

ISR-2 C-0 C-1

**ISR-2 C-0 C**  
Neta Eiger

**IChO**  
**General instructions**  
**Cover sheet**

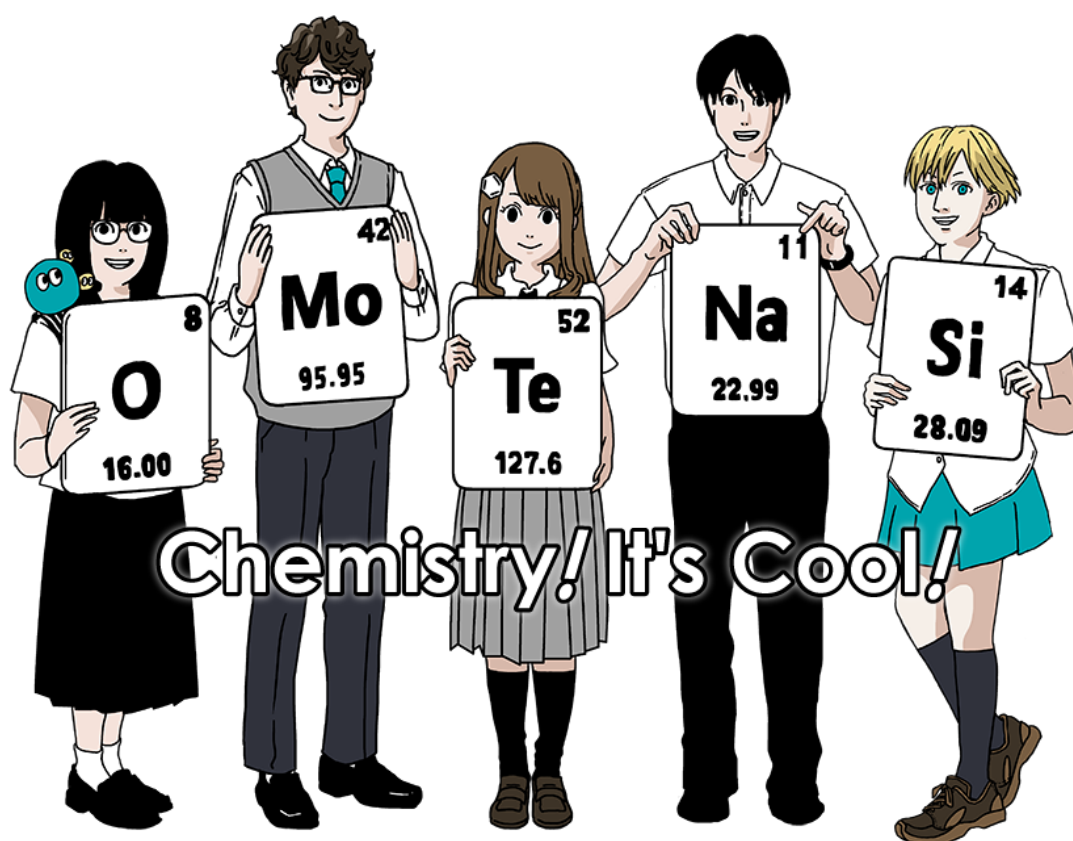
Please return this cover sheet together with all the related question sheets.

International Chemistry Olympiad 2021 Japan

53rd IChO2021 Japan

25th July - 2nd August, 2021

<https://www.icho2021.org>





## General Instruction

- You are allowed to use only pen to write the answer.
- Your calculator must be non-programmable.
- This examination has **9 problems**.
- You can solve the problems in any order.
- You will have **5 hours** to solve all problems.
- You can **begin** working only after the **START** command is given.
- All results must be written in the appropriate answer boxes with pen on the **answer sheets**. Use the back of the question sheets if you need scratch paper. Remember that answers written outside the answer boxes will not be graded.
- Write relevant calculations in the appropriate boxes when necessary. Full marks will be given for correct answers **only when your work is shown**.
- The invigilator will announce a **30-minute** warning before the **STOP** command.
- You **must stop** working when the **STOP** command is given. Failure to stop writing will lead to the nullification of your examination.
- The official English version of this examination is available on request only for clarification.
- You are not allowed to leave your working place without permission. If you need any assistance (broken calculator, need to visit a restroom, etc), raise your hand and wait until an invigilator arrives.

**GOOD LUCK!**

## Problems and Grading Information

	<b>Title</b>	<b>Total Score</b>	<b>Percentage</b>
<b>1</b>	Hydrogen at a Metal Surface	<b>24</b>	<b>11</b>
<b>2</b>	Isotope Time Capsule	<b>35</b>	<b>11</b>
<b>3</b>	Lambert–Beer Law?	<b>22</b>	<b>8</b>
<b>4</b>	The Redox Chemistry of Zinc	<b>32</b>	<b>11</b>
<b>5</b>	Mysterious Silicon	<b>60</b>	<b>12</b>
<b>6</b>	The Solid-State Chemistry of Transition Metals	<b>45</b>	<b>13</b>
<b>7</b>	Playing with Non-benzenoid Aromaticity	<b>36</b>	<b>13</b>
<b>8</b>	Dynamic Organic Molecules and Their Chirality	<b>26</b>	<b>11</b>
<b>9</b>	Likes and Dislikes of Capsules	<b>23</b>	<b>10</b>
		<b>Total</b>	<b>100</b>

## Physical Constants and Equations

### Constants

Speed of light in vacuum	$c = 2.99792458 \times 10^8 \text{ m s}^{-1}$
Planck constant	$h = 6.62607015 \times 10^{-34} \text{ J s}$
Elementary charge	$e = 1.602176634 \times 10^{-19} \text{ C}$
Electron mass	$m_e = 9.10938370 \times 10^{-31} \text{ kg}$
Electric constant (permittivity of vacuum)	$\epsilon_0 = 8.85418781 \times 10^{-12} \text{ F m}^{-1}$
Avogadro constant	$N_A = 6.02214076 \times 10^{23} \text{ mol}^{-1}$
Boltzmann constant	$k_B = 1.380649 \times 10^{-23} \text{ J K}^{-1}$
Faraday constant	$F = N_A \times e = 9.64853321233100184 \times 10^4 \text{ C mol}^{-1}$
Gas constant	$R = N_A \times k_B = 8.31446261815324 \text{ J K}^{-1} \text{ mol}^{-1}$ $= 8.2057366081 \times 10^{-2} \text{ L atm K}^{-1} \text{ mol}^{-1}$
Unified atomic mass unit	$u = 1 \text{ Da} = 1.66053907 \times 10^{-27} \text{ kg}$
Standard pressure	$p = 1 \text{ bar} = 10^5 \text{ Pa}$
Atmospheric pressure	$p_{\text{atm}} = 1.01325 \times 10^5 \text{ Pa}$
Zero degree Celsius	$0^\circ \text{C} = 273.15 \text{ K}$
Ångstrom	$1 \text{ Å} = 10^{-10} \text{ m}$
Picometer	$1 \text{ pm} = 10^{-12} \text{ m}$
Electronvolt	$1 \text{ eV} = 1.602176634 \times 10^{-19} \text{ J}$
Part-per-million	$1 \text{ ppm} = 10^{-6}$
Part-per-billion	$1 \text{ ppb} = 10^{-9}$
Part-per-trillion	$1 \text{ ppt} = 10^{-12}$
pi	$\pi = 3.141592653589793$
The base of the natural logarithm (Euler's number)	$e = 2.718281828459045$

## Equations

The ideal gas law	$PV = nRT$ , where $P$ is the pressure, $V$ is the volume, $n$ is the amount of substance, $T$ is the absolute temperature of ideal gas.
Coulomb's law	$F = k_e \frac{q_1 q_2}{r^2}$ , where $F$ is the electrostatic force, $k_e (\approx 9.0 \times 10^9 \text{ N m}^2 \text{ C}^{-2})$ is Coulomb's constant, $q_1$ and $q_2$ are the magnitudes of the charges, and $r$ is the distance between the charges.
The first law of thermodynamics	$\Delta U = q + w$ , where $\Delta U$ is the change in the internal energy, $q$ is the heat supplied, $w$ is the work done.
Enthalpy $H$	$H = U + PV$
Entropy based on Boltzmann's principle $S$	$S = k_B \ln W$ , where $W$ is the number of microstates.
The change of entropy $\Delta S$	$\Delta S = \frac{q_{\text{rev}}}{T}$ , where $q_{\text{rev}}$ is the heat for the reversible process.
Gibbs free energy $G$	$G = H - TS$ $\Delta_r G^\circ = -RT \ln K = -zFE^\circ$ , where $K$ is the equilibrium constant, $z$ is the number of electrons, $E^\circ$ is the standard electrode potential.
Reaction quotient $Q$	$\Delta_r G = \Delta_r G^\circ + RT \ln Q$ For a reaction $aA + bB \rightleftharpoons cC + dD$ $Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ , where $[A]$ is the concentration of A.

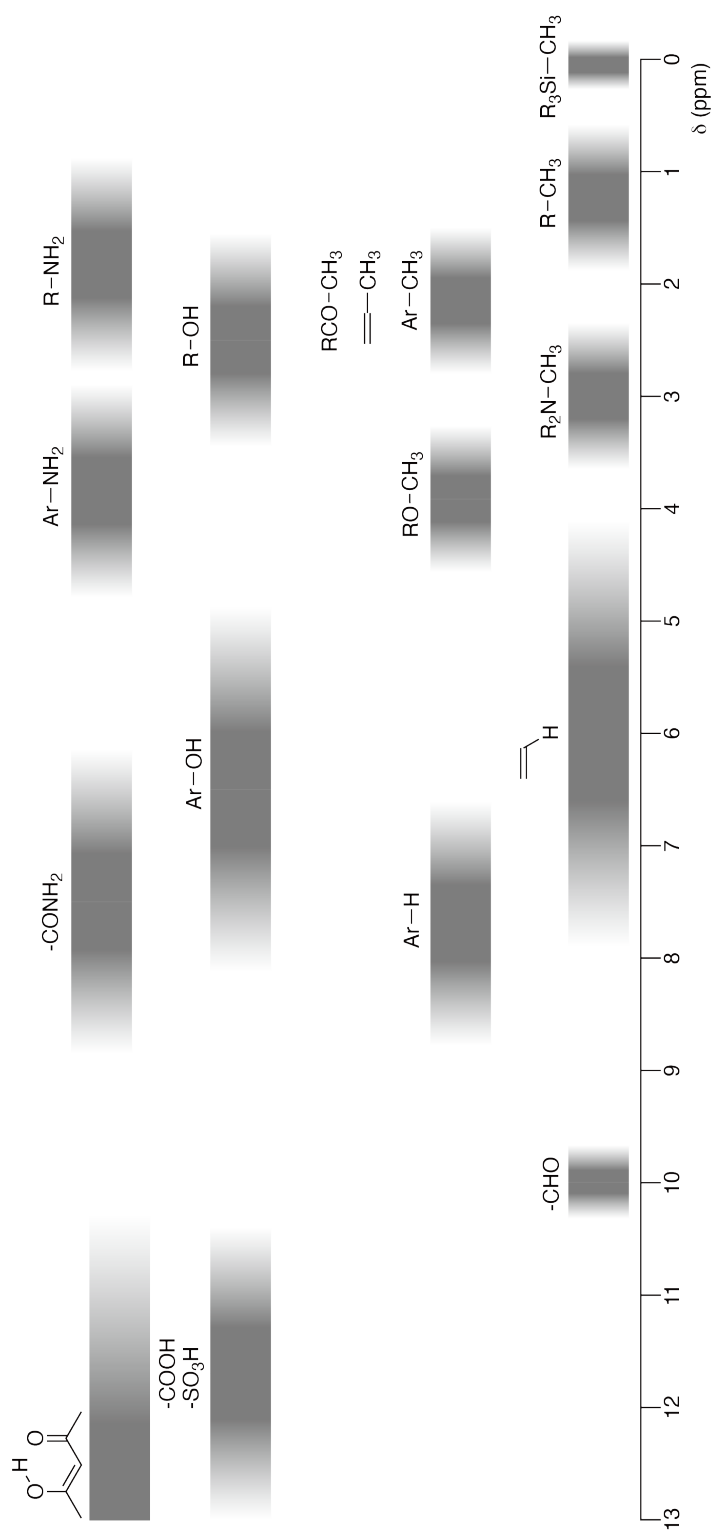
Heat change $\Delta q$	$\Delta q = nc_m \Delta T$ , where $c_m$ is the temperature-independent molar heat capacity.
Nernst equation for redox reaction	$E = E^\circ + \frac{RT}{zF} \ln \frac{C_{\text{ox}}}{C_{\text{red}}}$ , where $C_{\text{ox}}$ is the concentration of oxidized substance, $C_{\text{red}}$ is the concentration of reduced substance.
Arrhenius equation	$k = A \exp\left(-\frac{E_a}{RT}\right)$ , where $k$ is the rate constant, $A$ is the pre-exponential factor, $E_a$ is the activation energy. $\exp(x) = e^x$
Lambert-Beer equation	$A = \varepsilon lc$ , where $A$ is the absorbance, $\varepsilon$ is the molar absorption coefficient, $l$ is the optical path length, $c$ is the concentration of the solution.
Henderson-Hasselbalch equation	For an equilibrium $\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-$ , where equilibrium constant is $K_a$ , $\text{pH} = \text{p}K_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$
Energy of a photon	$E = h\nu = h\frac{c}{\lambda}$ , where $\nu$ is the frequency, $\lambda$ is the wavelength of the light.
The sum of a geometric series	When $x \neq 1$ , $1 + x + x^2 + \dots + x^n = \sum_{i=0}^n x^i = \frac{1 - x^{n+1}}{1 - x}$
Approximation equation that can be used to solve problems	When $x \ll 1$ , $\frac{1}{1 - x} \simeq 1 + x$

## Periodic Table

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H Hydrogen 1.008																	2 He Helium 4.003
3 Li Lithium 6.968	4 Be Beryllium 9.012															9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.306															17 Cl Chlorine 35.452	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.630	33 As Arsenic 74.922	34 Se Selenium 78.971	35 Br Bromine 79.904	36 Kr Krypton 83.798
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium [99]	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.414	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.760	52 Te Tellurium 127.60	53 I Iodine 126.904	54 Xe Xenon 131.293
55 Cs Caesium 132.905	56 Ba Barium 137.327	57-71 La-Lu Lanthanoids	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.084	79 Au Gold 196.967	80 Hg Mercury 200.592	81 Tl Thallium 204.384	82 Pb Lead 207.2	83 Bi Bismuth 208.98	84 Po Polonium [210]	85 At Astatine [210]	86 Rn Radon [222]
87 Fr Francium [223]	88 Ra Radium [226]	89-103 Ac-Lr Actinoids	104 Rf Rutherfordium [267]	105 Db Dubnium [268]	106 Sg Seaborgium [271]	107 Bh Bohrium [272]	108 Hs Hassium [277]	109 Mt Meitnerium [276]	110 Ds Darmstadtium [281]	111 Rg Roentgenium [280]	112 Cn Copernicium [285]	113 Nh Nihonium [278]	114 Fl Flerovium [289]	115 Mc Moscovium [289]	116 Lv Livermorium [293]	117 Ts Tennessine [293]	118 Og Oganesson [294]
57-71 La-Lu Lanthanoids		57 La Lanthanum 138.905	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.242	61 Pm Promethium [145]	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.500	67 Ho Holmium 164.930	68 Er Erbium 167.259	69 Tm Thulium 168.934	70 Yb Ytterbium 173.045	71 Lu Lutetium 174.967	
89-103 Ac-Lr Actinoids		89 Ac Actinium [227]	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium [237]	94 Pu Plutonium [239]	95 Am Americium [243]	96 Cm Curium [247]	97 Bk Berkelium [247]	98 Cf Californium [252]	99 Es Einsteinium [252]	100 Fm Fermium [257]	101 Md Mendelevium [258]	102 No Nobelium [259]	103 Lr Lawrencium [262]	

Key:  
 113  
 Nh  
 Nihonium  
 [278]  
 atomic number  
 Symbol  
 name  
 atomic weight [in parenthesis for the radioactive element]

**<sup>1</sup>H NMR Chemical Shifts**



$\Delta\delta$  for one alkyl group-substitution: *ca.* +0.4 ppm



ISR-2 C-1 C-1

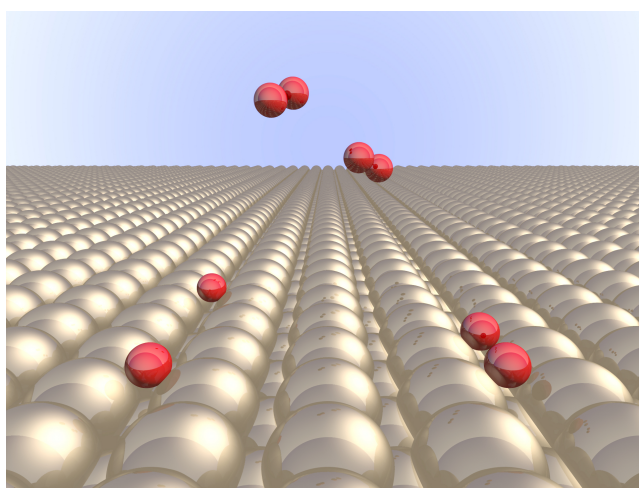
**ISR-2 C-1 C**  
Neta Eiger

**IChO**  
**Problem 1**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Hydrogen at a Metal Surface

11 % of the total							
Question	A.1	A.2	B.1	B.2	B.3	B.4	Total
Points	6	4	5	3	3	3	<b>24</b>
Score							



Hydrogen is expected to be a future energy source that does not depend on fossil fuels. Here, we will consider the hydrogen-storage process in a metal, which is related to hydrogen-transport and -storage technology.

### Part A

As hydrogen is **absorbed** (Hebrew: *nivla*) into the bulk of a metal via its surface, let us first consider the **adsorption** process (Hebrew: *sfih'a*) of hydrogen at the metal surface,  $\text{H}_2(\text{g}) \rightarrow 2\text{H}(\text{ad})$ , where the gaseous and **adsorbed** states of hydrogen are represented as (g) and (ad), respectively.

Hydrogen molecules ( $\text{H}_2$ ) that reach the metal surface (M) dissociate at the surface and are **adsorbed** as H atoms (Fig. 1). Here, the potential energy of  $\text{H}_2$  is represented by two variables: the interatomic distance,  $d$ , and the height relative to the surface metal atom,  $z$ .

It is assumed that the axis along the two H atoms is parallel to the surface and that the center of gravity is always on the vertical dotted line in Fig. 1. Fig. 2 shows the potential energy contour plot for the dissociation at the surface.

The numerical values represent the potential energy in units of kJ per mole of  $\text{H}_2$ . The solid line spacing is  $20 \text{ kJ mol}^{-1}$ , the dashed line spacing is  $100 \text{ kJ mol}^{-1}$ , and the spacing between solid and dashed lines is  $80 \text{ kJ mol}^{-1}$ . The zero-point vibration energy is ignored.

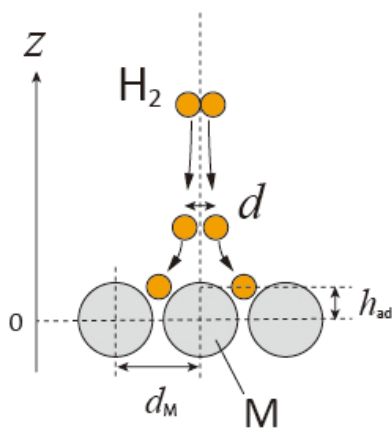


Fig.1 Definition of variables. Drawing is not in scale.

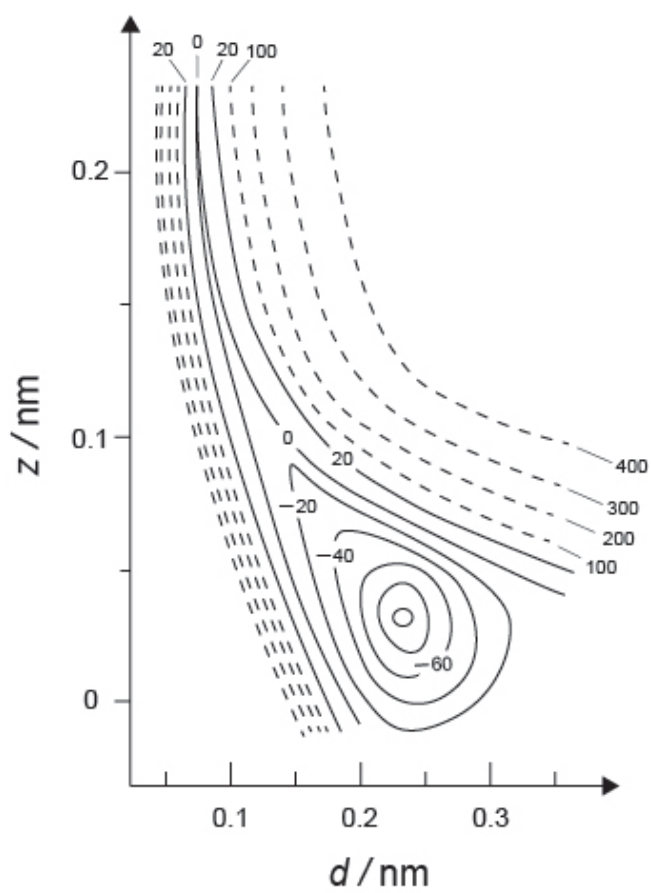


Fig.2

- A.1** For each of the following items (i)–(iii), **select** the closest value from A–G. 6pt
- (i) The interatomic distance for a gaseous  $\text{H}_2$  molecule
  - (ii) The interatomic distance between metal atoms ( $d_M$  in Fig. 1)
  - (iii) The distance of adsorbed H atoms from the surface ( $h_{ad}$  in Fig. 1)

A. 0.03 nm   B. 0.07 nm   C. 0.11 nm   D. 0.15 nm  
E. 0.19 nm   F. 0.23 nm   G. 0.27 nm

- A.2** For each of the following items (i)–(ii), **select** the closest value from A–H. 4pt
- (i) the energy required for the dissociation of gaseous  $\text{H}_2$  to gaseous H  
[ $\text{H}_2(\text{g}) \rightarrow 2\text{H}(\text{g})$ ]
  - (ii) the energy released during the adsorption of gaseous  $\text{H}_2$  [ $\text{H}_2(\text{g}) \rightarrow 2\text{H}(\text{ad})$ ]

A. 20  $\text{kJ mol}^{-1}$    B. 40  $\text{kJ mol}^{-1}$    C. 60  $\text{kJ mol}^{-1}$    D. 100  $\text{kJ mol}^{-1}$   
E. 150  $\text{kJ mol}^{-1}$    F. 200  $\text{kJ mol}^{-1}$    G. 300  $\text{kJ mol}^{-1}$    H. 400  $\text{kJ mol}^{-1}$

## Part B

The adsorbed hydrogen atoms are then either absorbed into the bulk, or recombine and desorb back into the gas phase, as shown in the reactions (1a) and (1b). H(ab) represents a hydrogen atom absorbed in the bulk.



The reaction rates per surface site for adsorption, desorption, and absorption are  $r_1[\text{s}^{-1}]$ ,  $r_2[\text{s}^{-1}]$  and  $r_3[\text{s}^{-1}]$ , respectively. They are expressed as:

$$r_1 = k_1 P_{\text{H}_2} (1 - \theta)^2 \quad (2)$$

$$r_2 = k_2 \theta^2 \quad (3)$$

$$r_3 = k_3 \theta \quad (4)$$

where  $k_1 [\text{s}^{-1} \text{Pa}^{-1}]$ ,  $k_2 [\text{s}^{-1}]$  and  $k_3 [\text{s}^{-1}]$  are the reaction rate constants and  $P_{\text{H}_2}$  is the pressure of  $\text{H}_2$ . Among the sites available on the surface,  $\theta$  ( $0 \leq \theta \leq 1$ ) is the fraction occupied by H atoms. It is assumed that adsorption and desorption are fast compared to absorption ( $r_1, r_2 \gg r_3$ ) and that  $\theta$  remains constant.

**B.1**  $r_3$  can be expressed as:

5pt

$$r_3 = \frac{k_3}{1 + \sqrt{\frac{1}{P_{\text{H}_2} C}}} \quad (5)$$

**Express**  $C$  using  $k_1$  and  $k_2$ .

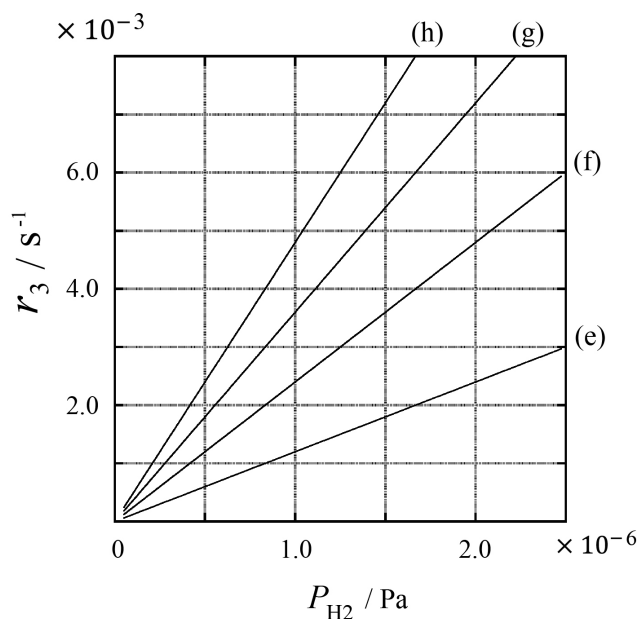
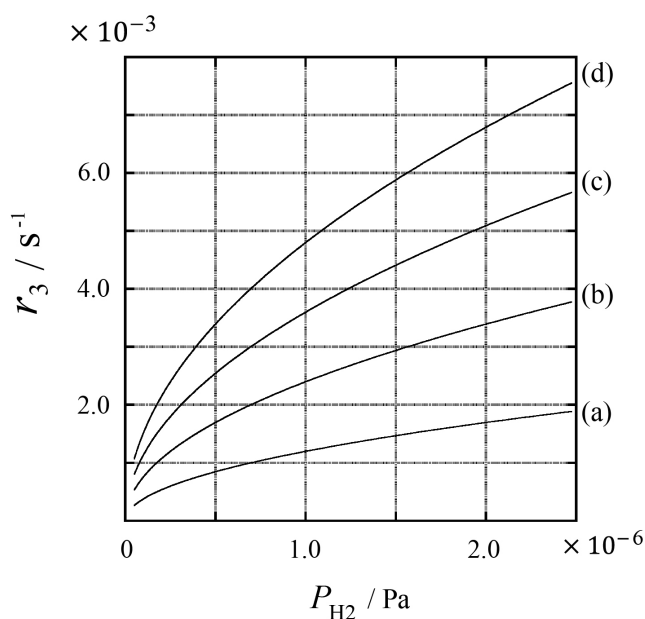
A metal sample with a surface area of  $S = 1.0 \times 10^{-3} \text{ m}^2$  was placed in a container ( $V = 1.0 \times 10^{-3} \text{ m}^3$ ) with  $\text{H}_2$  ( $P_{\text{H}_2} = 1.0 \times 10^2 \text{ Pa}$ ). The density of hydrogen-atom adsorption sites on the surface was  $N = 1.3 \times 10^{18} \text{ m}^{-2}$ . The surface temperature was kept at  $T = 400 \text{ K}$ .

As the reactions (1a) and (1b) proceeded,  $P_{\text{H}_2}$  decreased at a constant rate of  $v = 4.0 \times 10^{-4} \text{ Pa s}^{-1}$ . Assume that  $\text{H}_2$  is an ideal gas and that the volume of the metal sample is negligible.

**B.2** Calculate the amount of H atoms, in moles absorbed per unit area of the surface per unit time,  $A$  [ $\text{mol s}^{-1} \text{ m}^{-2}$ ]. 3pt

**B.3** At  $T = 400 \text{ K}$ ,  $C$  equals  $1.0 \times 10^2 \text{ Pa}^{-1}$ . Calculate the value of  $k_3$  at 400 K. If you did not obtain the answer to **B.2**, use  $A = 3.6 \times 10^{-7} \text{ mol s}^{-1} \text{ m}^{-2}$ . 3pt

**B.4** At a different  $T$ ,  $C = 2.5 \times 10^3 \text{ Pa}^{-1}$  and  $k_3 = 4.8 \times 10^{-2} \text{ s}^{-1}$  are given. For  $r_3$  as a function of  $P_{\text{H}_2}$  at this temperature, select the correct plot from (a)–(h). 3pt





ISR-2 C-1 A-1

**A1-1**  
ISR English (Israel)

## Hydrogen at a Metal Surface

### Part A

**A.1** (6 pt)

(i)	(ii)	(iii)

**A.2** (4 pt)

(i)	(ii)



ISR-2 C-1 A-2

**A1-2**  
ISR English (Israel)

**Part B**

**B.1** (5 pt)

$C =$  \_\_\_\_\_

**B.2** (3 pt)

$A =$  \_\_\_\_\_  $\text{mol s}^{-1} \text{m}^{-2}$





ISR-2 C-1 A-3

**A1-3**  
ISR English (Israel)

**B.3** (3 pt)

$$k_3 = \underline{\hspace{10em}} \text{ s}^{-1}$$

**B.4** (3 pt)

\_\_\_\_\_

ISR-2 C-2 C-1

**ISR-2 C-2 C**  
Neta Eiger

**ICHO**  
**Problem 2**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Isotope Time Capsule

11 % of the total					
Question	A.1	A.2	A.3	A.4	Total
Points	8	8	10	9	<b>35</b>
Score					



Molecular entities that differ only in isotopic composition, such as  $\text{CH}_4$  and  $\text{CH}_3\text{D}$ , are called isotopologues. Isotopologues are considered to have the same chemical characteristics. In nature, however, there exists a slight difference.

Assume that all of the substances shown in this question are in the gas phase.

Let us consider the following equilibrium:

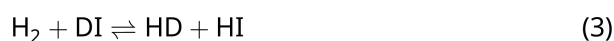


The entropy,  $S$ , increases with increasing the number of possible microscopic states of a system,  $W$ :

$$S = k_B \ln W \quad (2)$$

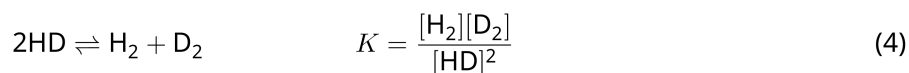
$W = 1$  for  $^{12}\text{C}^{16}\text{O}_2$  and  $^{12}\text{C}^{18}\text{O}_2$ . In contrast,  $W = 2$  for a  $^{12}\text{C}^{16}\text{O}^{18}\text{O}$  molecule because the oxygen atoms are distinguishable in this molecule. Since the right-hand side of the equilibrium shown in eq. 1 has two  $^{12}\text{C}^{16}\text{O}^{18}\text{O}$  molecules,  $W = 2^2 = 4$ .

**A.1** The enthalpy change,  $\Delta H$ , of eq. 3 is positive regardless of the temperature. 8pt



**Calculate** the equilibrium constants,  $K$ , for eq. 3 at very low (think of  $T \rightarrow 0$ ) and very high (think of  $T \rightarrow +\infty$ ) temperatures. Assume that the reaction remains unchanged at these temperatures and that  $\Delta H$  converges to a constant value for high temperatures.

The  $\Delta H$  of the following process can be explained by molecular vibrations.



At  $T = 0$  K, the vibrational energy of a diatomic molecule whose vibration frequency is  $\nu$  [ $\text{s}^{-1}$ ] is expressed as:

$$E = \frac{1}{2}h\nu \quad (5)$$

$$\nu = \frac{1}{2\pi} \sqrt{\frac{k}{\mu}} \quad (6)$$

Wherein  $k$  is the force constant and  $\mu$  is the reduced mass, which is expressed in terms of the mass of the two atoms in the diatomic molecule,  $m_1$  and  $m_2$ , according to:

$$\mu = \frac{m_1 m_2}{m_1 + m_2} \quad (7)$$

**A.2** The vibration of  $\text{H}_2$  is at  $4161.0 \text{ cm}^{-1}$  when reported as a wavenumber. 8pt  
**Calculate** the  $\Delta H$  of the following equation at  $T = 0$  K in units of  $\text{J mol}^{-1}$ .



Assume that:

- only the vibrational energy contributes to the  $\Delta H$ .
- the  $k$  values for  $\text{H}_2$ ,  $\text{HD}$ , and  $\text{D}_2$  are identical.
- the mass of H to be 1 Da and the mass of D to be 2 Da.

In a system in equilibrium, the molar ratio of  $\text{H}_2$ ,  $\text{HD}$ , and  $\text{D}_2$  depends on the temperature. Here,  $\Delta_{\text{D}_2}$  is defined as the change of the molar ratio of  $\text{D}_2$ .

$$\Delta_{\text{D}_2} = \frac{R_{\text{D}_2}}{R_{\text{D}_2}^*} - 1 \quad (9)$$

Here,  $R_{\text{D}_2}$  refers to  $\frac{[\text{D}_2]}{[\text{H}_2]}$  in the sample and  $R_{\text{D}_2}^*$  to  $\frac{[\text{D}_2]}{[\text{H}_2]}$  at  $T \rightarrow +\infty$ . It should be noted here that the distribution of isotopes becomes random at  $T \rightarrow +\infty$ .

**A.3** At the temperature where  $K$  in eq. 4 is 0.300, **Calculate**  $\Delta_{\text{D}_2}$  with natural D 10pt abundance when the isotopic exchange is in equilibrium. Assume that the natural abundance ratios of D and H are  $1.5576 \times 10^{-4}$  and  $1 - 1.5576 \times 10^{-4}$ , respectively.

In general, the molar ratio of the doubly substituted isotopologue, which contains two heavy isotope atoms in one molecule, increases with decreasing temperature. Let us consider the molar ratio of CO<sub>2</sub> molecules with molecular weights of 44 and 47, which are described as CO<sub>2</sub>[44] and CO<sub>2</sub>[47] below. The quantity  $\Delta_{47}$  is defined as:

$$\Delta_{47} = \frac{R_{47}}{R_{47}^*} - 1 \quad (10)$$

$R_{47}$  refers to  $\frac{[\text{CO}_2[47]]}{[\text{CO}_2[44]]}$  in the sample and  $R_{47}^*$  to  $\frac{[\text{CO}_2[47]]}{[\text{CO}_2[44]]}$  at  $T \rightarrow +\infty$ . The natural abundances of carbon and oxygen atoms are shown below; ignore isotopes that are not shown here.

	<sup>12</sup> C	<sup>13</sup> C
natural abundance	0.988888	0.011112

	<sup>16</sup> O	<sup>17</sup> O	<sup>18</sup> O
natural abundance	0.997621	0.0003790	0.0020000

The temperature dependence of  $\Delta_{47}$  is determined as follows, where  $T$  is given as the absolute temperature in units of K:

$$\Delta_{47} = \frac{36.2}{T^2} + 2.920 \times 10^{-4} \quad (11)$$

- A.4** Consider only the most common isotopologue of CO<sub>2</sub>[47] for the calculation. 9pt  
 The  $R_{47}$  of fossil plankton obtained from the Antarctic seabed was  $4.50865 \times 10^{-5}$ .  
**Estimate** the temperature using this  $R_{47}$ . This temperature is interpreted as the air temperature during the era in which the plankton lived.



ISR-2 C-2 A-1

**A2-1**  
ISR English (Israel)

## Isotope Time Capsule

**A.1** (8 pt)

$T \rightarrow 0 : K =$  \_\_\_\_\_ ,  $T \rightarrow +\infty : K =$  \_\_\_\_\_



ISR-2 C-2 A-2

**A2-2**  
ISR English (Israel)

**A.2** (8 pt)

$\Delta H =$  \_\_\_\_\_  $\text{J mol}^{-1}$





ISR-2 C-2 A-3

**A2-3**  
ISR English (Israel)

**A.3** (10 pt)

$\Delta_{D_2} =$  \_\_\_\_\_



ISR-2 C-2 A-4

**A2-4**  
ISR English (Israel)

**A.4** (9 pt)

$T =$  \_\_\_\_\_  $K$

ISR-2 C-3 C-1

**ISR-2 C-3 C**  
Neta Eiger

**IChO**  
**Problem 3**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

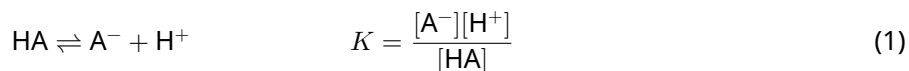
## Lambert-Beer Law?

8 % of the total				
Question	A.1	B.1	B.2	Total
Points	10	6	6	<b>22</b>
Score				

In this problem, ignore the absorption of the cell and the solvent. The temperatures of all solutions and gases are kept constant at 25 °C.

### Part A

An aqueous solution **X** was prepared using HA and NaA. The concentrations  $[A^-]$ ,  $[HA]$ , and  $[H^+]$  in solution **X** are  $1.00 \times 10^{-2} \text{ mol L}^{-1}$ ,  $1.00 \times 10^{-3} \text{ mol L}^{-1}$ , and  $1.00 \times 10^{-4} \text{ mol L}^{-1}$ , respectively, which are correlated via the following acid-base equilibrium:



The optical path length is  $l$  in Part A. Ignore the density change upon dilution. Assume that no chemical reactions other than eq 1 occur.

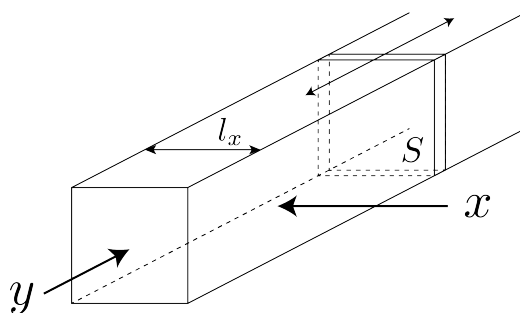
- A.1** The absorbance of **X** was  $A_1$  at a wavelength of  $\lambda_1$ . Then, solution **X** was diluted to twice its initial volume using hydrochloric acid solution of pH = 2.500. After the dilution, the absorbance was still  $A_1$  at  $\lambda_1$ . **Determine** the ratio  $\varepsilon_{HA}/\varepsilon_{A^-}$ , where  $\varepsilon_{HA}$  and  $\varepsilon_{A^-}$  represent the absorption coefficients of HA and of  $A^-$ , respectively, at  $\lambda_1$ . 10pt

**Part B**

Let us consider the following equilibrium in the gas phase.



Pure gas D is filled into a cuboid container that has a transparent movable wall with a cross-section of  $S$  (see the figure below) at a pressure  $P$ , and equilibrium is established while the total pressure is kept at  $P$ . The absorbance of the gas is  $A = \varepsilon(n/V)l$ , where  $\varepsilon$ ,  $n$ ,  $V$ , and  $l$  are the absorption coefficient, amount of the gas in moles, volume of the gas, and optical path length, respectively. Assume that all components of the gas mixture behave as ideal gases.



Use the following definitions if necessary.

	Initial state		After equilibrium	
	D	M	D	M
Partial pressure	$P$	0	$p_D$	$p_M$
Amount in moles	$n_0$	0	$n_D$	$n_M$
Volume	$V_0$		$V$	

**B.1** The absorbance of the gas at  $\lambda_{B1}$  measured from direction  $x$  ( $l = l_x$ ) was  $A_{B1}$  both at the initial state and after the equilibrium. **Determine** the ratio  $\varepsilon_D/\varepsilon_M$  at  $\lambda_{B1}$ , where  $\varepsilon_D$  and  $\varepsilon_M$  represent the absorption coefficients of D and of M, respectively. 6pt

**B.2** The absorbance of the gas at  $\lambda_{B2}$  measured from direction  $y$  was  $A_{B2}$  both at the initial state ( $l = l_{y0}$ ) and after the equilibrium ( $l = l_y$ ). **Determine** the ratio  $\varepsilon_D/\varepsilon_M$  at  $\lambda_{B2}$ . 6pt



ISR-2 C-3 A-1

**A3-1**  
ISR English (Israel)

## Lambert-Beer Law?

### Part A

**A.1** (10 pt)

(Continued on the next page)



ISR-2 C-3 A-2

**A3-2**  
ISR English (Israel)

A.1 (cont.)

$\epsilon_{HA}/\epsilon_{A^-} =$   
\_\_\_\_\_



ISR-2 C-3 A-3

**A3-3**  
ISR English (Israel)

**Part B**

**B.1** (6 pt)

$\epsilon_D/\epsilon_M =$  \_\_\_\_\_





ISR-2 C-3 A-4

**A3-4**  
ISR English (Israel)

**B.2** (6 pt)

$\epsilon_D/\epsilon_M =$   
\_\_\_\_\_

ISR-2 C-4 C-1

**ISR-2 C-4 C**  
Neta Eiger

**ICHO**  
**Problem 4**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## The Redox Chemistry of Zinc

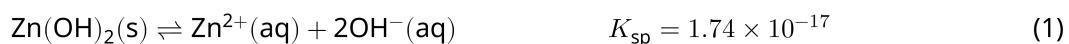
11 % of the total							
Question	A.1	A.2	B.1	B.2	B.3	B.4	Total
Points	6	5	4	3	5	9	<b>32</b>
Score							



Zinc has long been used in alloys of brass and steel materials. The zinc contained in industrial wastewater is separated by precipitation to detoxify the water, and the obtained precipitate is reduced to recover and reuse it as metallic zinc.

### Part A

The dissolution equilibrium of zinc hydroxide  $\text{Zn(OH)}_2(\text{s})$  at 25 °C and the relevant equilibrium constants are given in eq. 1-4.



The solubility,  $S$ , of zinc (concentration of zinc in a saturated aqueous solution) is given in eq. 5.

$$S = [\text{Zn}^{2+}(\text{aq})] + [\text{Zn}(\text{OH})_2(\text{aq})] + [\text{Zn}(\text{OH})_4^{2-}(\text{aq})] \quad (5)$$

**A.1** When the equilibria in eq. 1–4 are established, **calculate** the pH range in which  $[\text{Zn}(\text{OH})_2(\text{aq})]$  is the greatest among  $[\text{Zn}^{2+}(\text{aq})]$ ,  $[\text{Zn}(\text{OH})_2(\text{aq})]$  and  $[\text{Zn}(\text{OH})_4^{2-}(\text{aq})]$ . 6pt

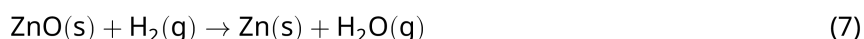
**A.2** A saturated aqueous solution of  $\text{Zn}(\text{OH})_2(\text{s})$  with  $\text{pH} = 7.00$  was prepared and filtered.  $\text{NaOH}$  was added to this filtrate to increase its  $\text{pH}$  to 12.00. **Calculate** the molar percentage of zinc that precipitates when increasing the  $\text{pH}$  from 7.00 to 12.00. Ignore the volume and temperature changes. 5pt

## Part B

Next, the recovered zinc hydroxide is heated to obtain zinc oxide according to the reaction below:

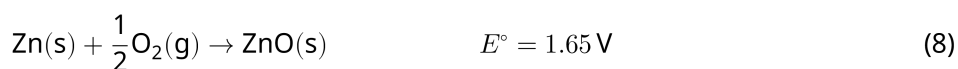


The zinc oxide is then reduced to metallic zinc by reaction with hydrogen:



**B.1** In order for reaction (7) to proceed at a hydrogen pressure kept at 1 bar, it is necessary to reduce the partial pressure of the generated water vapor. **Calculate** the upper limit for the partial pressure of water vapor to allow reaction (7) to proceed at 300 °C. 4pt  
Here, the Gibbs formation energies of zinc oxide and water vapor at 300 °C and 1 bar for all gaseous species are  $\Delta G_{\text{ZnO}}(300^\circ\text{C}) = -2.90 \times 10^2 \text{ kJ mol}^{-1}$  and  $\Delta G_{\text{H}_2\text{O}}(300^\circ\text{C}) = -2.20 \times 10^2 \text{ kJ mol}^{-1}$ , respectively.

Metallic zinc is used as a negative electrode (anode) material for metal-air batteries. The electrode consists of Zn and ZnO. It uses the following redox reaction to generate electricity with the electromotive force (e.m.f.) at 25 °C and pressure of 1 bar,  $E^\circ$ .



**B.2** A zinc-air battery was discharged at 20 mA for 24 hours. **Calculate** the change in mass of the negative electrode (anode) of the battery. 3pt



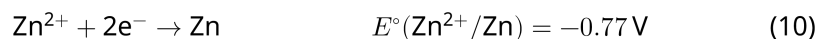
Mt. Fuji

- B.3** Consider the change of e.m.f. of a zinc-air battery depending on the environment. **Calculate** the e.m.f. at the summit of Mt. Fuji, where the temperature and altitude are  $-38\text{ }^{\circ}\text{C}$  (February) and  $3776\text{ m}$ , respectively. The atmospheric pressure is represented by 5pt

$$P [\text{bar}] = 1.013 \times \left( 1 - \frac{0.0065h}{T + 0.0065h + 273.15} \right)^{5.257} \quad (9)$$

at altitude  $h$  [m] and temperature  $T$  [ $^{\circ}\text{C}$ ]. The molar ratio of oxygen in the atmosphere is 21%. The Gibbs energy change of reaction (8) is  $\Delta G_{\text{ZnO}}(-38\text{ }^{\circ}\text{C}) = -3.26 \times 10^2 \text{ kJ mol}^{-1}$  at  $-38\text{ }^{\circ}\text{C}$  and 1 bar.

- B.4** **Calculate** the Gibbs energy change for reaction (6) at  $25\text{ }^{\circ}\text{C}$ . Note that the standard reduction potentials,  $E^{\circ}(\text{Zn}^{2+}/\text{Zn})$  and  $E^{\circ}(\text{O}_2/\text{H}_2\text{O})$  at  $25\text{ }^{\circ}\text{C}$  and 1 bar are given as (10) and (11), respectively. 9pt





ISR-2 C-4 A-1

**A4-1**  
ISR English (Israel)

## The Redox Chemistry of Zinc

### Part A

**A.1** (6 pt)

\_\_\_\_\_ < pH < \_\_\_\_\_



ISR-2 C-4 A-2

**A4-2**  
ISR English (Israel)

**A.2** (5 pt)

\_\_\_\_\_ %



ISR-2 C-4 A-3

**A4-3**  
ISR English (Israel)

**Part B**

**B.1** (4 pt)

$p_{\text{H}_2\text{O}} =$  \_\_\_\_\_ bar

**B.2** (3 pt)

\_\_\_\_\_ g





ISR-2 C-4 A-4

**A4-4**  
ISR English (Israel)

**B.3** (5 pt)

\_\_\_\_\_ v



ISR-2 C-4 A-5

**A4-5**  
ISR English (Israel)

**B.4** (9 pt)

$\Delta G^\circ =$  \_\_\_\_\_  $\text{J mol}^{-1}$

ISR-2 C-5 C-1

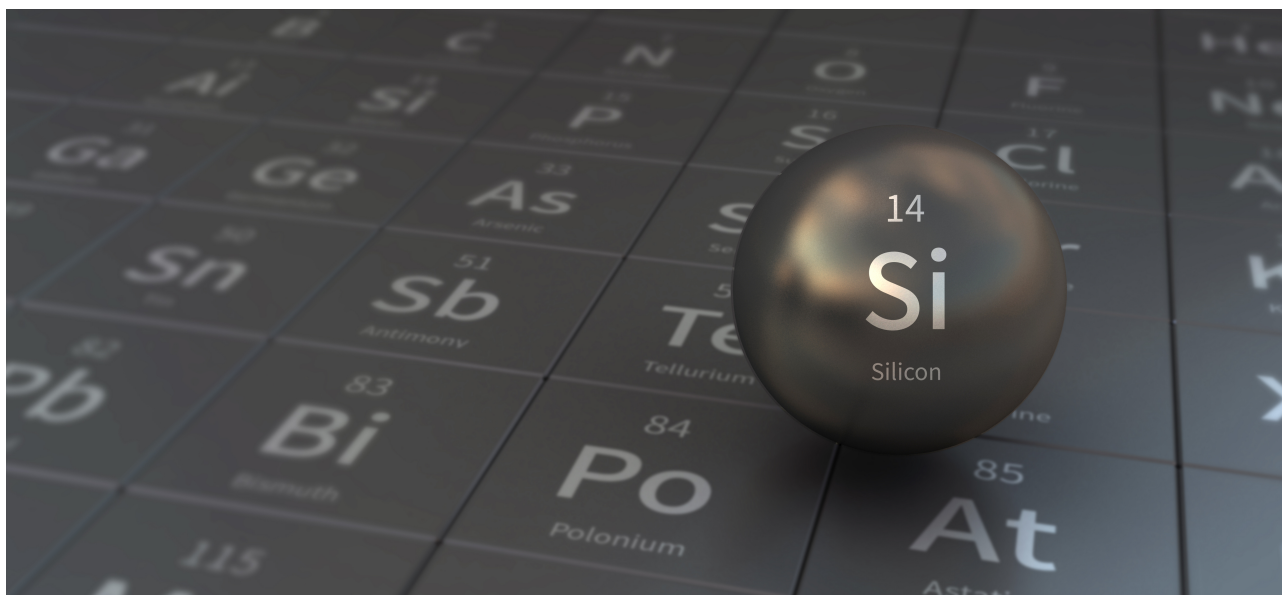
**ISR-2 C-5 C**  
Neta Eiger

**ICHO**  
**Problem 5**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Mysterious Silicon

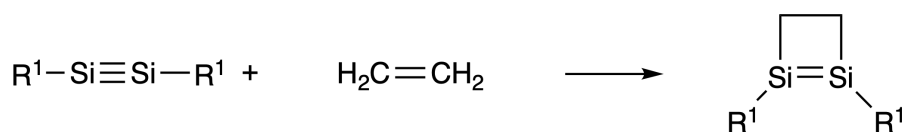
12 % of the total								
Question	A.1	A.2	A.3	A.4	B.1	B.2	B.3	Total
Points	9	7	6	10	5	15	8	<b>60</b>
Score								



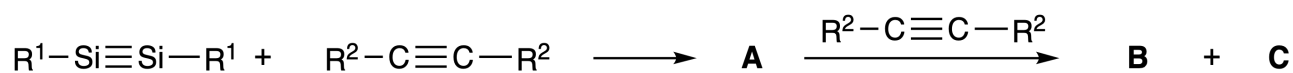
Although silicon is also a group 14 element like carbon, their properties differ significantly.

### Part A

Unlike the carbon-carbon triple bond, the silicon-silicon triple bond in a compound formulated as  $R^1-Si \equiv Si-R^1$  ( $R$ : organic substituent) is extremely reactive. For example, it reacts with ethylene to form a cyclic product that contains a four-membered ring.



When  $R^1-Si \equiv Si-R^1$  is treated with an alkyne ( $R^2-C \equiv C-R^2$ ), the four-membered-ring compound **A** is formed as an initial intermediate. Further reaction of another molecule of  $R^2-C \equiv C-R^2$  with **A** affords isomers **B** and **C**, both of which have benzene-like cyclic conjugated structures, so-called 'disilabenzenes' that contain a six-membered ring and can be formulated as  $(R^1-Si)_2(R^2-C)_4$ .



The  $^{13}\text{C}$  NMR analysis of the corresponding six-membered ring skeletons  $\text{Si}_2\text{C}_4$  shows two signals for **B** and one signal for **C**.

**A.1** **Draw** the structural formulae of **A**, **B**, and **C** using  $R^1$ ,  $R^2$ , Si, and C, with one of the possible resonance structures. 9pt

**A.2** **Calculate** the aromatic stabilization energy (ASE) for benzene and **C** (in the case of  $R^1 = R^2 = \text{H}$ ) as positive values, considering the enthalpy change in some hydrogenation reactions of unsaturated systems shown below (Fig. 1). 7pt

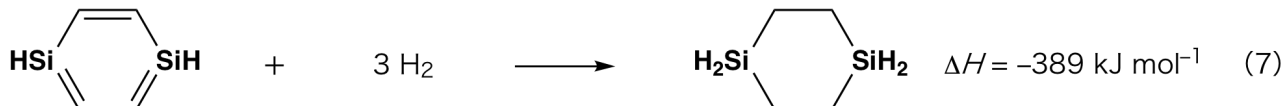
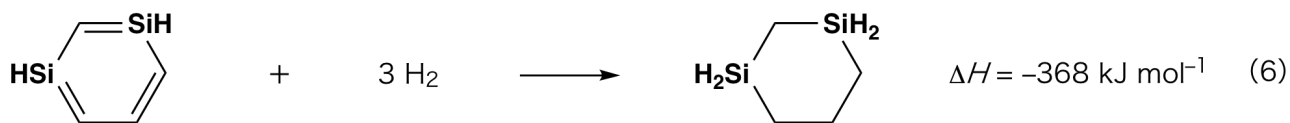
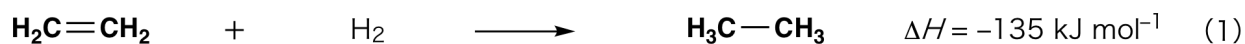


Fig. 1

When a xylene solution of **C** is heated, it undergoes isomerization to give an equilibrium mixture of compounds **D** and **E**. The molar ratio is **D** : **E** = 1 : 40.0 at 50.0 °C and **D** : **E** = 1 : 20.0 at 120.0 °C.

**A.3 Calculate**  $\Delta H$  for the transformation of **D** to **E**. Assume that  $\Delta H$  does not depend on temperature. 6pt

The isomerization from **C** to **D** and to **E** proceeds via transformations of  $\pi$ -bonds into  $\sigma$ -bonds without breaking any  $\sigma$ -bonds. A  $^{13}\text{C}$  NMR analysis revealed one signal for the  $\text{Si}_2\text{C}_4$  skeleton of **D** and two signals for that of **E**. The skeleton of **D** does not contain any three-membered rings, while **E** has two three-membered rings that share an edge.

**A.4 Draw** the structural formulae of **D** and **E** using  $\text{R}^1$ ,  $\text{R}^2$ , Si, and C. 10pt

## Part B

Silicon is able to form highly coordinated compounds (> four substituents) with electronegative elements such as fluorine. As metal fluorides are often used as fluorination reagents, highly coordinated silicon fluorides also act as fluorination reagents.

The fluorination reaction of  $\text{CCl}_4$  using  $\text{Na}_2\text{SiF}_6$  was carried out as follows.

- **Standardization of  $\text{Na}_2\text{SiF}_6$  solution :**

- Preparation

Aqueous solution **F**: 0.855 g of  $\text{Na}_2\text{SiF}_6$  ( $188.053 \text{ g mol}^{-1}$ ) dissolved in water (total volume: 200 mL).

Aqueous solution **G**: 6.86 g of  $\text{Ce}_2(\text{SO}_4)_3$  ( $568.424 \text{ g mol}^{-1}$ ) dissolved in water (total volume: 200 mL).

- Procedure

Precipitation titration of a solution **F** (50.0 mL) by dropwise adding solution **G** in the presence of xylenol orange, which coordinates to  $\text{Ce}^{3+}$ , as an indicator.

After adding 18.8 mL of solution **G**, the color of the solution changes from yellow to magenta. The generated precipitate is a binary compound that contains  $\text{Ce}^{3+}$ , and the only resulting silicon compound is  $\text{Si}(\text{OH})_4$ .

**B.1 Write** the balanced equation for the reaction of  $\text{Na}_2\text{SiF}_6$  with  $\text{Ce}_2(\text{SO}_4)_3$ . 5pt

- **Reaction of  $\text{CCl}_4$  with  $\text{Na}_2\text{SiF}_6$ :**

(Substance losses by *e.g.* evaporation are negligible during the following operations.)

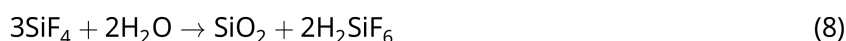
$\text{Na}_2\text{SiF}_6$  ( $x$  [g]) was added to  $\text{CCl}_4$  (500.0 g) and heated to 300 °C in a sealed pressure-resistant reaction vessel. The unreacted  $\text{Na}_2\text{SiF}_6$  and generated  $\text{NaCl}$  were removed by filtration. The filtrate was diluted to a total volume of 1.00 L with  $\text{CCl}_4$  (solution **H**).

The  $^{29}\text{Si}$  and  $^{19}\text{F}$  NMR spectra of solution **H** showed  $\text{SiF}_4$  as the only silicon compound. In the  $^{19}\text{F}$  NMR spectrum, in addition to  $\text{SiF}_4$ , signals corresponding to  $\text{CFCl}_3$ ,  $\text{CF}_2\text{Cl}_2$ ,  $\text{CF}_3\text{Cl}$ , and  $\text{CF}_4$  were observed (*cf.* Table 1). The integration ratios in the  $^{19}\text{F}$  NMR spectrum are proportional to the number of fluorine nuclei.

Table 1

<sup>19</sup> F NMR data	CFCl <sub>3</sub>	CF <sub>2</sub> Cl <sub>2</sub>	CF <sub>3</sub> Cl	CF <sub>4</sub>
Integration ratio	45.0	65.0	18.0	2.0

SiF<sub>4</sub> is hydrolyzed to form H<sub>2</sub>SiF<sub>6</sub> according to the following eq. 8:



Solution **H** (10 mL) was added to an excess amount of water, which resulted in the complete hydrolysis of SiF<sub>4</sub>. After separation, the H<sub>2</sub>SiF<sub>6</sub> generated from the hydrolysis in the aqueous solution was neutralized and completely converted to Na<sub>2</sub>SiF<sub>6</sub> (aqueous solution **J**).

The precipitate of unreacted Na<sub>2</sub>SiF<sub>6</sub> and NaCl, which was removed by filtration in the initial step (underlined), was completely dissolved in water to give an aqueous solution (solution **K**; 10.0 L).

Then, additional precipitation titrations using solution **G** were carried out, and the endpoints of the titrations with **G** were as follows:

•For solution **J** (entire amount): 61.6 mL.

•For 100 mL of solution **K**: 44.4 mL.

It should be noted here that the coexistence of NaCl or SiO<sub>2</sub> has no effect on the precipitation titration.

**B.2** Calculate the mass of the NaCl produced in the reaction vessel (information underlined), and calculate the mass ( $x$  [g]) of the Na<sub>2</sub>SiF<sub>6</sub> used as a starting material. 15pt

**B.3** 77.8% of the CCl<sub>4</sub> used as a starting material was unreacted. Calculate the mass of CF<sub>3</sub>Cl generated. 8pt



ISR-2 C-5 A-1

**A5-1**  
ISR English (Israel)

## Mysterious Silicon

### Part A

**A.1** (9 pt)

**A** (3 pt)

**B** (3 pt)

**C** (3 pt)

**A.2** (7 pt)

$C_6H_6$  : \_\_\_\_\_  $\text{kJ mol}^{-1}$ , **C** : \_\_\_\_\_  $\text{kJ mol}^{-1}$





ISR-2 C-5 A-2

**A5-2**  
ISR English (Israel)

**A.3** (6 pt)

$\Delta H =$  \_\_\_\_\_  $\text{kJ mol}^{-1}$

**A.4** (10 pt)

**D** (5 pt)

**E** (5 pt)



ISR-2 C-5 A-3

**A5-3**  
ISR English (Israel)

**Part B**

**B.1** (5 pt)

---

**B.2** (15 pt)

(Continued on the next page)



ISR-2 C-5 A-4

**A5-4**  
ISR English (Israel)

**B.2 (cont.)**

NaCl : \_\_\_\_\_ g, Na<sub>2</sub>SiF<sub>6</sub> : \_\_\_\_\_ g



ISR-2 C-5 A-5

**A5-5**  
ISR English (Israel)

**B.3** (8 pt)

$\text{CF}_3\text{Cl}$  : \_\_\_\_\_ g

ISR-2 C-6 C-1

**ISR-2 C-6 C**  
Neta Eiger

**IChO**  
**Problem 6**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## The Solid-State Chemistry of Transition Metals

13 % of the total											
Question	A.1	A.2	A.3	B.1	B.2	B.3	B.4	C.1	C.2	C.3	Total
Points	6	3	3	6	4	4	4	5	5	5	<b>45</b>
Score											



Volcano at Sakurajima island

### Part A

Japan is one of the countries with the highest number of volcanos worldwide. When silicate minerals crystallize from magma, a part of the transition-metal ions ( $M^{n+}$ ) in the magma is incorporated into the silicate minerals. The  $M^{n+}$  studied in this problem are coordinated by oxide ions ( $O^{2-}$ ) and adopt a four-coordinate tetrahedral ( $T_d$ ) geometry in the magma and six-coordinate octahedral ( $O_h$ ) geometry in the silicate minerals, both of which exhibit a high-spin electron configuration.

The distribution coefficient of  $M^{n+}$  between the silicate minerals and magma,  $D$ , can be expressed by:

$$D = \frac{[M]_s}{[M]_l}$$

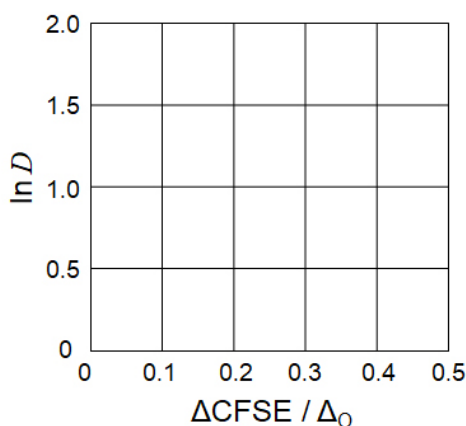
where  $[M]_s$  and  $[M]_l$  are the concentrations of  $M^{n+}$  in the silicate minerals and the magma, respectively. The table below shows the  $D$  values of  $Cr^{2+}$  and  $Mn^{2+}$  as examples.

	$Cr^{2+}$	$Mn^{2+}$
$D$	7.2	1.1

Let  $\Delta_o$  and  $CFSE^O$  be the energy separation of the d-orbitals of  $M^{n+}$  and the crystal-field stabilization energy in a  $O_h$  field, respectively. Let  $\Delta_T$  and  $CFSE^T$  be those in a  $T_d$  field.

**A.1 Calculate**  $|CFSE^O - CFSE^T| = \Delta CFSE$  in terms of  $\Delta_o$  for  $Cr^{2+}$ ,  $Mn^{2+}$ , and  $Co^{2+}$ ; assume  $\Delta_T = 4/9\Delta_o$ . 6pt

**A.2** A linear relationship is observed by plotting  $\ln D$  against  $\Delta CFSE / \Delta_o$  in the Cartesian coordinate system shown below. **Estimate**  $D$  for  $Co^{2+}$ . 3pt



Metal oxides MO (M: Ca, Ti, V, Mn, or Co) crystallize in a rock-salt structure wherein the  $M^{n+}$  adopts an  $O_h$  geometry with a high-spin electron configuration. The lattice enthalpy of these oxides is mainly governed by the Coulomb interactions based on the radius and charge of the ions and some contributions from the CFSE of  $M^{n+}$  in the  $O_h$  field.

**A.3 Choose** the appropriate set of lattice enthalpies [ $kJ mol^{-1}$ ] from one of the options (a) to (f). 3pt

	CaO	TiO	VO	MnO	CoO
(a)	3460	3878	3913	3810	3916
(b)	3460	3916	3878	3810	3913
(c)	3460	3913	3916	3810	3878
(d)	3810	3878	3913	3460	3916
(e)	3810	3916	3878	3460	3913
(f)	3810	3913	3916	3460	3878

**Part B**

A mixed oxide **A**, which contains  $\text{La}^{3+}$  and  $\text{Cu}^{2+}$ , crystallizes in a tetragonal unit cell shown in Fig.1. In the  $[\text{CuO}_6]$  octahedron, the Cu–O length along the z-axis ( $l_z$ ) is longer than that of the x-axis ( $l_x$ ), and  $[\text{CuO}_6]$  is distorted from the regular  $O_h$  geometry. This distortion removes the degeneracy of the  $e_g$  orbitals ( $d_{x^2-y^2}$  and  $d_{z^2}$ ).

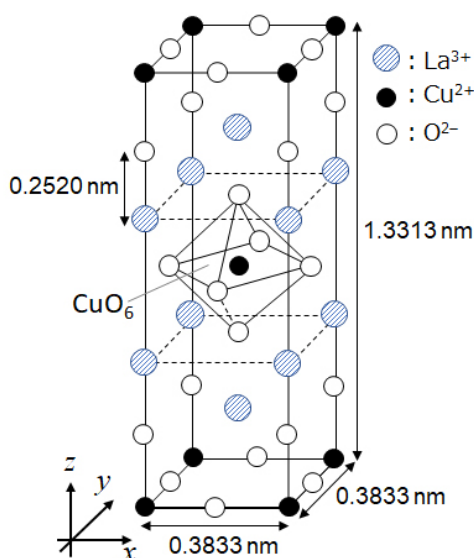


Fig. 1

**A** can be synthesized by thermal decomposition (pyrolysis) of complex **B**, which is formed by mixing metal chlorides in dilute aqueous ammonia solution containing squaric acid  $\text{C}_4\text{H}_2\text{O}_4$ , i.e., a diacid.

The pyrolysis behavior of **B** in dry air shows a weight loss of 29.1% up to 200 °C due to the loss of crystallization water, followed by another weight loss up to 700 °C due to the release of  $\text{CO}_2$ . The total weight loss during the formation of **A** from **B** is 63.6%. It should be noted that only water and  $\text{CO}_2$  are released in the pyrolysis reaction.

<b>B.1</b>	<b>Write</b> the chemical formulae for <b>A</b> and <b>B</b> .	6pt
<b>B.2</b>	<b>Calculate</b> $l_x$ and $l_z$ using Fig. 1.	4pt
<b>B.3</b>	For $\text{Cu}^{2+}$ in the distorted $[\text{CuO}_6]$ octahedron in <b>A</b> of Fig. 1, <b>write</b> the names of the split $e_g$ orbitals ( $d_{x^2-y^2}$ and $d_{z^2}$ ) in (i) and (ii), and <b>draw</b> the electron configuration in the dotted box in your answer sheet.	4pt



**A** is an insulator. When one  $\text{La}^{3+}$  is substituted with one  $\text{Sr}^{2+}$ , one hole is generated in the crystal lattice that can now conduct electricity. As a result, the  $\text{Sr}^{2+}$ -doped **A** shows superconductivity below 38 K.

When a substitution reaction took place for **A**,  $2.05 \times 10^{27}$  holes  $\cdot m^{-3}$  were generated.

**B.4** **Calculate** the percentage of  $\text{Sr}^{2+}$  substituted for  $\text{La}^{3+}$  based on the mole ratio 4pt  
in the substitution reaction. Note that the valences of the constituent ions and the crystal structure are not altered by the substitution reaction.

### Part C

$\text{Cu}_2(\text{CH}_3\text{CO}_2)_4$  is composed of four  $\text{CH}_3\text{CO}_2^-$  coordinated to two  $\text{Cu}^{2+}$  (Fig. 2A).  $\text{Cu}_2(\text{CH}_3\text{CO}_2)_4$  exhibits high levels of structural symmetry, with two axes passing through the carbon atoms of the four  $\text{CH}_3\text{CO}_2^-$  and an axis passing through the two  $\text{Cu}^{2+}$ , all of which are oriented orthogonal relative to each other.

When a dicarboxylate ligand is used instead of  $\text{CH}_3\text{CO}_2^-$ , a “cage complex” is formed. The cage complex  $\text{Cu}_4(\mathbf{L1})_4$  (Fig. 2C) is composed of planar dicarboxylate **L1** (Fig. 2B) and  $\text{Cu}^{2+}$ . The angle  $\theta$  between the coordination directions of the two carboxylates, indicated by the arrows in Fig. 2B, determines the structure of the cage complex. The  $\theta$  is  $0^\circ$  for **L1**. Note that hydrogen atoms are not shown in Fig. 2.

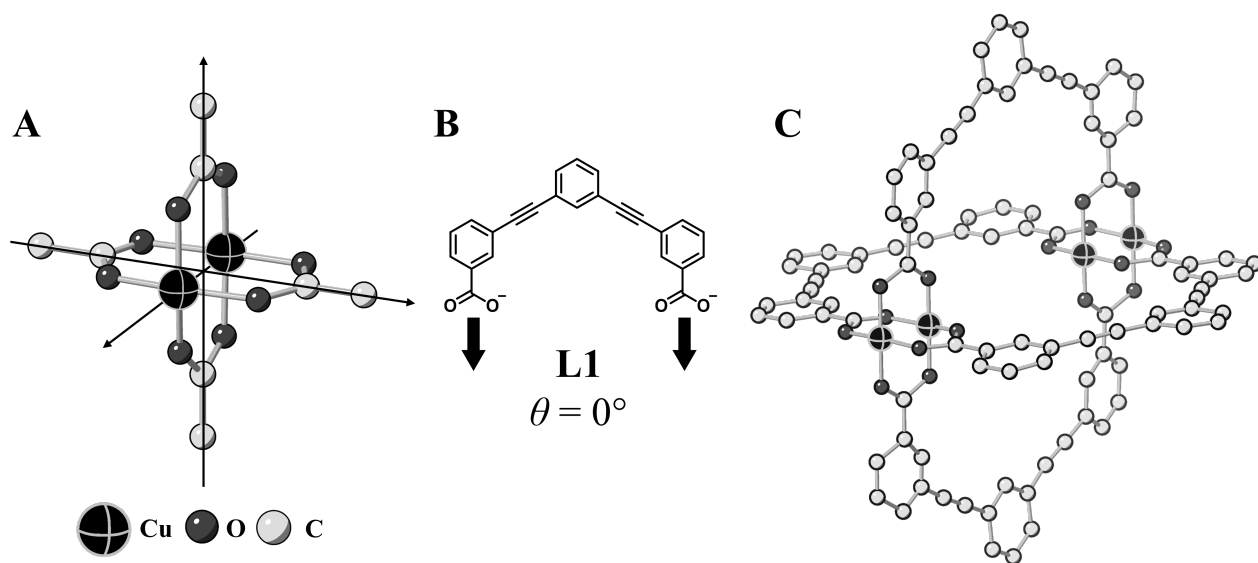
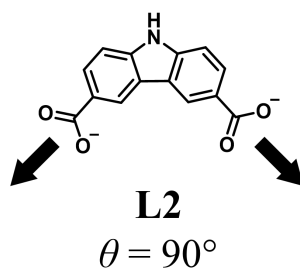


Fig. 2

- C.1** The  $\theta$  of the planar dicarboxylate **L2** below is fixed to  $90^\circ$ . If the composition of the cage complex formed from **L2** and  $\text{Cu}^{2+}$  is  $\text{Cu}_n(\text{L2})_m$ , give the smallest integer combination of  $n$  and  $m$ . 5pt  
 Assume that only the  $\text{CO}_2^-$  groups of **L2** form a coordination bond to  $\text{Cu}^{2+}$  ions.



A zinc complex,  $Zn_4O(CH_3CO_2)_6$ , contains four tetrahedral  $Zn^{2+}$ , six  $CH_3CO_2^-$ , and one  $O^{2-}$  (Fig. 3A). In  $Zn_4O(CH_3CO_2)_6$ , the  $O^{2-}$  is located at the origin, and the three axes passing through the carbon atoms of  $CH_3CO_2^-$  are oriented orthogonal relative to each other.

When *p*-benzenedicarboxylate (Fig. 3B, **L3**,  $\theta = 180^\circ$ ) is used instead of  $CH_3CO_2^-$ , the  $Zn^{2+}$  clusters are linked to each other to form a crystalline solid (**X**) that is called a "porous coordination polymer" (Fig. 3C). The composition of **X** is  $[Zn_4O(L3)_3]_n$ , and it has a cubic crystal structure with nano-sized pores (Hebrew: *h'orim*).

One pore is represented as a sphere in Fig. 3D, and each tetrahedral  $Zn^{2+}$  cluster is represented as a dark gray polyhedron in Fig. 3C and 3D. Note that hydrogen atoms are not shown in Fig. 3.

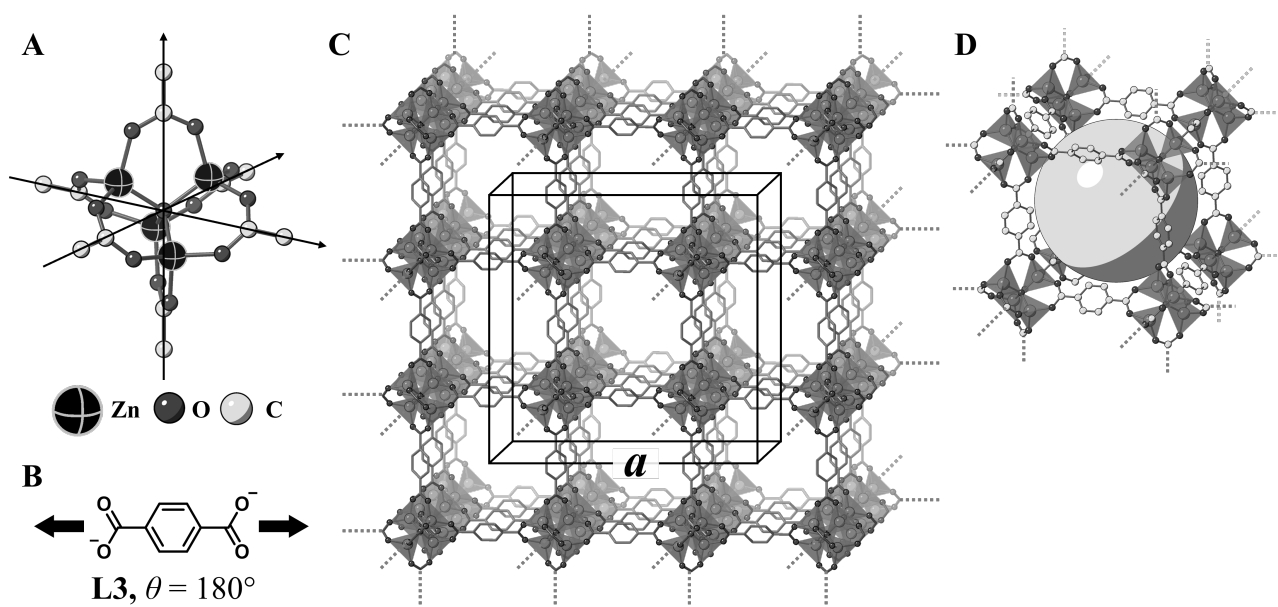


Fig. 3

**C.2** **X** has a cubic unit cell with a side length of  $a$  (Fig. 3C) and a density of  $0.592 \text{ g cm}^{-3}$ . **Calculate**  $a$  in [cm]. 5pt

**C.3** **X** contains a considerable number of pores, and 1 g of **X** can accommodate  $3.0 \times 10^2 \text{ mL}$  of  $CO_2$  gas in the pores at 1 bar and  $25^\circ C$ . **Calculate** the average number of  $CO_2$  molecules per pore. 5pt



ISR-2 C-6 A-1

**A6-1**  
ISR English (Israel)

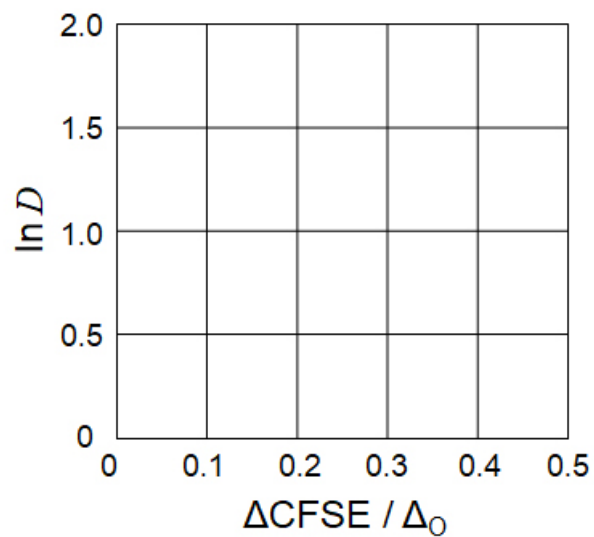
## The Solid-State Chemistry of Transition Metals

### Part A

**A.1** (6 pt)

$\text{Cr}^{2+}$  : \_\_\_\_\_  $\Delta_0$ ,  $\text{Mn}^{2+}$  : \_\_\_\_\_  $\Delta_0$ ,  $\text{Co}^{2+}$  : \_\_\_\_\_  $\Delta_0$

**A.2** (3 pt)



$D$ : \_\_\_\_\_

**A.3** (3 pt)

\_\_\_\_\_



ISR-2 C-6 A-3

**A6-3**  
ISR English (Israel)

**Part B**

**B.1** (6 pt)

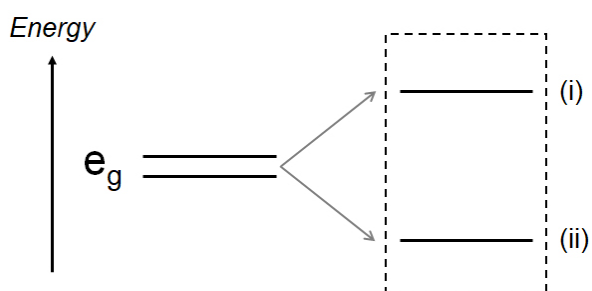
A : \_\_\_\_\_, B : \_\_\_\_\_

**B.2** (4 pt)

$l_x =$  \_\_\_\_\_ nm,  $l_z =$  \_\_\_\_\_ nm

**B.3** (4 pt)

(i) : \_\_\_\_\_ , (ii) : \_\_\_\_\_



**B.4** (4 pt)

\_\_\_\_\_ %



ISR-2 C-6 A-5

**A6-5**  
ISR English (Israel)

**Part C**

**C.1** (5 pt)

$n =$  \_\_\_\_\_,  $m =$  \_\_\_\_\_

**C.2** (5 pt)

$a =$  \_\_\_\_\_ cm





ISR-2 C-6 A-6

**A6-6**  
ISR English (Israel)

**C.3** (5 pt)

---

ISR-2 C-7 C-1

**ISR-2 C-7 C**  
Neta Eiger

**ICHO**  
**Problem 7**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Playing with Non-benzenoid Aromaticity

13 % of the total					
Question	A.1	A.2	A.3	B.1	Total
Points	5	2	19	10	<b>36</b>
Score					

Prof. Nozoe (1902–1996) opened the research field of non-benzenoid aromatic compounds, which are now ubiquitous in organic chemistry.



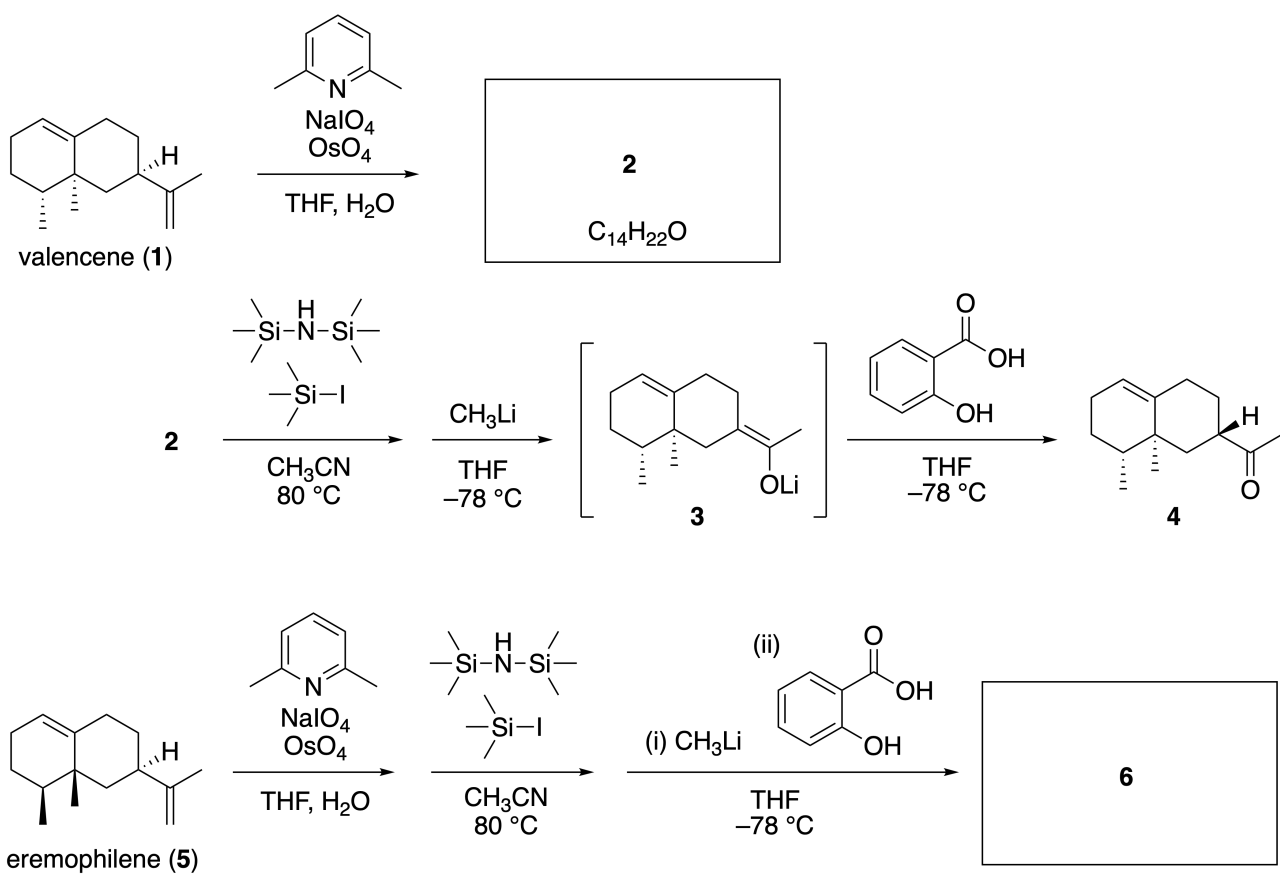
Photo courtesy: Tohoku Univ.

### Part A

Lineariifolianone is a natural product with a unique structure, which was isolated from *Inula linariifolia*. From valencene (**1**), a one-step conversion yields **2**, before a three-step conversion via **3** yields ketone **4**. Eremophilene (**5**) is converted into **6** by performing the same four-step conversion.



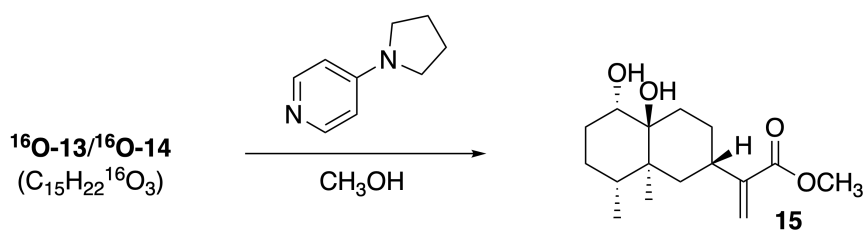
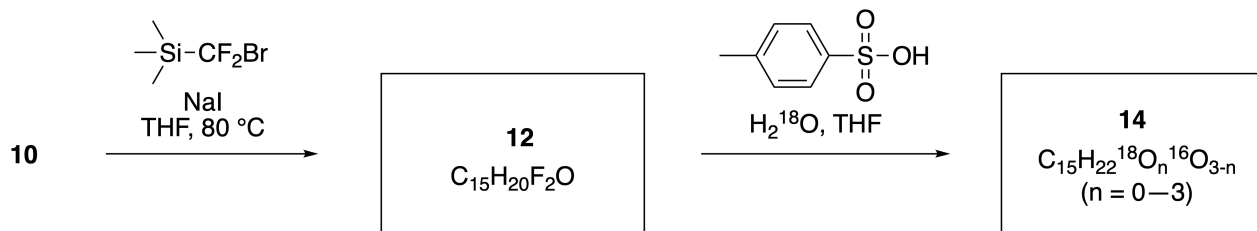
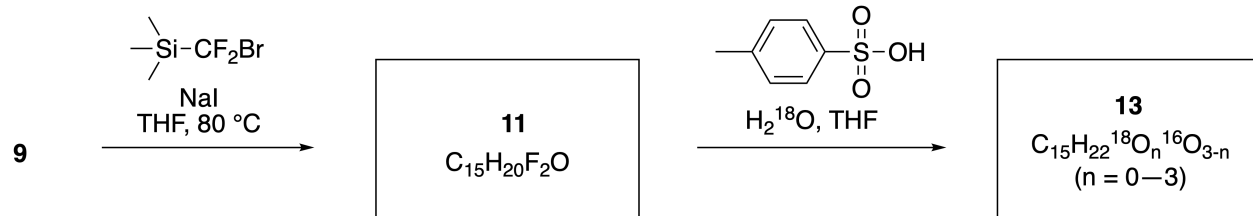
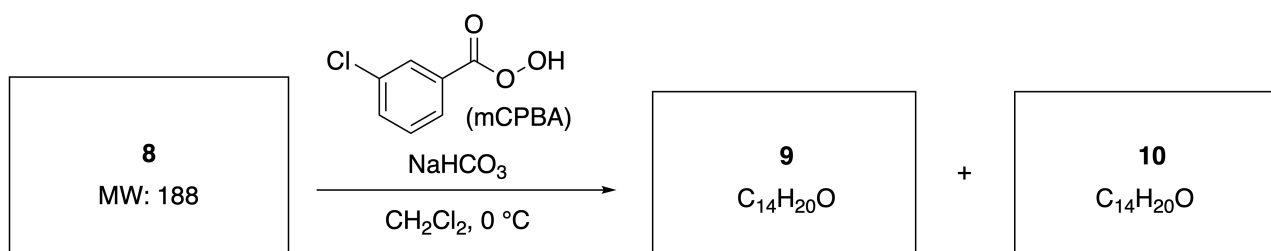
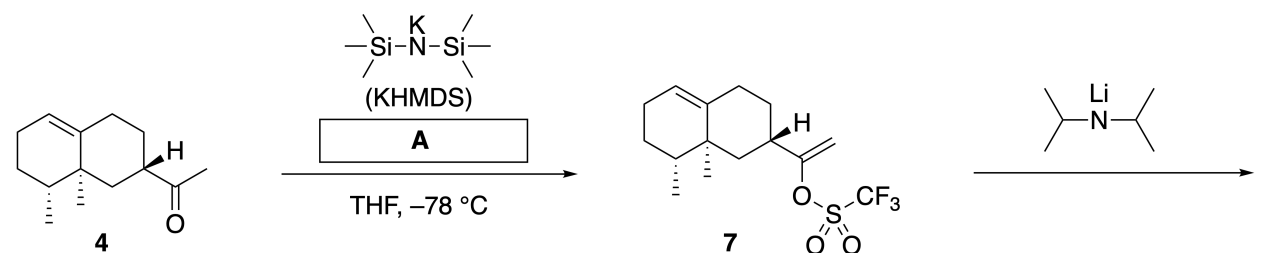
*Inula linariifolia*



**A.1** Draw the structures of **2** and **6** and clearly identify the stereochemistry where necessary. 5pt

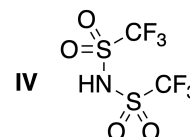
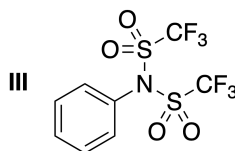
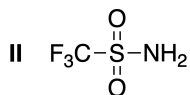
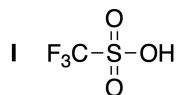
Then, ketone **4** is converted into ester **15**. Compound **8** (molecular weight: 188) retains all the stereocenters in **7**. Compounds **9** and **10** have five stereocenters and no carbon-carbon double bonds.

Assume that  $\text{H}_2^{18}\text{O}$  is used instead of  $\text{H}_2^{16}\text{O}$  for the synthesis of  $^{18}\text{O}$ -labelled-linearifolianones **13** and **14** from **11** and **12**, respectively. Compounds **13** and **14** are  $^{18}\text{O}$ -labelled isotopomers. Ignoring isotopic labelling, both **13** and **14** provide the same product **15** with identical stereochemistry.



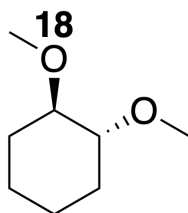
**A.2** Choose the appropriate structure for **A**.

2pt



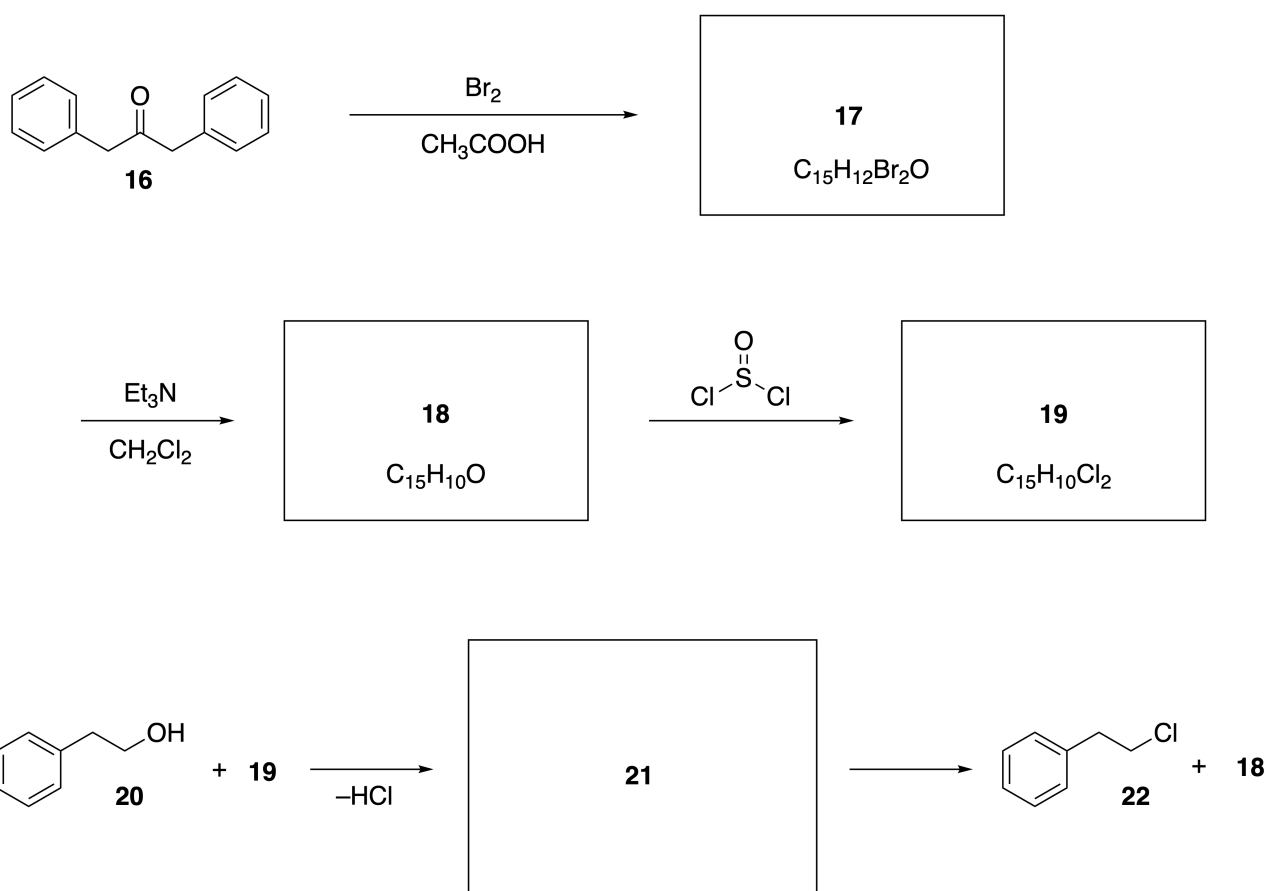
**A.3** Draw the structures of **8-14** and clearly **identify** the stereochemistry where necessary. Also, **indicate** the introduced  $^{18}\text{O}$  atoms for **13** and **14** as shown in the example below.

19pt



**Part B**

Compound **19** is synthesized as shown below. In relation to non-benzenoid aromaticity, **19** can be used as an activator for alcohols, and **20** was converted to **22** via ion-pair intermediate **21**. Although the formation of **21** was observed by NMR, **21** gradually decomposes to give **18** and **22**.



$^1\text{H}$  NMR ( $\text{CD}_3\text{CN}$ , ppm) **20**:  $\delta$  7.4–7.2 (5H), 3.7 (2H), 2.8 (2H), 2.2 (1H)

**21**:  $\delta$  8.5–7.3 (15H), 5.5 (2H), 3.4 (2H)

**B.1** **Draw** the structures of **17–19** and **21**. Identifying the stereochemistry is not necessary. 10pt



ISR-2 C-7 A-1

**A7-1**  
ISR English (Israel)

## Playing with Non-benzenoid Aromaticity

### Part A

**A.1** (5 pt)

**2** (2 pt)

**6** (3 pt)

**A.2** (2 pt)





**A.3** (19 pt)

**8** (3 pt)



**9** (2 pt)

**10** (2 pt)



**11** (2 pt)

**12** (2 pt)



**13** (4 pt)

**14** (4 pt)



ISR-2 C-7 A-3

**A7-3**  
ISR English (Israel)

**Part B**

**B.1** (10 pt)

**17** (2 pt)

**18** (2 pt)

**19** (3 pt)

**21** (3 pt)

ISR-2 C-8 C-1

**ISR-2 C-8 C**  
Neta Eiger

**ICHO**  
**Problem 8**  
**Cover sheet**

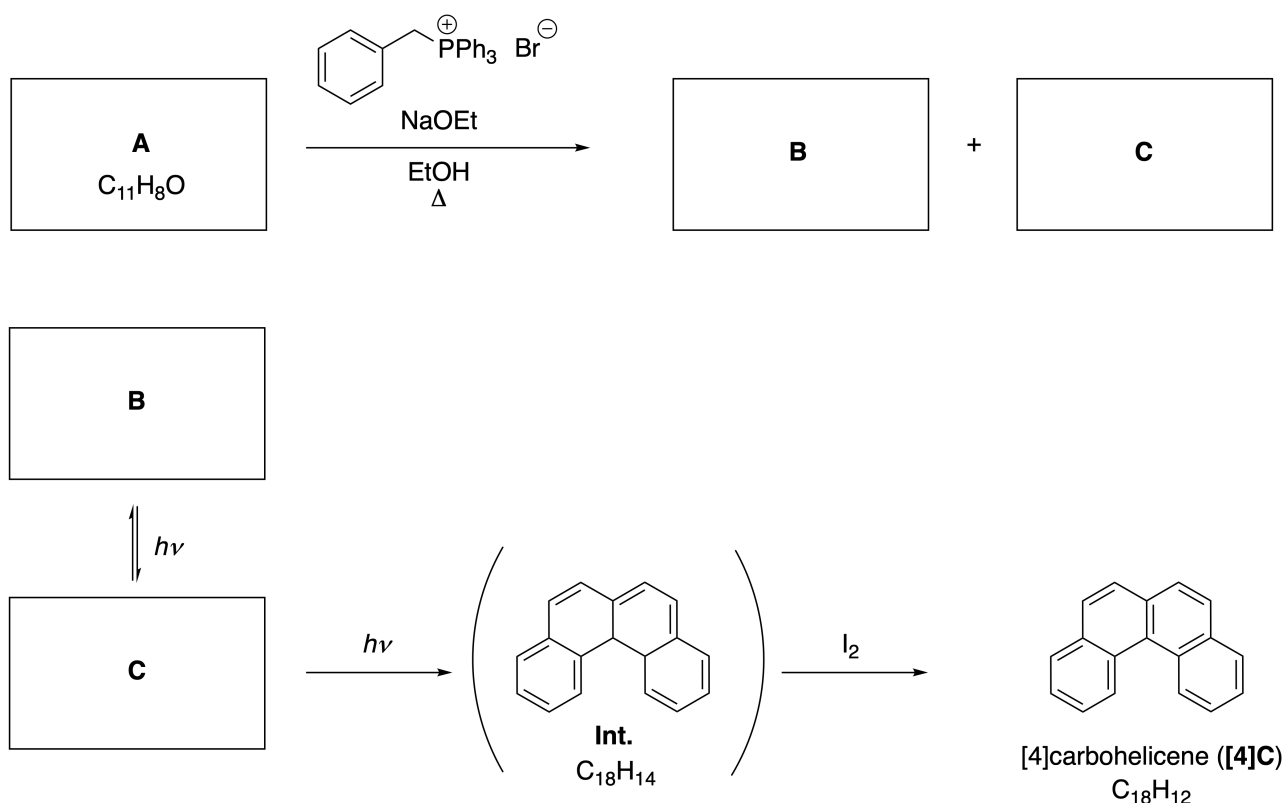
Please return this cover sheet together with all the related question sheets.

## Dynamic Organic Molecules and Their Chirality

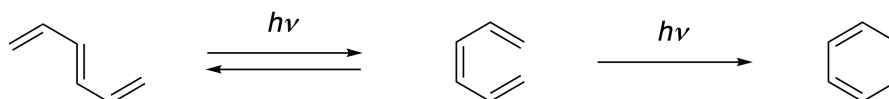
11 % of the total						
Question	A.1	A.2	A.3	B.1	B.2	Total
Points	9	3	7	3	4	<b>26</b>
Score						

### Part A

Polycyclic aromatic hydrocarbons with successive ortho-connections are called [n]carbohelicenes (here, n represents the number of six-membered rings) (see below). [4]Carbohelicene (**[4]C**) is efficiently prepared by a route using a photoreaction as shown below, via an intermediate (**Int.**) that is readily oxidized by iodine.



The photoreaction proceeds in a manner similar to the following example.

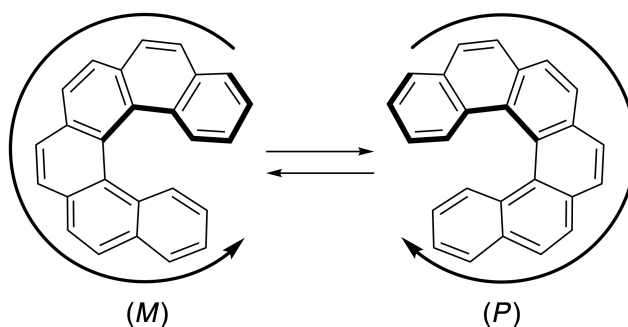


**Note:** For all of Question 8, please draw alternating single and double bonds in your answers to the problems as depicted in the examples of carbohelicene. Do not use circles for conjugated  $\pi$  systems.

**A.1** Draw the structures of **A-C**. Stereoisomers should be distinguished. 9pt

**A.2** Attempts to synthesize [5]carbohelicene from the same phosphonium salt and an appropriate starting compound resulted in the formation of only a trace amount of [5]carbohelicene, instead affording product **D** whose molecular weight was 2 Da lower than that of [5]carbohelicene. The  $^1\text{H}$  NMR chemical shifts of **D** are listed below. Draw the structure of **D**. [**D** ( $\delta$ , ppm in  $\text{CS}_2$ , r.t.), 8.85 (2H), 8.23 (2H), 8.07 (2H), 8.01 (2H), 7.97 (2H), 7.91 (2H)] 3pt

[5]- and larger [n]carbohelicenes have helical chirality, and interconversion between enantiomers of these helicenes is significantly slow at room temperature. The chirality of [n]carbohelicenes is defined as (*M*) or (*P*) as shown below.



[n]Carbohelicenes with n larger than 4 can be enantiomerically separated by a chiral column chromatography, which was developed by Prof. Yoshio Okamoto.

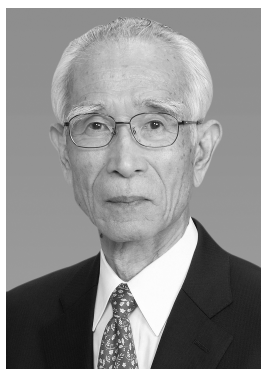
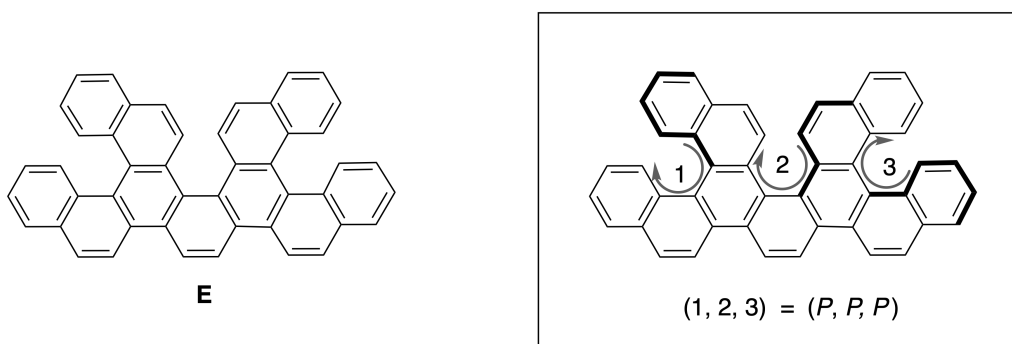
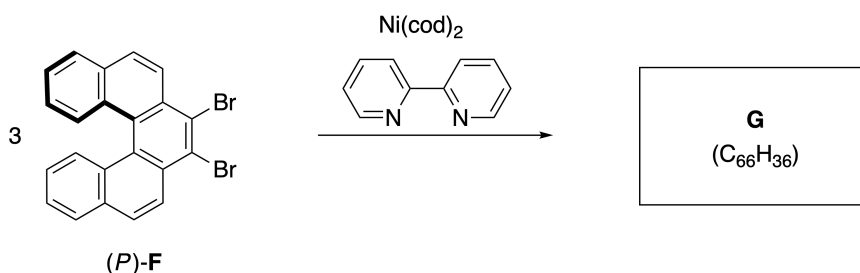
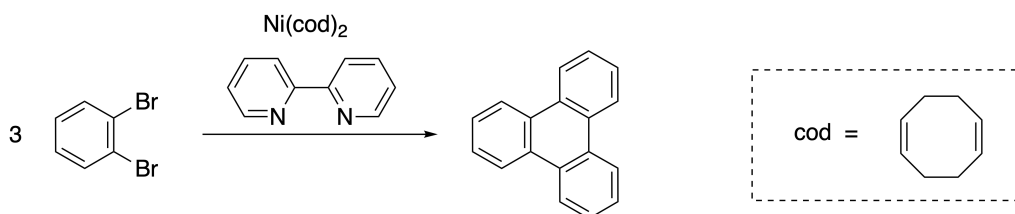


Photo courtesy: The Japan Prize Foundation

Multiple helicenes are molecules that contain two or more helicene-like structures. If its helical chirality is considered, several stereoisomers exist in a multiple helicene. For example, compound **E** contains three [5]carbohelicene-like moieties in one molecule. One of the stereoisomers is described as (*P, P, P*) as shown below.

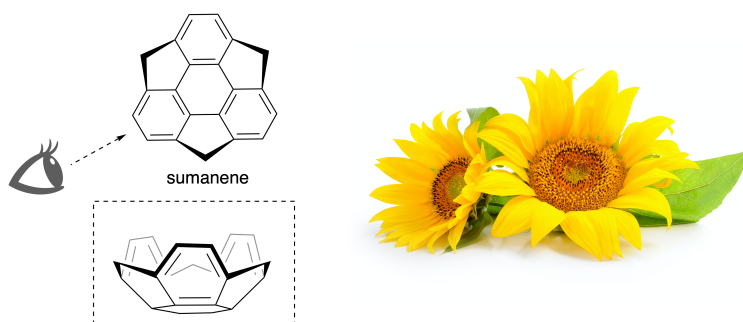


- A.3** The nickel-mediated trimerization of 1,2-dibromobenzene generates triphenylene. 7pt  
When the same reaction is applied to an enantiomer of **F**, (*P*)-**F**, multiple helicene **G** ( $C_{66}H_{36}$ ) is obtained. Given that interconversion between stereoisomers does not occur during the reaction, **identify all** the possible stereoisomers of **G** formed in this process, without duplication.  
As a reference, one isomer should be drawn completely with the chirality defined as in the example above, with numerical labels; the other stereoisomers should be listed with location numbers and *M* and *P* labels according to the same numbering. For instance, the other stereoisomers of **E** should be listed as (1, 2, 3) = (*P, M, P*), (*P, M, M*), (*P, P, M*), (*M, M, M*), (*M, M, P*), (*M, P, P*), and (*M, P, M*).

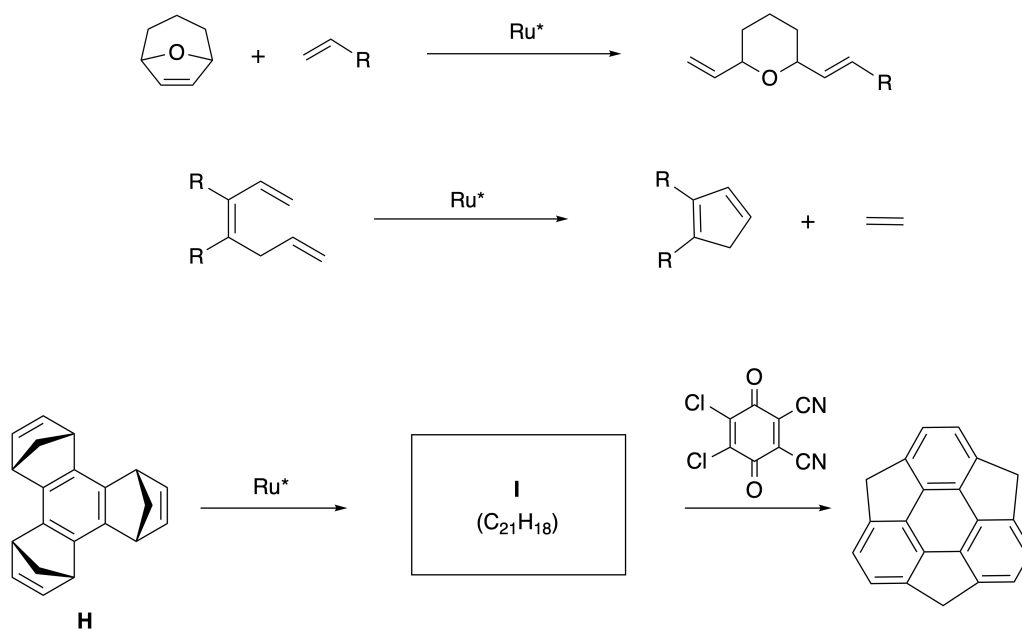


**Part B**

Sumanene is a bowl-shaped hydrocarbon that was first reported in Japan in 2003. The name "sumanene" derives from a Sanskrit-Hindi word "suman" that means sunflower. The synthesis of sumanene was achieved by a reaction sequence that consists of a ring-opening and a ring-closing metathesis.

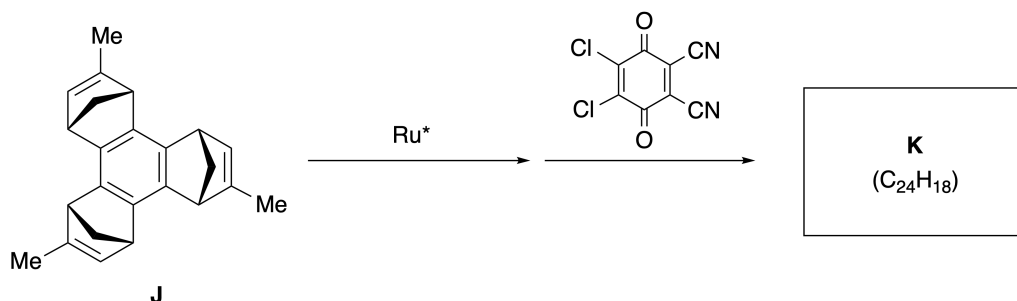


Representative metathesis reactions catalyzed by a ruthenium catalyst ( $\text{Ru}^*$ ) are shown below.



**B.1** Draw the structure of intermediate **I** (its stereochemistry is not required).

3pt



- B.2** Starting from the optically active precursor **J**, the same reaction sequence gives the optically active sumanene derivative **K**. The stereocenters in **J** suffer no inversion during the metathesis reaction. **Draw** the structure of **K** with the appropriate stereochemistry. 4pt





ISR-2 C-8 A-1

**A8-1**  
ISR English (Israel)

## Dynamic Organic Molecules and Their Chirality

### Part A

**A.1** (9 pt)

**A** (3 pt)

**B** (3 pt)

**C** (3 pt)

**A.2** (3 pt)



ISR-2 C-8 A-2

**A8-2**  
ISR English (Israel)

**A.3** (7 pt)



ISR-2 C-8 A-3

**A8-3**  
ISR English (Israel)

**Part B**

**B.1** (3 pt)

**B.2** (4 pt)

ISR-2 C-9 C-1

**ISR-2 C-9 C**  
Neta Eiger

**ICHO**  
**Problem 9**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

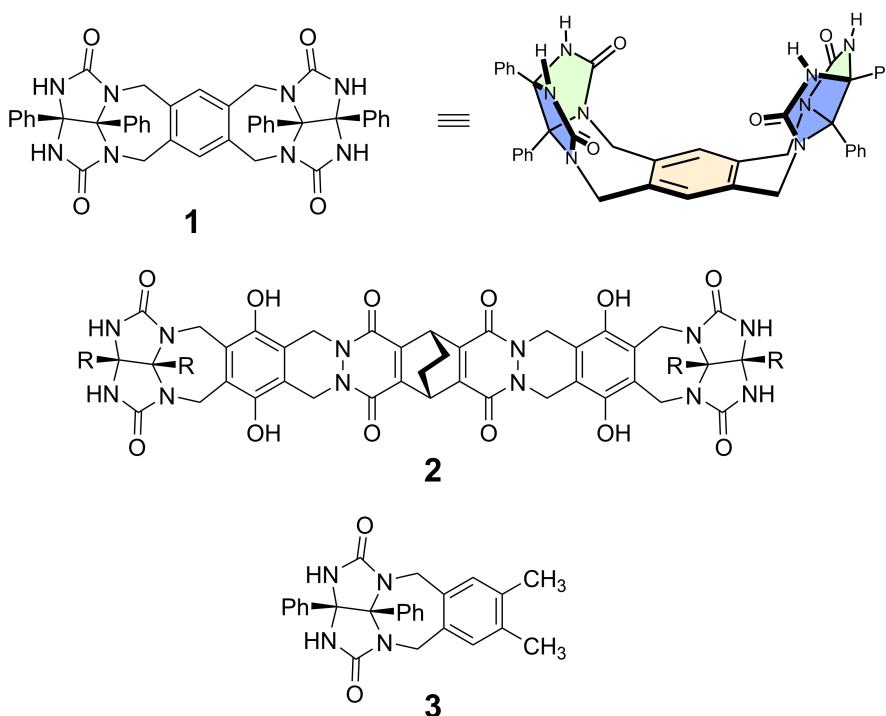
## Likes and Dislikes of Capsule

10 % of the total						
Question	A.1	A.2	A.3	A.4	A.5	Total
Points	13	2	2	3	3	<b>23</b>
Score						

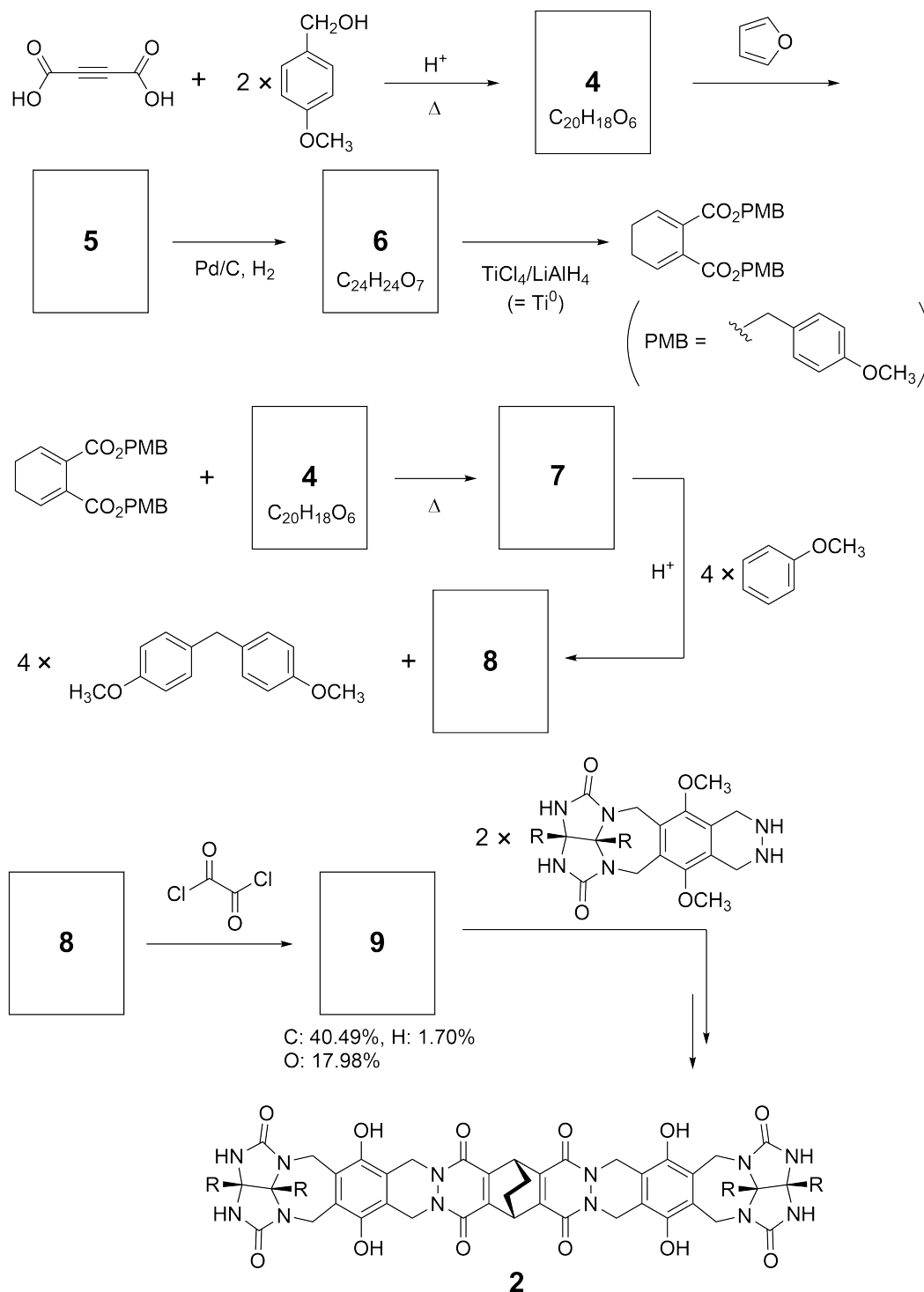
Good kids don't do this, but if you unseam a tennis ball, you can disassemble it into two U-shaped pieces.



Based on this idea, compounds **1** and **2** were synthesized as U-shaped molecules with different sizes. Compound **3** was prepared as a comparison of **1** and the encapsulation behavior of these compounds was investigated.



The synthetic route to **2** is shown below. The elemental composition of compound **9**: C; 40.49%, H; 1.70%, and O; 17.98% by mass.





ISR-2 C-9 Q-3

# Q9-3

ISR English (Israel)

**A.1** **Draw** the structures of **4-9**; the stereochemistry can be neglected. Use "PMB" as a substituent instead of drawing the whole structure of *p*-methoxybenzyl group shown in the scheme above. 13pt

In the mass spectrum of **1**, the ion peak corresponding to its dimer (**1<sub>2</sub>**) was clearly observed, whereas an ion peak for **3<sub>2</sub>** was not observed in the spectrum of **3**. In the <sup>1</sup>H NMR spectra of a solution of **1<sub>2</sub>**, all the NH protons derived from **1** were observed to be chemically equivalent, and their chemical shift was significantly different from that of the NH protons of **3**.

These data indicate that hydrogen bonds are formed between the NH moieties of **1** and atoms **X** of another molecule of **1** to form the dimeric capsule.

**A.2** **Circle** all the appropriate atom(s) **X** in **1**. 2pt

**A.3** **Give** the number of the hydrogen bonds in the dimeric capsule (**1<sub>2</sub>**). 2pt

The dimeric capsule of **1** ( $1_2$ ) has an internal space wherein an appropriate small molecule Z can be encapsulated. This phenomenon is expressed by the following equation:



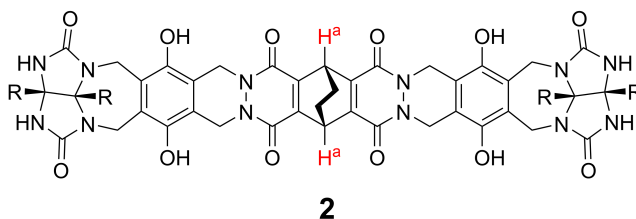
The equilibrium constant of the encapsulation of Z into  $1_2$  is given as below:

$$K_a = \frac{[Z@1_2]}{[Z][1_2]} \quad (2)$$

Encapsulation of a molecule into a capsule could be monitored by NMR spectroscopy. For example,  $1_2$  in  $C_6D_6$  gave different signals in the  $^1H$  NMR spectra before and after addition of  $CH_4$ .

Compound **2** also forms a rigid and larger dimeric capsule ( $2_2$ ). The  $^1H$  NMR spectrum of  $2_2$  was measured in  $C_6D_6$ ,  $C_6D_5F$ , and a  $C_6D_6/C_6D_5F$  solvent mixture, with all other conditions being kept constant.

The chemical shifts for the  $H^a$  proton of **2** in the above solvents are summarized below, and no other signals from the  $H^a$  in **2**, except for the listed, were observed. Assume that the interior of the capsule is always filled with the largest possible number of solvent molecules and that each signal corresponds to one species of the filled capsule.



solvent	$\delta$ (ppm) of $H^a$
$C_6D_6$	4.60
$C_6D_5F$	4.71
$C_6D_6 / C_6D_5F$	4.60, 4.71, 4.82

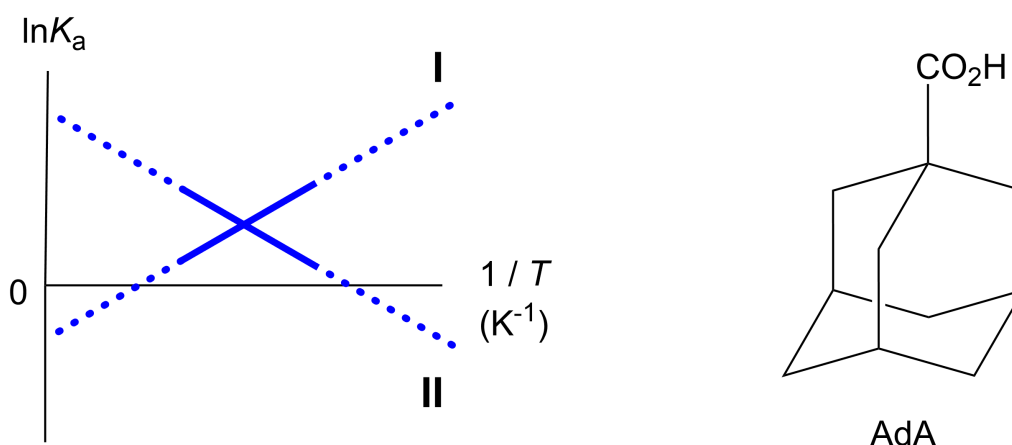
**A.4 Determine** the number of  $C_6D_6$  and  $C_6D_5F$  molecules encapsulated in  $2_2$ , for each  $H^a$  signal. 3pt



$^1\text{H}$  NMR measurements in  $\text{C}_6\text{D}_6$  revealed that  $\mathbf{2}_2$  can incorporate one molecule of 1-adamantanecarboxylic acid (AdA), and the association constants ( $K_a$ ) which are expressed below were determined for various temperatures.  $[\text{solvent@}\mathbf{2}_2]$  denotes a species containing one or more solvent molecules.

$$K_a = \frac{[\text{Z@}\mathbf{2}_2]}{[\text{Z}][\text{solvent@}\mathbf{2}_2]} \quad (3)$$

Similarly, the  $K_a$  values of  $\text{CH}_4$  and  $\mathbf{1}_2$  given as eq (2) at various temperatures in  $\text{C}_6\text{D}_6$  were also determined by  $^1\text{H}$  NMR measurements. The plots of the two association constants (as  $\ln K_a$  vs  $1/T$ ) are shown below.



No  $\text{C}_6\text{D}_6$  molecule is encapsulated in  $\mathbf{1}_2$ .

**A.5** **Choose** the correct option (A or B) for each of the gaps (1)–(5) in the following paragraph: 3pt

In line **II** (in the graph above), the entropy change ( $\Delta S$ ) is ( 1 ) and enthalpy change ( $\Delta H$ ) is ( 2 ), indicating that the driving force for the encapsulation in line **II** is ( 3 ). Therefore, line **I** corresponds to ( 4 ), and line **II** corresponds to ( 5 ).

	A	B
(1)	positive	negative
(2)	positive	negative
(3)	$\Delta S$	$\Delta H$
(4)	$\mathbf{1}_2$ and $\text{CH}_4$	$\mathbf{2}_2$ and AdA
(5)	$\mathbf{1}_2$ and $\text{CH}_4$	$\mathbf{2}_2$ and AdA



ISR-2 C-9 A-1

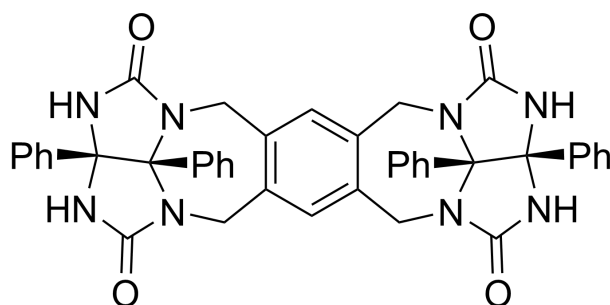
**A9-1**  
ISR English (Israel)

## Likes and Dislikes of Capsule

**A.1** (13 pt)

<b>4</b> (2 pt)	<b>5</b> (3 pt)
<b>6</b> (2 pt)	<b>7</b> (2 pt)
<b>8</b> (2 pt)	<b>9</b> (2 pt)

**A.2** (2 pt)



**A.3** (2 pt)

**A.4** (3 pt)

$\delta$ (ppm) of H <sup>a</sup>	number of C <sub>6</sub> D <sub>6</sub> molecules	number of C <sub>6</sub> D <sub>5</sub> F molecules
4.60 ppm		
4.71 ppm		
4.82 ppm		

**A.5** (3 pt)

(1) : \_\_\_\_\_ (2) : \_\_\_\_\_ (3) : \_\_\_\_\_

(4) : \_\_\_\_\_ (5) : \_\_\_\_\_