = 1 : 20.0$ at 120.0 °C.

A.3 **Calculate** $\Delta H$ for the transformation of D to E. Assume that $\Delta H$ does not depend on temperature.

The isomerization from C to D and to E proceeds via transformations of $\pi$-bonds into $\sigma$-bonds without breaking any $\sigma$-bonds.

A $^{13}$C NMR analysis revealed one signal for the Si$_2$C$_4$ skeleton of D and two signals for that of E.

The skeleton of D does not contain any three-membered rings, while E has two three-membered rings that share an edge.

A.4 **Draw** the structural formulae of D and E using $R^1$, $R^2$, Si, and C.

**Part B**

Silicon is able to form highly coordinated compounds (> four substituents) with electronegative elements such as fluorine. As metal fluorides are often used as fluorination reagents, highly coordinated silicon fluorides also act as fluorination reagents.

The fluorination reaction of CCl$_4$ using Na$_2$SiF$_6$ was carried out as follows.

- **Standardization of Na$_2$SiF$_6$ solution**:
  - Preparation
    Aqueous solution F: 0.855 g of Na$_2$SiF$_6$ (188.053 g mol$^{-1}$) dissolved in water (total volume: 200 mL).
    Aqueous solution G: 6.86 g of Ce$_2$(SO$_4$)$_3$ (568.424 g mol$^{-1}$) dissolved in water (total volume: 200 mL).
  - Procedure
    Precipitation titration of a solution F (50.0 mL) by dropwise adding solution G in the presence of xylenol orange, which coordinates to Ce$^{3+}$, as an indicator.
    After adding 18.8 mL of solution G, the colour of the solution changes from yellow to magenta.
    The precipitate formed is a binary compound that contains Ce$^{3+}$, and the only resulting silicon compound is Si(OH)$_4$.

- **Reaction of CCl$_4$ with Na$_2$SiF$_6$**:
  (Material losses by *e.g.*, evaporation, are negligible during the following operations.)
  Na$_2$SiF$_6$($x$ g) was added to CCl$_4$ (500.0 g) and heated to 300 °C in a sealed pressure-resistant reaction vessel.
  The unreacted Na$_2$SiF$_6$ and generated NaCl were removed by filtration. The filtrate was diluted to a total volume of 1.00 L with CCl$_4$ (solution H).
  The $^{29}$Si and $^{19}$F NMR spectra of solution H showed SiF$_4$ as the only silicon compound.
In the $^{19}$F NMR spectrum, in addition to SiF$_4$, signals corresponding to CFCl$_3$, CF$_2$Cl$_2$, CF$_3$Cl, and CF$_4$ were observed (see Table 1). The integration ratios in the $^{19}$F NMR spectrum are proportional to the number of fluorine nuclei.

<table>
<thead>
<tr>
<th>$^{19}$F NMR data</th>
<th>CFCl$_3$</th>
<th>CF$_2$Cl$_2$</th>
<th>CF$_3$Cl</th>
<th>CF$_4$</th>
</tr>
</thead>
<tbody>
<tr>
<td>Integration ratio</td>
<td>45.0</td>
<td>65.0</td>
<td>18.0</td>
<td>2.0</td>
</tr>
</tbody>
</table>

SiF$_4$ is hydrolyzed to form H$_2$SiF$_6$ according to the following eq. 8:

$$3\text{SiF}_4 + 2\text{H}_2\text{O} \rightarrow \text{SiO}_2 + 2\text{H}_2\text{SiF}_6$$ (8)

Solution H (10 mL) was added to an excess amount of water, which resulted in the complete hydrolysis of SiF$_4$.

After separation, the H$_2$SiF$_6$ generated from the hydrolysis in the aqueous solution was neutralized and completely converted to Na$_2$SiF$_6$ (aqueous solution J).

The precipitate of unreacted Na$_2$SiF$_6$ and NaCl, which was removed by filtration in the initial step (underlined), was completely dissolved in water to give an aqueous solution (solution K; 10.0 L).

Then, additional precipitation titrations using solution G were carried out, and the endpoints of the titrations with G were as follows:
- For solution J (entire amount): 61.6 mL.
- For 100 mL of solution K: 44.4 mL.

Note: the presence of NaCl or SiO$_2$ has no effect on the precipitation titration.

**B.2 Calculate** the mass of the NaCl produced in the reaction vessel (information underlined), and **calculate** the mass (x [g]) of the Na$_2$SiF$_6$ used as a starting material.

**B.3** 77.8% of the CCl$_4$ used as a starting material was unreacted. **Calculate** the mass of CF$_3$Cl generated.
Mysterious Silicon

Part A

A.1 (9 pt)

A (3 pt)          B (3 pt)          C (3 pt)

A.2 (7 pt)

C₆H₆ :   \text{kJ mol}^{-1}, \text{C : }  \text{kJ mol}^{-1}
A.3 (6 pt)

\[ \Delta H = \text{kJ mol}^{-1} \]
| A.4 (10 pt) | D (5 pt) | E (5 pt) |
Part B

B.1 (5 pt)
B.2 (cont.)

\[ \text{NaCl : } \quad g \quad \text{Na}_2\text{SiF}_6 : \quad g \]
B.3 (8 pt)

CF$_3$Cl: __________ g
Please return this cover sheet together with all the related question sheets.
Part A

Japan has a very large number of volcanos. When silicate minerals crystallize from magma, some transition-metal ions \((M^{n+})\) in the magma are incorporated into the silicate minerals.

The \(M^{n+}\) in this problem are coordinated by oxide ions \((O^{2-})\) and adopt a four-coordinate tetrahedral \((T_d)\) geometry in the magma and six-coordinate octahedral \((O_h)\) geometry in the silicate minerals, both of which exhibit a high-spin electron configuration.

The distribution coefficient of \(M^{n+}\) between the silicate minerals and magma, \(D\), can be expressed as:

\[
D = \frac{[M]_s}{[M]_l}
\]

\([M]_s\) and \([M]_l\) are the concentrations of \(M^{n+}\) in the silicate minerals and the magma, respectively. The table below gives \(D\) for \(Cr^{2+}\) and \(Mn^{2+}\) examples.

<table>
<thead>
<tr>
<th></th>
<th>Cr(^{2+})</th>
<th>Mn(^{2+})</th>
</tr>
</thead>
<tbody>
<tr>
<td>(D)</td>
<td>7.2</td>
<td>1.1</td>
</tr>
</tbody>
</table>
Δ₀ and CFSE⁰ represent the energy separation of the d-orbitals of M⁺⁺⁺ and the crystal-field stabilization energy in a Oₙ field, respectively. Δₜ and CFSET are these in a Tₜ field.

A.1 Calculate $|\text{CFSE}^O - \text{CFSE}^T| = \Delta\text{CFSE}$ in terms of Δ₀ for Cr²⁺, Mn²⁺, and Co²⁺; assume Δₜ = 4/9Δ₀.

A.2 A linear relationship is observed when ln(D) is plotted against ΔCFSE / Δ₀ in the Cartesian coordinate system shown below. Estimate D for Co²⁺.

![Graph showing linear relationship between ln(D) and ΔCFSE / Δ₀](image)

Metal oxides MO (M: Ca, Ti, V, Mn, or Co) crystallize in a rock-salt structure wherein the M⁺⁺⁺ adopts an Oₙ geometry with a high-spin electron configuration.

The lattice enthalpy of these oxides is mainly governed by the Coulomb interactions based on the radius and charge of the ions and some contributions from the CFSE of M⁺⁺⁺ in the Oₙ field.

A.3 Choose the appropriate set of lattice enthalpies [kJ mol⁻¹] from one of the options (a) to (f).

<table>
<thead>
<tr>
<th></th>
<th>CaO</th>
<th>TiO</th>
<th>VO</th>
<th>MnO</th>
<th>CoO</th>
</tr>
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<tr>
<td>(a)</td>
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<td>3810</td>
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<tr>
<td>(b)</td>
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<td>3913</td>
</tr>
<tr>
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<td>3913</td>
</tr>
<tr>
<td>(f)</td>
<td>3810</td>
<td>3913</td>
<td>3916</td>
<td>3460</td>
<td>3878</td>
</tr>
</tbody>
</table>
**Part B**

A mixed oxide $A$, which contains $\text{La}^{3+}$ and $\text{Cu}^{2+}$, crystallizes in a tetragonal unit cell shown in Fig. 1.

In the $[\text{CuO}_6]$ octahedron, the $\text{Cu}$–$\text{O}$ length along the $z$-axis ($l_z$) is longer than that of the $x$-axis ($l_x$), and $[\text{CuO}_6]$ is distorted from the regular $O_h$ geometry.

This distortion removes the degeneracy of the $e_g$ orbitals ($d_{x^2-y^2}$ and $d_{z^2}$).

Fig. 1

$A$ can be synthesized by thermal decomposition (pyrolysis) of complex $B$, which is formed by mixing metal chlorides in dilute aqueous ammonia solution containing squaric acid $C_4H_2O_4$ (a diacid).

Pyrolysis $B$ in dry air results in a weight loss of 29.1% up to 200 °C due to the loss of waters of crystallization. This is followed by release of $\text{CO}_2$ at temperatures up to 700 °C.

The total weight loss during the formation of $A$ from $B$ is 63.6%. It should be noted that only water and $\text{CO}_2$ are released in the pyrolysis reaction.

| B.1 | Write the chemical formulae for $A$ and $B$. | 6pt |
| B.2 | Calculate $l_x$ and $l_z$ using Fig. 1. | 4pt |
| B.3 | For $\text{Cu}^{2+}$ in the distorted $[\text{CuO}_6]$ octahedron in $A$ of Fig. 1: write the names of the split $e_g$ orbitals ($d_{x^2-y^2}$ and $d_{z^2}$) in (i) and (ii). draw the electron configuration in the dotted box on your answer sheet. | 4pt |
**A** is an insulator. When one La$^{3+}$ is substituted with one Sr$^{2+}$, one hole is generated in the crystal lattice that can conduct electricity. As a result, the Sr$^{2+}$-doped **A** shows superconductivity below 38 K.

When a substitution reaction took place in **A**, $2.05 \times 10^{27}$ holes m$^{-3}$ were generated.

**B.4 Calculate** the percentage of Sr$^{2+}$ substituted for La$^{3+}$ based on the mole ratio in the substitution reaction. Note that the valences of the constituent ions and the crystal structure are not altered by the substitution reaction.

**Part C**

Cu$_2$(CH$_3$CO$_2$)$_4$ is composed of four CH$_3$CO$_2^-$ coordinated to two Cu$^{2+}$ (Fig. 2A). Cu$_2$(CH$_3$CO$_2$)$_4$ exhibits high levels of structural symmetry. Two axes pass through the carbon atoms of the four CH$_3$CO$_2^-$ and an axis passes through the two Cu$^{2+}$. All of these are oriented orthogonal relative to each other.

When a dicarboxylate ligand is used instead of CH$_3$CO$_2^-$, a “cage complex” is formed. The cage complex Cu$_4$(L1)L$_1$ is composed of planar dicarboxylate L1 (Fig. 2B) and Cu$^{2+}$ (Fig. 2C).

The angle $\theta$ between the coordination directions of the two carboxylates, indicated by the arrows in Fig. 2B, determines the structure of the cage complex. The $\theta$ is 0° for L1. Note that hydrogen atoms are not shown in Fig. 2.
C.1 The \( \theta \) of the planar dicarboxylate \( L_2 \) below is fixed to 90°. Assume the composition of the cage complex formed from \( L_2 \) and \( \text{Cu}^{2+} \) is \( \text{Cu}_{n}(L_2)_m \). Give the smallest integer combination of \( n \) and \( m \). Assume that only the \( \text{CO}_2^- \) groups of \( L_2 \) coordinate to \( \text{Cu}^{2+} \) ions.
A zinc complex, $\text{Zn}_4\text{O}(\text{CH}_3\text{CO}_2)^6$, contains four tetrahedral $\text{Zn}^{2+}$, six $\text{CH}_3\text{CO}_2^-$, and one $\text{O}^{2-}$ (Fig. 3A). In $\text{Zn}_4\text{O}(\text{CH}_3\text{CO}_2)^6$, the $\text{O}^{2-}$ is located at the origin, and the three axes passing through the carbon atoms of $\text{CH}_3\text{CO}_2^-$ are oriented orthogonal relative to each other.

When $p$-benzenedicarboxylate (Fig. 3B, $\text{L}_3$, $\theta = 180^\circ$) is used instead of $\text{CH}_3\text{CO}_2^-$, the $\text{Zn}^{2+}$ clusters are linked to each other to form a crystalline solid ($\text{X}$) that is called a “porous coordination polymer” (Fig. 3C).

The composition of $\text{X}$ is $[\text{Zn}_4\text{O}(\text{L}_3)]_n$, and it has a cubic crystal structure with nano-sized pores. One pore is represented as a sphere in Fig. 3D.

Each tetrahedral $\text{Zn}^{2+}$ cluster is represented as dark grey polyhedron in Fig. 3C and 3D. Hydrogen atoms are not shown in Fig. 3.

**Fig. 3**

---

**C.2** $\text{X}$ has a cubic unit cell with a side length of $a$ (Fig. 3C) and a density of 0.592 g cm$^{-3}$. **Calculate** $a$ in [cm].

---

**C.3** $\text{X}$ contains a considerable number of pores. 1 g of $\text{X}$ can accommodate $3.0 \times 10^2$ mL of $\text{CO}_2$ gas in the pores at 1 bar and 25 °C. **Calculate** the average number of $\text{CO}_2$ molecules per pore.
The Solid-State Chemistry of Transition Metals

Part A

A.1 (6 pt)

\[ \text{Cr}^{2+} : \Delta_0, \text{Mn}^{2+} : \Delta_0, \text{Co}^{2+} : \Delta_0 \]
A.2 (3 pt)

\[ D : \quad \text{[Blank]} \]

A.3 (3 pt)

\[ \quad \text{[Blank]} \]
Part B

B.1 (6 pt)

A: ________________________, B: _______________________

B.2 (4 pt)

\[ l_x = \quad \text{nm}, \quad l_z = \quad \text{nm} \]
B.3 (4 pt)
(i) : ____________________________ , (ii) :

\[
\begin{array}{c}
\text{Energy} \\
\text{e}_g \\
\text{(i)} \\
\text{(ii)}
\end{array}
\]

B.4 (4 pt)
Part C

C.1 (5 pt)
\[ n = \quad , \quad m = \quad \]

C.2 (5 pt)
\[ a = \quad \text{cm} \]
C.3 (5 pt)
IChO
Problem 7
Cover sheet

Please return this cover sheet together with all the related question sheets.
Prof. Nozoe (1902–1996) initiated research into non-benzenoid aromatics, compounds which are now ubiquitous in organic chemistry.

<table>
<thead>
<tr>
<th>Question</th>
<th>A.1</th>
<th>A.2</th>
<th>A.3</th>
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</table>

Score

Part A

_Inula linariifolia_
A.1 The natural product lineariifolianone, was isolated from *Inula linariifolia*. Valencene (1) can be converted in a one-step process to 2. 2 is converted in three-step process via 3 to ketone 4. Eremophilene (5) is converted into 6 using the same four-step process. Draw structures of 2 and 6 clearly show stereochemistry where necessary.

Then, ketone 4 is converted by a multi-step process into ester 15.

Compound 8 (molecular weight: 188) retains all stereocenters in 7. Compounds 9 and 10 have five stereocenters and no carbon-carbon double bonds.

Assume that H$_2^{18}$O is used instead of H$_2^{16}$O for the synthesis of $^{18}$O-labelled-lineariifolianones 13 and 14 from 11 and 12, respectively. Compounds 13 and 14 are $^{18}$O-labelled isotopomers. If the isotopic labelling is ignored, both 13 and 14 provide the same product 15 with identical stereochemistry.
\[ \text{IChO} \]

**NZL-2 C-7 Q-3**

NZL English (New Zealand)

---

**Reactions and Structures**

1. **4**
   - THF, \(-78 \, ^\circ \text{C}\)
   - Reaction with **A**

2. **8**
   - MW: 188
   - **8** + **9** + **10**
   - Reactions:
     - Chloranil (mCPBA) and 
     - \(\text{NaHCO}_3\) in 
     - 

3. **9**
   - THF, \(80 \, ^\circ \text{C}\)
   - Reaction with **11**

4. **10**
   - THF, \(80 \, ^\circ \text{C}\)
   - Reaction with **12**

5. **13**
   - \(\text{C}_1\text{H}_2\text{F}_2\text{O}\)
   - \(\text{H}_2\text{O}^{18}\), THF

6. **14**
   - \(\text{C}_1\text{H}_2\text{F}_2\text{O}\)
   - \(\text{H}_2\text{O}^{18}\), THF

---

**Isotope Labeled Compounds**

- **16O-13/16O-14**
  - \(\text{C}_1\text{H}_2\text{F}_2\text{O}\)
  - \(\text{CH}_3\text{OH}\)

---

**Additional Structures**

- **15**

---

*Chemical reactions and structures as observed in the image.*
A.2  Choose the appropriate structure for A.

I. $\text{F}_3\text{C-S-OH}$  
II. $\text{F}_3\text{C-S-NH}_2$  
III. $\text{F}_3\text{C-S-CF}_3$  
IV. $\text{F}_3\text{C-S-CF}_3$

A.3  Draw the structures of 8–14 and clearly identify the stereochemistry where necessary. Also, indicate the introduced $^{18}\text{O}$ atoms for 13 and 14 as shown in the example below.
Part B

Compound 19 is synthesized as shown below. In relation to non-benzenoid aromaticity, 19 can be used as an activator for alcohols, and 20 was converted to 22 via ion-pair intermediate 21. Although the formation of 21 was observed by NMR, 21 gradually decomposes to give 18 and 22.

\[ \text{17} \quad \text{C}_{15}\text{H}_{12}\text{Br}_2\text{O} \]

\[ \text{18} \quad \text{C}_{15}\text{H}_{10}\text{O} \]

\[ \text{19} \quad \text{C}_{15}\text{H}_{10}\text{Cl}_2 \]

\[ \text{20} + \text{19} \rightarrow \text{21} \quad \text{HCl} \rightarrow \text{22} + \text{18} \]

\(^1\text{H NMR (CD}_3\text{CN, ppm)}\)

20: \( \delta \) 7.4–7.2 (5H), 3.7 (2H), 2.8 (2H), 2.2 (1H)

21: \( \delta \) 8.5–7.3 (15H), 5.5 (2H), 3.4 (2H)

B.1 Draw the structures of 17–19 and 21. Identifying the stereochemistry is not necessary.
## Playing with Non-benzenoid Aromaticity

### Part A

<table>
<thead>
<tr>
<th>A.1 (5 pt)</th>
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### A.2 (2 pt)
A.3 (19 pt)

8 (3 pt)

9 (2 pt) 10 (2 pt)

11 (2 pt) 12 (2 pt)

13 (4 pt) 14 (4 pt)
### Part B

**B.1** (10 pt)

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Please return this cover sheet together with all the related question sheets.
Dynamic Organic Molecules and Their Chirality

11% of the total

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<th>A.1</th>
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Score

Part A
Polycyclic aromatic hydrocarbons with successive ortho-connections are called [n]carbohelicenes (n represents the number of six-membered rings) (see below). [4]Carbohelicene ([4]C) is efficiently prepared, using a photoreaction of the type shown below, via an intermediate (Int.) that is readily oxidized by iodine.

The photoreaction proceeds in a manner similar to the following example.
Note: For all of Question 8, please draw alternating single and double bonds in your answers as shown in the examples of carbohelicene. Do not use circles for conjugated \( \pi \) systems.

A.1 Draw the structures of A–C. Stereoisomers should be distinguished. 9pt

A.2 Attempts to synthesize [5]carbohelicene from the phosphonium salt above and an appropriate starting compound resulted in the formation of only a trace amount of [5]carbohelicene. This reaction gave product D having a molecular weight 2 Da lower than [5]carbohelicene. The \(^1\text{H} \) NMR chemical shifts of D are shown below.

\[ \text{D} (\delta, \text{ppm in } \text{CS}_2, \text{r.t.}), 8.85 (2\text{H}), 8.23 (2\text{H}), 8.07 (2\text{H}), 8.01 (2\text{H}), 7.97 (2\text{H}), 7.91 (2\text{H}) \]

Draw the structure of D. 3pt

[5]- and larger [n]carbohelicenes have helical chirality and interconversion between enantiomers of these helicenes is significantly slow at room temperature. The configuration of [n]carbohelicenes is defined as (M) or (P) as shown below.

![configuration diagram](image)

[n]Carbohelicenes with n larger than 4 can be separated into enantiomers by a chiral column chromatography, which was developed by Prof. Yoshio Okamoto.

Photo courtesy: The Japan Prize Foundation
Multiple helicenes are molecules that contain two or more helicene-like structures. If its helical chirality is considered, several stereoergic (stereoisomeric) centres exist in a multiple helicene. For example, compound E contains three [5]carbohelicene-like moieties in one molecule. One of the stereoisomers is described as \((P, P, P)\) as shown below.

![Diagram of compound E](image)

\((1, 2, 3) = (P, P, P)\)

A.3 The nickel-mediated trimerization of 1,2-dibromobenzene generates triphenylene as shown below. When the same reaction is applied to an enantiomer of F, \((P)-F\), multiple helicene G \((C_{66}H_{36})\) is obtained.

Assuming that interconversion between stereoisomers does not occur during the reaction, identify all possible stereoisomers of G formed in this process, without duplication.

One isomer should be drawn completely with the chirality defined as in the example, using numerical labels. Other stereoisomers should be listed with location numbers and M and P labels according to the same numbering. For instance, the other stereoisomers of E should be listed as \((1, 2, 3) = (P, M, P), (P, M, M), (P, P, M), (M, M, M), (M, M, P), (M, P, P), and (M, P, M)\).
Sumanene is a bowl-shaped hydrocarbon that was first reported in Japan in 2003. The name "sumanene" derives from a Sanskrit-Hindi word "suman" that means sunflower. The synthesis of sumanene was achieved by a reaction sequence that consists of a ring-opening and a ring-closing metathesis.

Representative metathesis reactions catalyzed by a ruthenium catalyst ($\text{Ru}^*$) are shown below.

B.1 **Draw** the structure of intermediate I (stereochemistry not required). 3pt
B.2 Starting from the optically active precursor $J$, the same reaction sequence gives the optically active sumanene derivative $K$. The stereocenters in $J$ are not inverted during the metathesis reaction. **Draw** the structure of $K$ with appropriate stereochemistry.
Dynamic Organic Molecules and Their Chirality

Part A

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<td>B (3 pt)</td>
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A.2 (3 pt)

---

Part A
A.3 (7 pt)
Part B

B.1 (3 pt)

B.2 (4 pt)
IChO
Problem 9
Cover sheet

Please return this cover sheet together with all the related question sheets.
Likes and Dislikes of Capsule

<table>
<thead>
<tr>
<th>Question</th>
<th>A.1</th>
<th>A.2</th>
<th>A.3</th>
<th>A.4</th>
<th>A.5</th>
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<td>3</td>
<td>23</td>
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</table>

Good kids don't do this, but if you cut a tennis ball on the seam, you can disassemble it into two U-shaped pieces.

Based on this idea, compounds 1 and 2 were synthesized as U-shaped molecules of different sizes. Compound 3 was prepared as a model of 1. The encapsulation behaviour of these was investigated.
The synthetic route to 2 is shown below. The elemental composition of compound 9 is C; 40.49%, H; 1.70%, and O; 17.98% by mass.
A.1 Draw the structures of 4–9. It is not necessary to show stereochemistry. Use "PMB" as an abbreviation so that you do not need to draw the structure of $p$-methoxybenzyl group shown in the scheme above.

In the mass spectrum of 1, an ion peak corresponding to its dimer (1$_2$) was clearly observed, whereas an ion peak for the dimer of 3 (3$_2$) was not observed in the spectrum of 3. In the $^1$H NMR spectra of a solution of 1$_2$, all N-H protons derived from 1 were observed to be chemically equivalent, and their chemical shift was significantly different from that of the N-H protons of 3. This indicates that hydrogen bonds are formed between the NH groups of 1 and atoms X of another molecule of 1 to form the dimeric capsule.

A.2 Circle all the appropriate atom(s) X in 1.

A.3 Give the number of the hydrogen bonds in the dimeric capsule (1$_2$).
The dimeric capsule of \(1 (I_2)\) has an internal space wherein an appropriate small molecule \(Z\) can be encapsulated. This phenomenon is expressed by the following reaction equation:

\[
Z + I_2 \rightarrow Z@I_2
\]  

(1)

The equilibrium constant for the encapsulation of \(Z\) into \(I_2\) is given as below:

\[
K_a = \frac{[Z@I_2]}{[Z][I_2]}
\]  

(2)

Encapsulation of a molecule can be monitored by NMR spectroscopy. For example, \(I_2\) in \(C_6D_6\) gave different signals in the \(^1H\) NMR spectra before and after addition of \(CH_4\).

Compound \(2\) also forms a rigid and larger dimeric capsule \((2_2)\). The \(^1H\) NMR spectrum of \(2_2\) was measured in \(C_6D_6\), \(C_6D_5F\), and a \(C_6D_6/C_6D_5F\) solvent mixture, with all other conditions being kept constant. The chemical shifts for the \(H^a\) proton of \(2\) in the above solvents are summarized below. No other signals from the \(H^a\) in \(2\), except for those listed, were observed. Assume that the interior of the capsule is always filled with the largest possible number of solvent molecules and that each signal corresponds to one species of the filled capsule.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>(\delta) (ppm) of (H^a)</th>
</tr>
</thead>
<tbody>
<tr>
<td>(C_6D_6)</td>
<td>4.60</td>
</tr>
<tr>
<td>(C_6D_5F)</td>
<td>4.71</td>
</tr>
<tr>
<td>(C_6D_6 / C_6D_5F)</td>
<td>4.60, 4.71, 4.82</td>
</tr>
</tbody>
</table>

A.4 **Determine** the number of \(C_6D_6\) and \(C_6D_5F\) molecules encapsulated in \(2_2\) giving each \(H^a\) signal.
$^1$H NMR measurements in C$_6$D$_6$ revealed that 2$_2$ can incorporate one molecule of 1-adamantanecarboxylic acid (AdA).

The association constants ($K_a$) which are expressed below were determined for various temperatures. [solvent@2$_2$] denotes a species containing one or more solvent molecules.

$$K_a = \frac{[Z@2_2]}{[Z][\text{solvent@2}_2]} \quad (3)$$

Similarly, $K_a$ for association of CH$_4$ and 1$_2$ (see eq (2)), was determined at various temperatures by $^1$H NMR measurements in C$_6$D$_6$. The plots of the two association constants (as ln $K_a$ vs 1/T) are shown below.

No C$_6$D$_6$ molecule is encapsulated in 1$_2$. In line II, the entropy change ($\Delta S$) is (1) and enthalpy change ($\Delta H$) is (2), indicating that the driving force for the encapsulation in line II is (3). Therefore, line I corresponds to (4), and line II corresponds to (5).

A.5 Choose the correct option, from A and B, for each of the gaps (1)–(5) in the paragraph above.

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<th>B</th>
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<td>negative</td>
</tr>
<tr>
<td>(2)</td>
<td>positive</td>
<td>negative</td>
</tr>
<tr>
<td>(3)</td>
<td>$\Delta S$</td>
<td>$\Delta H$</td>
</tr>
<tr>
<td>(4)</td>
<td>1$_2$ and CH$_4$</td>
<td>2$_2$ and AdA</td>
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<tr>
<td>(5)</td>
<td>1$_2$ and CH$_4$</td>
<td>2$_2$ and AdA</td>
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</table>
# Likes and Dislikes of Capsule

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<td>9 (2 pt)</td>
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</tbody>
</table>
A.2 (2 pt)

A.3 (2 pt)

A.4 (3 pt)

\[
\begin{array}{|c|c|c|}
\hline
\delta (\text{ppm}) \text{ of } H^a & \text{numbers of C}_6\text{D}_6 & \text{numbers of C}_6\text{D}_5\text{F} \\
\hline
4.60 \text{ ppm} & \text{ } & \text{ } \\
\hline
4.71 \text{ ppm} & \text{ } & \text{ } \\
\hline
4.82 \text{ ppm} & \text{ } & \text{ } \\
\hline
\end{array}
\]

A.5 (3 pt)

(1) : (2) : (3) :

(4) : (5) :