

GBR-1 C-0 C-1

**GBR-1 C-0 C**  
Alexander Thow

**IChO**  
**General instructions**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.



GBR-1 C-0 G-1

# G0-1

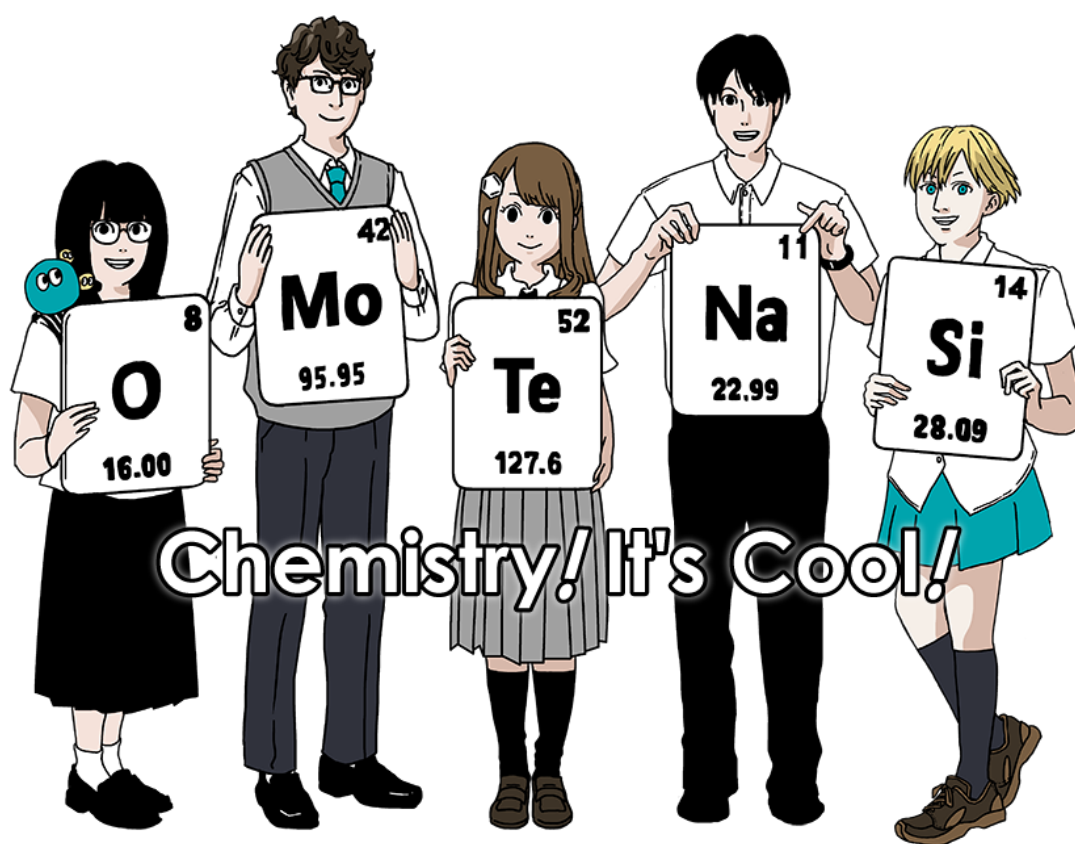
British English (United Kingdom)

International Chemistry Olympiad 2021 Japan

53rd IChO2021 Japan

25th July - 2nd August, 2021

<https://www.icho2021.org>





GBR-1 C-0 G-2

# GO-2

British English (United Kingdom)

## General Instructions

- You are only allowed to use pen to write the answers.
- Your calculator must be non-programmable.
- This examination has **9 problems**.
- You can solve the problems in any order.
- You will have **5 hours** to solve all problems.
- You can **begin** working only after the **START** command is given.
- All results must be written in the appropriate answer boxes with pen on the **answer sheets**. Use the back of the question sheets if you need rough paper. Remember that answers written outside the answer boxes will not be graded.
- Write relevant calculations in the appropriate boxes when necessary. Full marks will be given for correct answers only when your work is shown.
- The invigilator will announce a **30-minute** warning before the **STOP** command.
- You **must stop** working when the **STOP** command is given. Failure to stop writing will lead to the nullification of your examination.
- The official English version of this examination is available on request only for clarification.
- You are not allowed to leave your working place without permission. If you need any assistance (broken calculator, need to go to the toilet, etc.), raise your hand and wait until an invigilator arrives.

**GOOD LUCK!**

## Problems and Grading Information

	<b>Title</b>	<b>Total Score</b>	<b>Percentage</b>
<b>1</b>	Hydrogen at a Metal Surface	<b>24</b>	<b>11</b>
<b>2</b>	Isotope Time Capsule	<b>35</b>	<b>11</b>
<b>3</b>	Lambert–Beer Law?	<b>22</b>	<b>8</b>
<b>4</b>	The Redox Chemistry of Zinc	<b>32</b>	<b>11</b>
<b>5</b>	Mysterious Silicon	<b>60</b>	<b>12</b>
<b>6</b>	The Solid-State Chemistry of Transition Metals	<b>45</b>	<b>13</b>
<b>7</b>	Playing with Non-benzenoid Aromaticity	<b>36</b>	<b>13</b>
<b>8</b>	Dynamic Organic Molecules and Their Chirality	<b>26</b>	<b>11</b>
<b>9</b>	Likes and Dislikes of Capsules	<b>23</b>	<b>10</b>
		<b>Total</b>	<b>100</b>



## Physical Constants and Equations

### Constants

Speed of light in vacuum	$c = 2.99792458 \times 10^8 \text{ m s}^{-1}$
Planck constant	$h = 6.62607015 \times 10^{-34} \text{ J s}$
Elementary charge	$e = 1.602176634 \times 10^{-19} \text{ C}$
Electron mass	$m_e = 9.10938370 \times 10^{-31} \text{ kg}$
Electric constant (permittivity of vacuum)	$\epsilon_0 = 8.85418781 \times 10^{-12} \text{ F m}^{-1}$
Avogadro constant	$N_A = 6.02214076 \times 10^{23} \text{ mol}^{-1}$
Boltzmann constant	$k_B = 1.380649 \times 10^{-23} \text{ J K}^{-1}$
Faraday constant	$F = N_A \times e = 9.64853321 \times 10^4 \text{ C mol}^{-1}$
Gas constant	$R = N_A \times k_B = 8.3144626 \text{ J K}^{-1} \text{ mol}^{-1}$ $= 8.2057366 \times 10^{-2} \text{ L atm K}^{-1} \text{ mol}^{-1}$
Unified atomic mass unit	$u = 1 \text{ Da} = 1.66053907 \times 10^{-27} \text{ kg}$
Standard pressure	$p = 1 \text{ bar} = 10^5 \text{ Pa}$
Atmospheric pressure	$p_{\text{atm}} = 1.01325 \times 10^5 \text{ Pa}$
Zero degrees Celsius	$0^\circ \text{C} = 273.15 \text{ K}$
Ångstrom	$1 \text{ Å} = 10^{-10} \text{ m}$
Picometer	$1 \text{ pm} = 10^{-12} \text{ m}$
Electronvolt	$1 \text{ eV} = 1.602176634 \times 10^{-19} \text{ J}$
Part-per-million	$1 \text{ ppm} = 10^{-6}$
Part-per-billion	$1 \text{ ppb} = 10^{-9}$
Part-per-trillion	$1 \text{ ppt} = 10^{-12}$
pi	$\pi = 3.141592653589793$
The base of the natural logarithm (Euler's number)	$e = 2.718281828459045$

## Equations

The ideal gas law	$PV = nRT$ , where $P$ is the pressure, $V$ is the volume, $n$ is the amount of substance, $T$ is the absolute temperature of ideal gas.
Coulomb's law	$F = k_e \frac{q_1 q_2}{r^2}$ , where $F$ is the electrostatic force, $k_e (\approx 9.0 \times 10^9 \text{ N m}^2 \text{ C}^{-2})$ is Coulomb's constant, $q_1$ and $q_2$ are the magnitudes of the charges, and $r$ is the distance between the charges.
The first law of thermodynamics	$\Delta U = q + w$ , where $\Delta U$ is the change in the internal energy, $q$ is the heat supplied, $w$ is the work done.
Enthalpy $H$	$H = U + PV$
Entropy based on Boltzmann's principle $S$	$S = k_B \ln W$ , where $W$ is the number of microstates.
The change of entropy $\Delta S$	$\Delta S = \frac{q_{\text{rev}}}{T}$ , where $q_{\text{rev}}$ is the heat for the reversible process.
Gibbs free energy $G$	$G = H - TS$ $\Delta_r G^\circ = -RT \ln K = -zFE^\circ$ , where $K$ is the equilibrium constant, $z$ is the number of electrons, $E^\circ$ is the standard electrode potential.
Reaction quotient $Q$	$\Delta_r G = \Delta_r G^\circ + RT \ln Q$ For a reaction $aA + bB \rightleftharpoons cC + dD$ $Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ , where $[A]$ is the concentration of A.

Heat change $\Delta q$	$\Delta q = nc_m \Delta T$ , where $c_m$ is the temperature-independent molar heat capacity.
Nernst equation for a redox reaction	$E = E^\circ + \frac{RT}{zF} \ln \frac{C_{\text{ox}}}{C_{\text{red}}}$ , where $C_{\text{ox}}$ is the concentration of the oxidized substance, $C_{\text{red}}$ is the concentration of the reduced substance.
Arrhenius equation	$k = A \exp\left(-\frac{E_a}{RT}\right)$ , where $k$ is the rate constant, $A$ is the pre-exponential factor, $E_a$ is the activation energy. $\exp(x) = e^x$
Lambert-Beer equation	$A = \varepsilon lc$ , where $A$ is the absorbance, $\varepsilon$ is the molar absorption coefficient, $l$ is the optical path length, and $c$ is the concentration of the solution.
Henderson-Hasselbalch equation	For an equilibrium $\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-$ , where equilibrium constant is $K_a$ , $\text{pH} = \text{p}K_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$
Energy of a photon	$E = h\nu = h\frac{c}{\lambda}$ , where $\nu$ is the frequency, $\lambda$ is the wavelength of the light.
The sum of a geometric series	When $x \neq 1$ , $1 + x + x^2 + \dots + x^n = \sum_{i=0}^n x^i = \frac{1 - x^{n+1}}{1 - x}$
Approximation equation that can be used to solve problems	When $x \ll 1$ , $\frac{1}{1 - x} \simeq 1 + x$



GBR-1 C-0 G-6

# G0-6

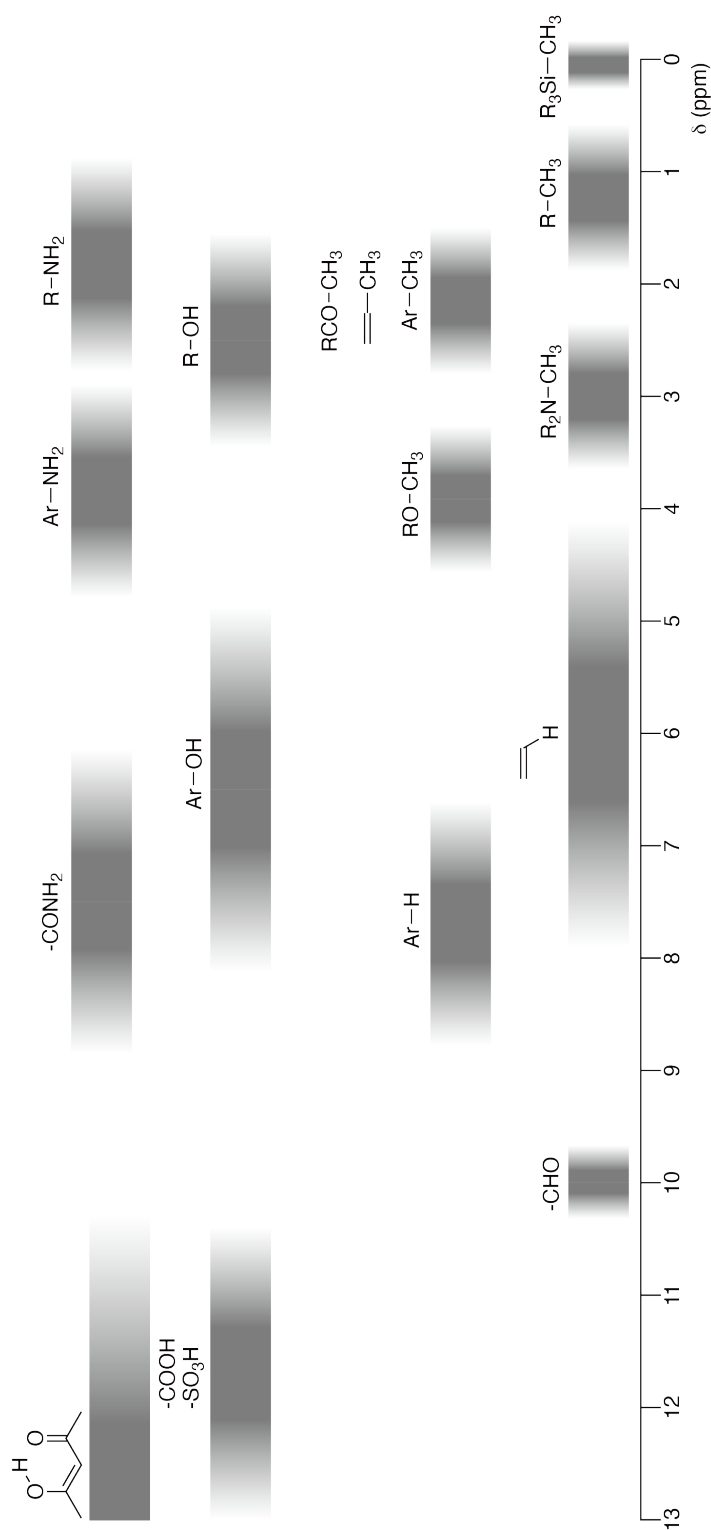
British English (United Kingdom)

## The Periodic Table

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H Hydrogen 1.008																	2 He Helium 4.003
3 Li Lithium 6.968	4 Be Beryllium 9.012															9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.306															17 Cl Chlorine 35.452	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.630	33 As Arsenic 74.922	34 Se Selenium 78.971	35 Br Bromine 79.904	36 Kr Krypton 83.798
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium [99]	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.414	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.760	52 Te Tellurium 127.60	53 I Iodine 126.904	54 Xe Xenon 131.293
55 Cs Caesium 132.905	56 Ba Barium 137.327	57-71 La-Lu Lanthanoids	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.084	79 Au Gold 196.967	80 Hg Mercury 200.592	81 Tl Thallium 204.384	82 Pb Lead 207.2	83 Bi Bismuth 208.98	84 Po Polonium [210]	85 At Astatine [210]	86 Rn Radon [222]
87 Fr Francium [223]	88 Ra Radium [226]	89-103 Ac-Lr Actinoids	104 Rf Rutherfordium [267]	105 Db Dubnium [268]	106 Sg Seaborgium [271]	107 Bh Bohrium [272]	108 Hs Hassium [277]	109 Mt Meitnerium [276]	110 Ds Darmstadtium [281]	111 Rg Roentgenium [280]	112 Cn Copernicium [285]	113 Nh Nihonium [278]	114 Fl Flerovium [289]	115 Mc Moscovium [289]	116 Lv Livermorium [293]	117 Ts Tennessine [293]	118 Og Oganesson [294]
		57-71 La-Lu Lanthanoids	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.242	61 Pm Promethium [145]	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.500	67 Ho Holmium 164.930	68 Er Erbium 167.259	69 Tm Thulium 168.934	70 Yb Ytterbium 173.045	71 Lu Lutetium 174.967	
		89-103 Ac-Lr Actinoids	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium [237]	94 Pu Plutonium [239]	95 Am Americium [243]	96 Cm Curium [247]	97 Bk Berkelium [247]	98 Cf Californium [252]	99 Es Einsteinium [252]	100 Fm Fermium [257]	101 Md Mendelevium [258]	102 No Nobelium [259]	103 Lr Lawrencium [262]	

Key:  
 113  
 Nh  
 Nihonium  
 [278]  
 atomic number  
 Symbol  
 name  
 atomic weight [in parenthesis for the radioactive element]

**<sup>1</sup>H NMR Chemical Shifts**



$\Delta\delta$  for one alkyl group-substitution: ca. +0.4 ppm



GBR-1 C-1 C-1

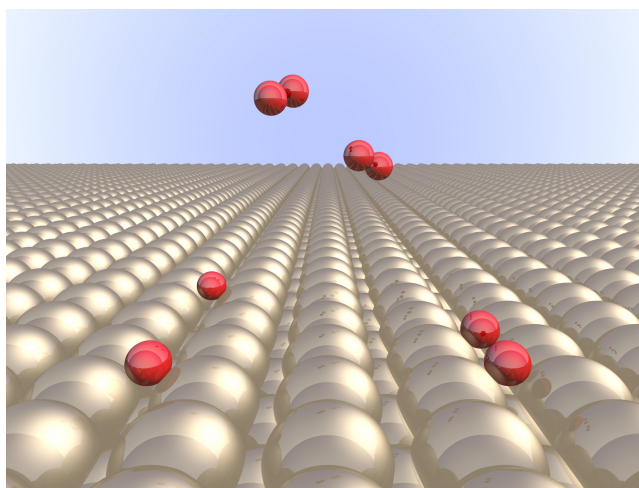
**GBR-1 C-1 C**  
Alexander Thow

**IChO**  
**Problem 1**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Hydrogen at a Metal Surface

11 % of the total							
Question	A.1	A.2	B.1	B.2	B.3	B.4	Total
Points	6	4	5	3	3	3	<b>24</b>
Score							



Hydrogen is anticipated to be a future energy source that does not depend on fossil fuels. Here, we will consider the hydrogen storage process in a metal, which is related to hydrogen transportation and storage technology.

### Part A

As hydrogen is absorbed into the bulk of a metal via its surface, let us first consider the adsorption process of hydrogen at the metal surface,  $\text{H}_2(\text{g}) \rightarrow 2\text{H}(\text{ad})$ , where the gaseous and adsorbed states of hydrogen are represented as (g) and (ad), respectively.

Hydrogen molecules ( $\text{H}_2$ ) that reach the metal surface (M) dissociate at the surface and are adsorbed as H atoms (Fig. 1). Here, the potential energy of  $\text{H}_2$  is represented by two variables: the interatomic distance,  $d$ , and the height relative to the surface metal atom,  $z$ . It is assumed that the axis along the two H atoms is parallel to the surface and that the centre of gravity is always on the vertical dotted line in Fig. 1.

Fig. 2 shows the potential energy contour plot for the dissociation at the surface. The numerical values represent the potential energy in units of  $\text{kJ mol}^{-1}$ . The solid line spacing is  $20 \text{ kJ mol}^{-1}$ , the dashed line spacing is  $100 \text{ kJ mol}^{-1}$ , and the spacing between solid and dashed lines is  $80 \text{ kJ mol}^{-1}$ . The zero-point vibration energy is ignored.

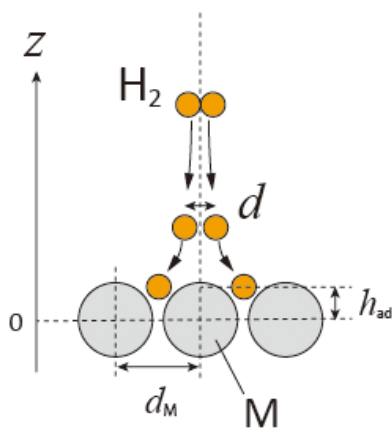


Fig.1 Definition of variables. Drawing is not to scale.

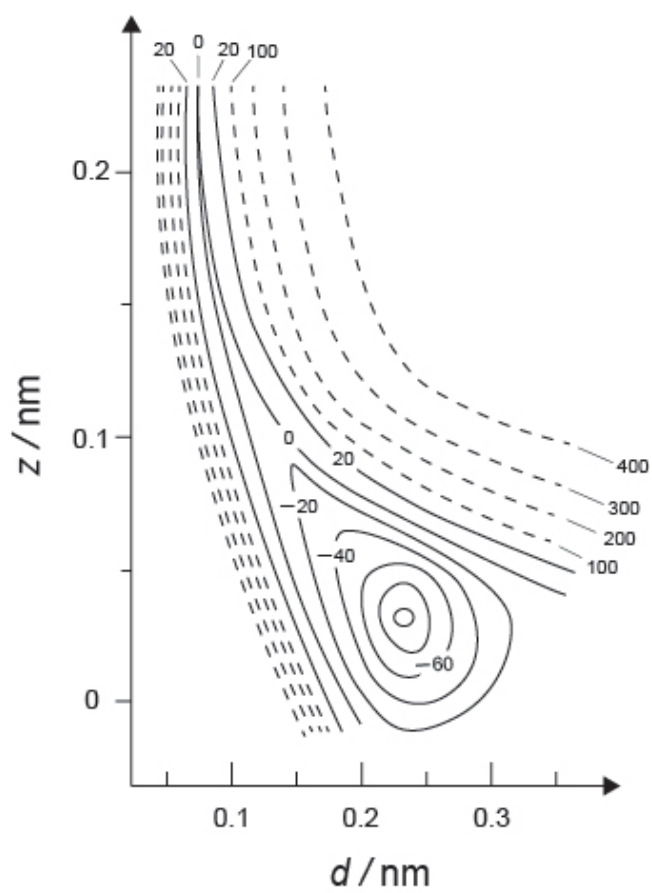
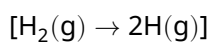


Fig.2

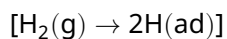
- A.1** For each of the following items (i)–(iii), **select** the closest value from A–G. 6pt
- (i) The interatomic distance for a gaseous  $\text{H}_2$  molecule
  - (ii) The interatomic distance between metal atoms ( $d_M$  in Fig. 1)
  - (iii) The distance of adsorbed H atoms from the surface ( $h_{\text{ad}}$  in Fig. 1)

A. 0.03 nm	B. 0.07 nm	C. 0.11 nm	D. 0.15 nm
E. 0.19 nm	F. 0.23 nm	G. 0.27 nm	

- A.2** For each of the following items (i)–(ii), **select** the closest value from A–H. 4pt
- (i) the energy required for the dissociation of gaseous  $\text{H}_2$  to gaseous H



- (ii) the energy released during the adsorption of gaseous  $\text{H}_2$



A. 20 $\text{kJ mol}^{-1}$	B. 40 $\text{kJ mol}^{-1}$	C. 60 $\text{kJ mol}^{-1}$	D. 100 $\text{kJ mol}^{-1}$
E. 150 $\text{kJ mol}^{-1}$	F. 200 $\text{kJ mol}^{-1}$	G. 300 $\text{kJ mol}^{-1}$	H. 400 $\text{kJ mol}^{-1}$

## Part B

The adsorbed hydrogen atoms are then either absorbed into the bulk, or recombine and desorb back into the gas phase, as shown in the reactions (1a) and (1b). H(ab) represents a hydrogen atom absorbed in the bulk.



The reaction rates per surface site for adsorption, desorption, and absorption are  $r_1[\text{s}^{-1}]$ ,  $r_2[\text{s}^{-1}]$  and  $r_3[\text{s}^{-1}]$ , respectively. They are expressed as:

$$r_1 = k_1 P_{\text{H}_2} (1 - \theta)^2 \quad (2)$$

$$r_2 = k_2 \theta^2 \quad (3)$$

$$r_3 = k_3 \theta \quad (4)$$

where  $k_1 [\text{s}^{-1} \text{Pa}^{-1}]$ ,  $k_2 [\text{s}^{-1}]$  and  $k_3 [\text{s}^{-1}]$  are the reaction rate constants and  $P_{\text{H}_2}$  is the pressure of  $\text{H}_2$ .

Among the sites available on the surface,  $\theta$  ( $0 \leq \theta \leq 1$ ) is the fraction occupied by H atoms. It is assumed that adsorption and desorption are fast compared to absorption ( $r_1, r_2 \gg r_3$ ) and that  $\theta$  remains constant.

**B.1**  $r_3$  can be expressed as:

5pt

$$r_3 = \frac{k_3}{1 + \sqrt{\frac{1}{P_{\text{H}_2} C}}} \quad (5)$$

**Express**  $C$  using  $k_1$  and  $k_2$ .

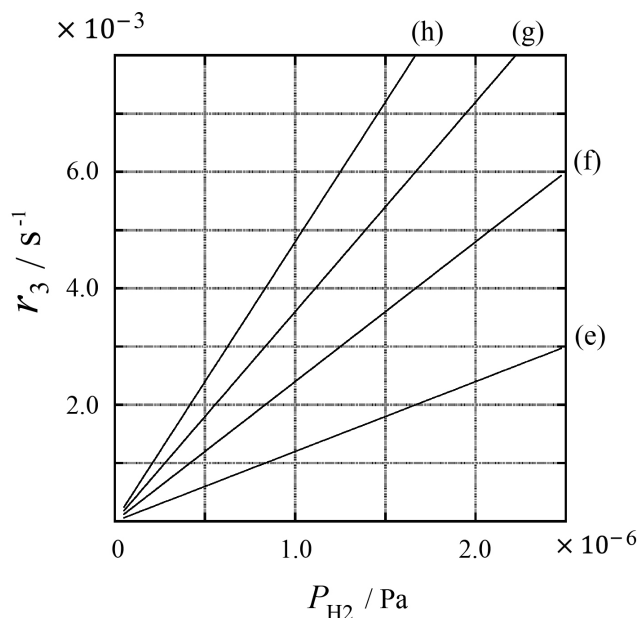
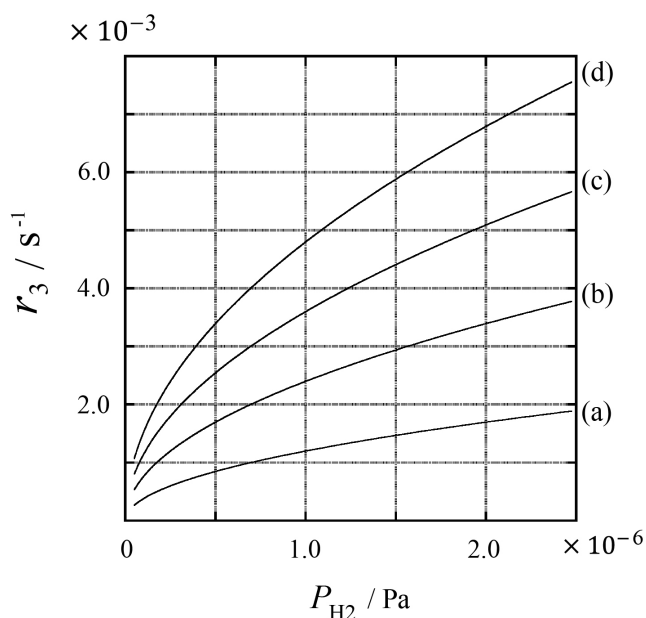
A metal sample with a surface area of  $S = 1.0 \times 10^{-3} \text{ m}^2$  was placed in a 1L container ( $1\text{L} = 1.0 \times 10^{-3} \text{ m}^3$ ) with  $\text{H}_2$  at pressure  $P_{\text{H}_2} = 1.0 \times 10^2 \text{ Pa}$ . The density of hydrogen-atom adsorption sites on the surface was  $N = 1.3 \times 10^{18} \text{ m}^{-2}$ . The surface temperature was kept at  $T = 400 \text{ K}$ .

As the reaction (1) proceeded,  $P_{\text{H}_2}$  decreased at a constant rate of  $v = 4.0 \times 10^{-4} \text{ Pa s}^{-1}$ . Assume that  $\text{H}_2$  is an ideal gas and that the volume of the metal sample is negligible.

**B.2 Calculate** in moles the amount of H atoms absorbed per unit area of the surface per unit time,  $A$  [ $\text{mol s}^{-1} \text{ m}^{-2}$ ]. 3pt

**B.3 Calculate** the value of  $k_3$  at 400 K, where  $C = 1.0 \times 10^2 \text{ Pa}^{-1}$ . 3pt  
If you did not obtain the answer to **B.2**, use  $A = 3.6 \times 10^{-7} \text{ mol s}^{-1} \text{ m}^{-2}$ .

**B.4** At a different  $T$ ,  $C = 2.5 \times 10^3 \text{ Pa}^{-1}$  and  $k_3 = 4.8 \times 10^{-2} \text{ s}^{-1}$ . 3pt  
For  $r_3$  as a function of  $P_{\text{H}_2}$  at this temperature, **select** the correct plot from (a)-(h).





GBR-1 C-1 A-1

**A1-1**  
British English (United Kingdom)

## Hydrogen at a Metal Surface

### Part A

**A.1** (6 pt)

(i)	(ii)	(iii)

**A.2** (4 pt)

(i)	(ii)



GBR-1 C-1 A-2

**A1-2**  
British English (United Kingdom)

**Part B**

**B.1** (5 pt)

$C =$  \_\_\_\_\_

**B.2** (3 pt)

$A =$  \_\_\_\_\_  $\text{mol s}^{-1} \text{m}^{-2}$





GBR-1 C-1 A-3

# A1-3

British English (United Kingdom)

**B.3** (3 pt)

$$k_3 = \underline{\hspace{10em}} \text{ s}^{-1}$$

**B.4** (3 pt)

\_\_\_\_\_

GBR-1 C-2 C-1

**GBR-1 C-2 C**  
Alexander Thow

**ICHO**  
**Problem 2**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Isotope Time Capsule

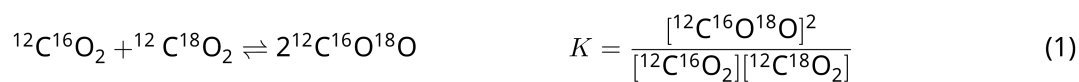
11 % of the total					
Question	A.1	A.2	A.3	A.4	Total
Points	8	8	10	9	<b>35</b>
Score					



Molecules that differ only in isotopic composition, such as  $\text{CH}_4$  and  $\text{CH}_3\text{D}$ , are called isotopologues. Isotopologues are widely considered to have the same chemical characteristics. In fact, however, there is a slight difference.

Assume that all of the substances in this Question are in the gas phase.

Consider the following equilibrium:



The entropy,  $S$ , increases with increasing the number of possible microscopic states of a system,  $W$ :

$$S = k_B \ln W \quad (2)$$

$W = 1$  for  $^{12}\text{C}^{16}\text{O}_2$  and  $^{12}\text{C}^{18}\text{O}_2$ .

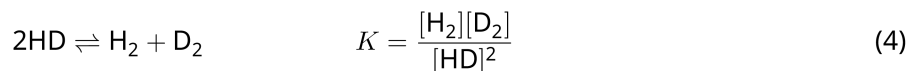
In contrast,  $W = 2$  for a  $^{12}\text{C}^{16}\text{O}^{18}\text{O}$  molecule because the oxygen atoms in this molecule are distinguishable. As the right-hand side of the equilibrium shown in eq. 1 has two  $^{12}\text{C}^{16}\text{O}^{18}\text{O}$  molecules,  $W = 2^2 = 4$ .

**A.1** The enthalpy change,  $\Delta H$ , of eq. 3 is positive regardless of the temperature. 8pt



**Calculate** the equilibrium constant,  $K$ , for eq. 3 at very low temperatures (as  $T \rightarrow 0$ ) and at very high temperatures (as  $T \rightarrow +\infty$ ). Assume that the reaction remains unchanged at these temperatures and that  $\Delta H$  converges to a constant value for high temperatures.

The  $\Delta H$  of the following process can be explained by considering molecular vibrations.



At  $T = 0$  K, the vibrational energy of a diatomic molecule whose vibration frequency is  $\nu$  [ $\text{s}^{-1}$ ] is expressed as:

$$E = \frac{1}{2}h\nu \quad (5)$$

$$\nu = \frac{1}{2\pi} \sqrt{\frac{k}{\mu}} \quad (6)$$

, where  $k$  is the force constant and  $\mu$  the reduced mass, which is expressed in terms of the mass of the two atoms in the diatomic molecule,  $m_1$  and  $m_2$ , according to:

$$\mu = \frac{m_1 m_2}{m_1 + m_2} \quad (7)$$

**A.2** The vibration of  $\text{H}_2$  is at  $4161.0 \text{ cm}^{-1}$ . 8pt  
**Calculate**  $\Delta H$  of the following equation at  $T = 0$  K in units of  $\text{J mol}^{-1}$ .



Assume that:

- only the vibrational energy contributes to the  $\Delta H$ .
- the  $k$  values for  $\text{H}_2$ ,  $\text{HD}$ , and  $\text{D}_2$  are identical.
- the mass of H is 1 Da and the mass of D is 2 Da.



GBR-1 C-2 Q-3

# Q2-3

British English (United Kingdom)

In a system in equilibrium, the molar ratios among  $\text{H}_2$ , HD, and  $\text{D}_2$  depend on the temperature.

We define  $\Delta_{\text{D}_2}$  as the change of the molar ratio of  $\text{D}_2$ .

$$\Delta_{\text{D}_2} = \frac{R_{\text{D}_2}}{R_{\text{D}_2}^*} - 1 \quad (9)$$

Here,  $R_{\text{D}_2}$  refers to  $\frac{[\text{D}_2]}{[\text{H}_2]}$  in the sample and  $R_{\text{D}_2}^*$  to  $\frac{[\text{D}_2]}{[\text{H}_2]}$  at  $T \rightarrow +\infty$ .

It should be noted here that the distribution of isotopes becomes random at  $T \rightarrow +\infty$ .

**A.3** **Calculate**  $\Delta_{\text{D}_2}$  when the isotopic exchange is equilibrated at the temperature where  $K$  in eq. 4 is 0.300. 10pt  
Assume that the natural abundance of D is  $1.5576 \times 10^{-4}$  and the natural abundance of H is  $(1 - 1.5576 \times 10^{-4})$ .

A doubly substituted isotopologue contains two heavy isotope atoms in one molecule.

The molar ratio of the doubly substituted isotopologue increases with decreasing temperature.

Consider the molar ratio of CO<sub>2</sub> molecules with molecular weights of 44 and 47, which are described as CO<sub>2</sub>[44] and CO<sub>2</sub>[47] below. The quantity Δ<sub>47</sub> is defined as:

$$\Delta_{47} = \frac{R_{47}}{R_{47}^*} - 1 \quad (10)$$

$R_{47}$  refers to  $\frac{[\text{CO}_2[47]]}{[\text{CO}_2[44]]}$  in the sample and  $R_{47}^*$  to  $\frac{[\text{CO}_2[47]]}{[\text{CO}_2[44]]}$  at  $T \rightarrow +\infty$ .

The natural abundances of carbon and oxygen atoms are shown below; ignore isotopes that are not shown here.

	<sup>12</sup> C	<sup>13</sup> C
natural abundance	0.988888	0.011112

	<sup>16</sup> O	<sup>17</sup> O	<sup>18</sup> O
natural abundance	0.997621	0.0003790	0.0020000

The temperature dependence of Δ<sub>47</sub> is determined as follows, where  $T$  is given as the absolute temperature in units of K:

$$\Delta_{47} = \frac{36.2}{T^2} + 2.920 \times 10^{-4} \quad (11)$$

- A.4** The  $R_{47}$  of fossil plankton obtained from the Antarctic seabed was  $4.50865 \times 10^{-5}$ . **9pt**  
**Estimate** the temperature using this  $R_{47}$ .  
 This temperature corresponds to the air temperature during the era in which the plankton lived.  
 Consider only the most common isotopologue of CO<sub>2</sub>[47] for the calculation.



GBR-1 C-2 A-1

**A2-1**  
British English (United Kingdom)

## Isotope Time Capsule

**A.1** (8 pt)

$T \rightarrow 0 : K =$  \_\_\_\_\_ ,  $T \rightarrow +\infty : K =$  \_\_\_\_\_



GBR-1 C-2 A-2

**A2-2**  
British English (United Kingdom)

**A.2** (8 pt)

$\Delta H =$  \_\_\_\_\_  $\text{J mol}^{-1}$





GBR-1 C-2 A-3

# A2-3

British English (United Kingdom)

**A.3** (10 pt)

$\Delta_{D_2} =$  \_\_\_\_\_



GBR-1 C-2 A-4

**A2-4**  
British English (United Kingdom)

**A.4** (9 pt)

$T =$  \_\_\_\_\_  $K$

GBR-1 C-3 C-1

**GBR-1 C-3 C**  
Alexander Thow

**IChO**  
**Problem 3**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.



GBR-1 C-3 Q-1

# Q3-1

British English (United Kingdom)

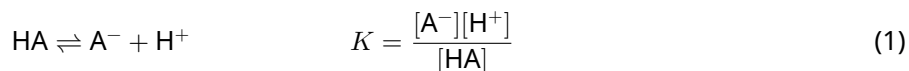
## Lambert-Beer Law?

8 % of the total				
Question	A.1	B.1	B.2	Total
Points	10	6	6	<b>22</b>
Score				

In this problem, ignore the absorption of the cell and the solvent. The temperatures of all solutions and gases are kept constant at 25 °C.

### Part A

An aqueous solution **X** was prepared using HA and NaA. The concentrations  $[A^-]$ ,  $[HA]$ , and  $[H^+]$  in solution **X** are  $1.00 \times 10^{-2} \text{ mol L}^{-1}$ ,  $1.00 \times 10^{-3} \text{ mol L}^{-1}$ , and  $1.00 \times 10^{-4} \text{ mol L}^{-1}$ , respectively, which are correlated via the following acid-base equilibrium:



In Part A, the optical path length is  $l$ . Ignore the density change upon dilution. Assume that no chemical reactions other than eq 1 occur.

- A.1** The absorbance of **X** was  $A_1$  at a wavelength of  $\lambda_1$ . 10pt  
Solution **X** was then diluted to twice its initial volume using hydrochloric acid with a pH = 2.500. After the dilution, the absorbance was still  $A_1$  at  $\lambda_1$ .
- Determine** the ratio  $\varepsilon_{HA}/\varepsilon_{A^-}$ , where  $\varepsilon_{HA}$  and  $\varepsilon_{A^-}$  represent the absorption coefficients (at  $\lambda_1$ ) of HA and  $A^-$ , respectively.

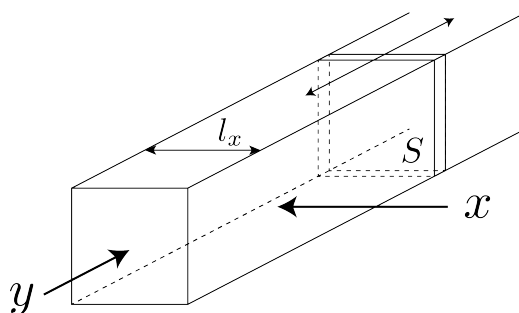
## Part B

Consider the following equilibrium in the gas phase.



A cuboidal container has a transparent movable wall with a cross-section of  $S$  (see the figure below). The container is filled with pure gas D at a pressure  $P$ , and equilibrium is established while the total pressure is maintained at  $P$ .

The absorbance of the gas is  $A = \varepsilon(n/V)l$ , where  $\varepsilon$  is the absorption coefficient,  $n$  the amount of the gas in moles,  $V$  the volume of the gas, and  $l$  the optical path length. Assume that all components of the gas mixture behave as ideal gases.



Use the following definitions if necessary.

	Initial state		After equilibrium	
	D	M	D	M
Partial pressure	$P$	0	$p_D$	$p_M$
Amount in moles	$n_0$	0	$n_D$	$n_M$
Volume	$V_0$		$V$	

**B.1** The absorbance of the gas at  $\lambda_{B1}$  measured in direction  $x$  (i.e.  $l = l_x$ ) was  $A_{B1}$  both at the initial state and after the equilibrium. 6pt

**Determine** the ratio  $\varepsilon_D/\varepsilon_M$  at  $\lambda_{B1}$ , where  $\varepsilon_D$  and  $\varepsilon_M$  represent the absorption coefficients of D and of M, respectively.

**B.2** The absorbance of the gas at  $\lambda_{B2}$  measured in direction  $y$  was  $A_{B2}$  both at the initial state (where  $l = l_{y0}$ ) and after the equilibrium (where  $l = l_y$ ). 6pt

**Determine** the ratio  $\varepsilon_D/\varepsilon_M$  at  $\lambda_{B2}$ .



GBR-1 C-3 A-1

**A3-1**  
British English (United Kingdom)

## Lambert-Beer Law?

### Part A

**A.1** (10 pt)

(Continued on the next page)



GBR-1 C-3 A-2

**A3-2**  
British English (United Kingdom)

A.1 (cont.)

$\epsilon_{\text{HA}}/\epsilon_{\text{A}^-} =$   
\_\_\_\_\_



GBR-1 C-3 A-3

**A3-3**  
British English (United Kingdom)

**Part B**

**B.1** (6 pt)

$\epsilon_D/\epsilon_M =$  \_\_\_\_\_





GBR-1 C-3 A-4

**A3-4**  
British English (United Kingdom)

**B.2** (6 pt)

$\epsilon_D/\epsilon_M =$   
\_\_\_\_\_

GBR-1 C-4 C-1

**GBR-1 C-4 C**  
Alexander Thow

**IChO**  
**Problem 4**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## The Redox Chemistry of Zinc

11 % of the total							
Question	A.1	A.2	B.1	B.2	B.3	B.4	Total
Points	6	5	4	3	5	9	<b>32</b>
Score							



Zinc has long been used in alloys for brass and steel materials. The zinc contained in industrial wastewater is separated by precipitation to purify the water. The precipitate is reduced to recover metallic zinc.

### Part A

The dissolution equilibrium of zinc hydroxide  $\text{Zn(OH)}_2(\text{s})$  at 25 °C and the relevant equilibrium constants are given in eq. 1-4.



The solubility,  $S$ , of zinc (concentration of zinc in a saturated aqueous solution) is given in eq. 5.

$$S = [\text{Zn}^{2+}(\text{aq})] + [\text{Zn}(\text{OH})_2(\text{aq})] + [\text{Zn}(\text{OH})_4^{2-}(\text{aq})] \quad (5)$$

**A.1** **Calculate** the pH range in which  $[\text{Zn}(\text{OH})_2(\text{aq})]$  is the greatest among  $[\text{Zn}^{2+}(\text{aq})]$ ,  $[\text{Zn}(\text{OH})_2(\text{aq})]$  and  $[\text{Zn}(\text{OH})_4^{2-}(\text{aq})]$  once the equilibria in eq. 1–4 are established. 6pt

**A.2** A saturated aqueous solution of  $\text{Zn}(\text{OH})_2(\text{s})$  with pH = 7.00 was prepared and filtered. NaOH was added to this filtrate to increase its pH to 12.00. 5pt

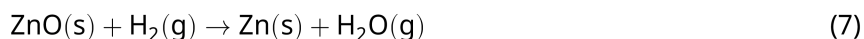
**Calculate** the molar percentage of zinc that precipitates when the pH is increased from 7.00 to 12.00.  
Ignore the volume and temperature changes.

## Part B

Next, the recovered zinc hydroxide is heated to obtain zinc oxide according to the reaction below:



The zinc oxide is then reduced to metallic zinc by reaction with hydrogen:



**B.1** It is necessary to reduce the partial pressure of the water vapour formed, so reaction (7) can proceed at a constant hydrogen pressure of 1 bar. 4pt

**Calculate** the upper limit for the partial pressure of water vapour to allow reaction (7) to proceed at 300 °C.

Gibbs free energies of formation at 300 °C and 1 bar for all gaseous species are:

$$\Delta G_{\text{ZnO}}(300^\circ\text{C}) = -2.90 \times 10^2 \text{ kJ mol}^{-1}$$

$$\Delta G_{\text{H}_2\text{O}}(300^\circ\text{C}) = -2.20 \times 10^2 \text{ kJ mol}^{-1}$$

Metallic zinc is used as a negative electrode (anode) material for metal-air batteries. The electrode consists of Zn and ZnO.

At 25 °C and pressure of 1 bar, the following redox reaction generates electricity with the electromotive force (e.m.f.),  $E^\circ$ .



**B.2** A zinc-air battery was discharged at 20 mA for 24 hours.

3pt

**Calculate** the change in mass of the negative electrode (anode) of the battery.



Mt. Fuji

**B.3** Consider the change in e.m.f. of a zinc-air battery due to the environment.

5pt

**Calculate** the e.m.f. at the summit of Mt. Fuji, where the temperature and altitude are  $-38^\circ\text{C}$  (February) and 3776 m, respectively.

The atmospheric pressure is represented by

$$P [\text{bar}] = 1.013 \times \left( 1 - \frac{0.0065h}{T + 0.0065h + 273.15} \right)^{5.257} \quad (9)$$

at altitude  $h$  [m] and temperature  $T$  [ $^\circ\text{C}$ ].

The molar proportion of oxygen in the atmosphere is 21%.

The Gibbs free energy change of reaction (8) at  $-38^\circ\text{C}$  and 1 bar is

$$\Delta G_{\text{ZnO}}(-38^\circ\text{C}) = -3.26 \times 10^2 \text{ kJ mol}^{-1}.$$



GBR-1 C-4 Q-4

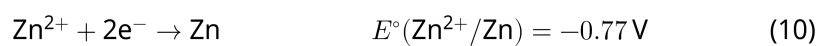
# Q4-4

British English (United Kingdom)

**B.4** **Calculate** the Gibbs free energy change for reaction (6) at 25 °C.

9pt

The standard reduction potentials at 25 °C and 1 bar,  $E^\circ(\text{Zn}^{2+}/\text{Zn})$  and  $E^\circ(\text{O}_2/\text{H}_2\text{O})$  are given as (10) and (11).





GBR-1 C-4 A-1

**A4-1**  
British English (United Kingdom)

## The Redox Chemistry of Zinc

### Part A

**A.1** (6 pt)

\_\_\_\_\_ < pH < \_\_\_\_\_



GBR-1 C-4 A-2

**A4-2**  
British English (United Kingdom)

**A.2** (5 pt)

\_\_\_\_\_ %





GBR-1 C-4 A-3

**A4-3**  
British English (United Kingdom)

**Part B**

**B.1** (4 pt)

$p_{\text{H}_2\text{O}} =$  \_\_\_\_\_ bar

**B.2** (3 pt)

\_\_\_\_\_ g



GBR-1 C-4 A-4

**A4-4**  
British English (United Kingdom)

**B.3** (5 pt)

\_\_\_\_\_ v



GBR-1 C-4 A-5

**A4-5**  
British English (United Kingdom)

**B.4** (9 pt)

$\Delta G^\circ =$  \_\_\_\_\_  $\text{J mol}^{-1}$

GBR-1 C-5 C-1

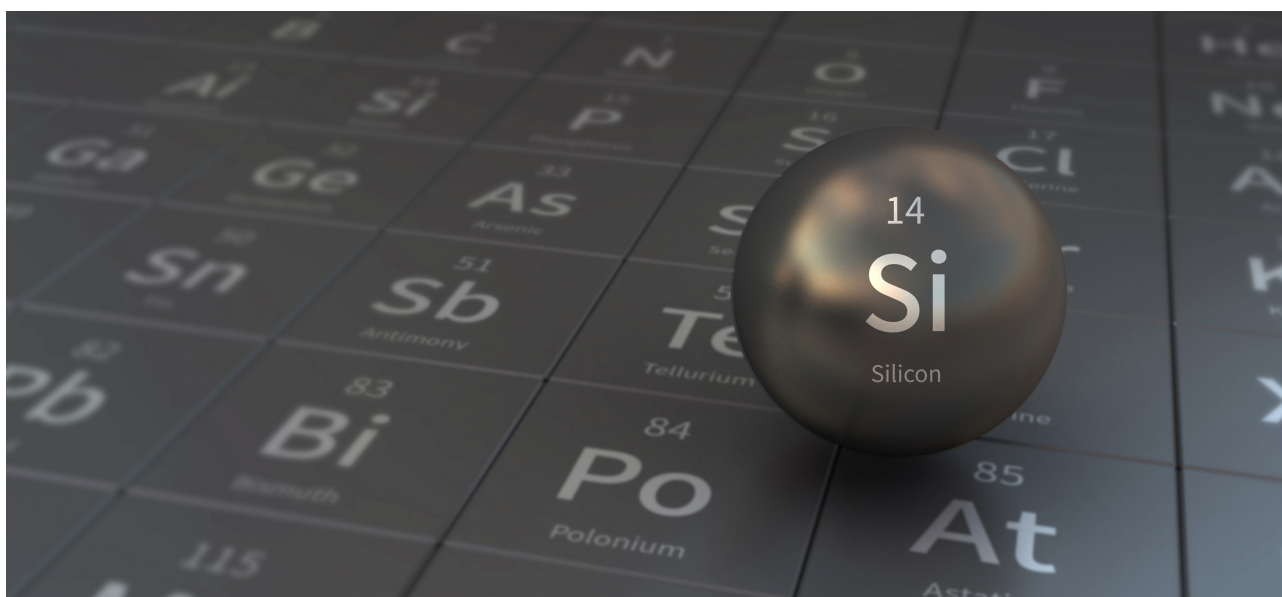
**GBR-1 C-5 C**  
Alexander Thow

**ICHO**  
**Problem 5**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Mysterious Silicon

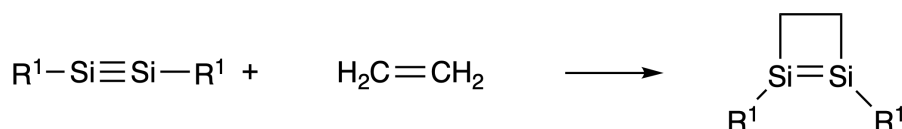
12 % of the total								
Question	A.1	A.2	A.3	A.4	B.1	B.2	B.3	Total
Points	9	7	6	10	5	15	8	<b>60</b>
Score								



The group 14 elements carbon and silicon differ significantly in their properties.

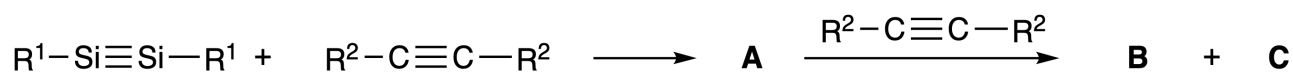
### Part A

Unlike the carbon-carbon triple bond, the silicon-silicon triple bond,  $R^1-Si \equiv Si-R^1$  (R: organic substituent), is extremely reactive. For example, it reacts with ethene to form a four-membered ring.



When  $R^1-Si \equiv Si-R^1$  is treated with an alkyne ( $R^2-C \equiv C-R^2$ ), the four-membered-ring compound **A** is formed as an initial intermediate. Further reaction of **A** with another molecule of  $R^2-C \equiv C-R^2$  gives isomers **B** and **C**, both of which have benzene-like cyclic conjugated structures.

These so-called 'disilabenzenes' contain a six-membered ring and can be formulated as  $(R^1-Si)_2(R^2-C)_4$ .



$^{13}\text{C}$  NMR analysis of the corresponding  $\text{Si}_2\text{C}_4$  six-membered ring skeletons shows two signals for **B** and one signal for **C**.

**A.1** **Draw** the structural formulae of **A**, **B**, and **C** using  $R^1$ ,  $R^2$ , Si, and C, as one of the possible resonance structures. 9pt

**A.2** **Calculate** the aromatic stabilisation energy (ASE) for benzene and **C** (in the case of  $R^1 = R^2 = \text{H}$ ) as positive values, given the enthalpy change of some hydrogenation reactions of unsaturated systems shown below (Fig. 1). 7pt

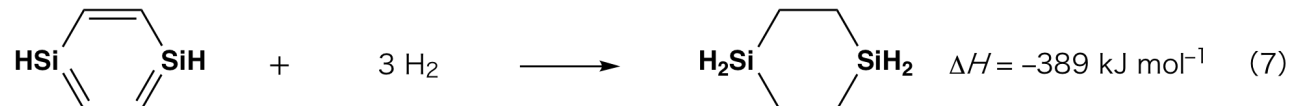
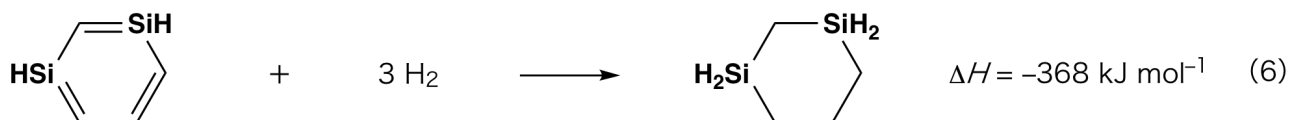


Fig. 1

When a xylene solution of **C** is heated, it undergoes isomerisation to give an equilibrium mixture of compounds **D** and **E**. The molar ratio is **D** : **E** = 1 : 40.0 at 50.0 °C and **D** : **E** = 1 : 20.0 at 120.0 °C.

**A.3** **Calculate**  $\Delta H$  for the transformation of **D** to **E**. 6pt  
Assume that  $\Delta H$  does not depend on temperature.

The isomerisation from **C** to **D** and to **E** proceeds via transformations of  $\pi$ -bonds into  $\sigma$ -bonds without breaking any  $\sigma$ -bonds.  $^{13}\text{C}$  NMR analysis revealed one signal for the  $\text{Si}_2\text{C}_4$  skeleton of **D** and two signals for that of **E**. The skeleton of **D** does not contain any three-membered rings, while **E** has two three-membered rings that share an edge.

**A.4** **Draw** the structural formulae of **D** and **E** using  $\text{R}^1$ ,  $\text{R}^2$ , Si, and C. 10pt

## Part B

Silicon is able to form highly coordinated compounds (more than four substituents) with electronegative elements such as fluorine. Like metal fluorides, highly coordinated silicon fluorides can also act as fluorination reagents.

The fluorination reaction of  $\text{CCl}_4$  using  $\text{Na}_2\text{SiF}_6$  was carried out as follows.

- **Standardisation of  $\text{Na}_2\text{SiF}_6$  solution :**

- Preparation

Aqueous solution **F**: 0.855 g of  $\text{Na}_2\text{SiF}_6$  ( $188.053 \text{ g mol}^{-1}$ ) dissolved in water (total volume: 200 mL).

Aqueous solution **G**: 6.86 g of  $\text{Ce}_2(\text{SO}_4)_3$  ( $568.424 \text{ g mol}^{-1}$ ) dissolved in water (total volume: 200 mL).

- Precipitation Titration Procedure

Solution **F** (50.0 mL) was titrated with solution **G** in the presence of xylenol orange, an indicator which coordinates to  $\text{Ce}^{3+}$ . After adding 18.8 mL of solution **G**, the colour of the solution changed from yellow to magenta. The precipitate formed is a binary compound that contains  $\text{Ce}^{3+}$ , and the only resulting silicon compound is  $\text{Si}(\text{OH})_4$ .

**B.1** **Write** the balanced equation for the reaction of  $\text{Na}_2\text{SiF}_6$  with  $\text{Ce}_2(\text{SO}_4)_3$ . 5pt

- **Reaction of  $\text{CCl}_4$  with  $\text{Na}_2\text{SiF}_6$ :**

(Substance losses, e.g. by evaporation, are negligible during the following steps.)

$\text{Na}_2\text{SiF}_6$  ( $x$  [g]) was added to  $\text{CCl}_4$  (500.0 g) and heated to 300 °C in a sealed vessel.

The unreacted  $\text{Na}_2\text{SiF}_6$  and  $\text{NaCl}$  produced were removed by filtration. The filtrate was diluted to a total volume of 1.00 L with  $\text{CCl}_4$  (solution **H**).

The  $^{29}\text{Si}$  and  $^{19}\text{F}$  NMR spectra of solution **H** showed  $\text{SiF}_4$  as the only silicon compound. In the  $^{19}\text{F}$  NMR spectrum, in addition to  $\text{SiF}_4$ , signals corresponding to  $\text{CFCl}_3$ ,  $\text{CF}_2\text{Cl}_2$ ,  $\text{CF}_3\text{Cl}$ , and  $\text{CF}_4$  were observed (cf. Table 1). The integrals in the  $^{19}\text{F}$  NMR spectrum are proportional to the number of fluorine nuclei.

Table 1

<sup>19</sup> F NMR data	CFCl <sub>3</sub>	CF <sub>2</sub> Cl <sub>2</sub>	CF <sub>3</sub> Cl	CF <sub>4</sub>
Integral	45.0	65.0	18.0	2.0

SiF<sub>4</sub> is hydrolysed to form H<sub>2</sub>SiF<sub>6</sub> according to the following eq. 8:



Solution **H** (10 mL) was added to excess water, which resulted in the complete hydrolysis of SiF<sub>4</sub>. After separation, the H<sub>2</sub>SiF<sub>6</sub> generated from the hydrolysis was neutralised and completely converted to Na<sub>2</sub>SiF<sub>6</sub> (aqueous solution **J**).

The precipitate of unreacted Na<sub>2</sub>SiF<sub>6</sub> and NaCl, which was removed by filtration in the initial step (previously underlined), was completely dissolved in water to give an aqueous solution (solution **K**; 10.0 L).

Then, additional precipitation titrations using solution **G** were carried out, and the endpoints of the titrations with **G** were as follows:

•For the entire amount of solution **J**: 61.6 mL.

•For 100 mL of solution **K**: 44.4 mL.

The presence of NaCl or SiO<sub>2</sub> has no effect on the precipitation titration.

**B.2** **Calculate** the mass of NaCl produced in the reaction vessel (information previously underlined), and **calculate** the mass ( $x$  [g]) of the Na<sub>2</sub>SiF<sub>6</sub> used as starting material. 15pt

**B.3** 77.8% of the CCl<sub>4</sub> used as a starting material did not react. **Calculate** the mass of CF<sub>3</sub>Cl generated. 8pt





GBR-1 C-5 A-1

**A5-1**  
British English (United Kingdom)

## Mysterious Silicon

### Part A

**A.1** (9 pt)

**A** (3 pt)

**B** (3 pt)

**C** (3 pt)

**A.2** (7 pt)

$C_6H_6$  : \_\_\_\_\_  $\text{kJ mol}^{-1}$ , **C** : \_\_\_\_\_  $\text{kJ mol}^{-1}$



GBR-1 C-5 A-2

**A5-2**  
British English (United Kingdom)

**A.3** (6 pt)

$\Delta H =$  \_\_\_\_\_  $\text{kJ mol}^{-1}$

**A.4** (10 pt)

**D** (5 pt)

**E** (5 pt)



GBR-1 C-5 A-3

**A5-3**  
British English (United Kingdom)

**Part B**

**B.1** (5 pt)

---

**B.2** (15 pt)

(Continued on the next page)



GBR-1 C-5 A-4

**A5-4**  
British English (United Kingdom)

**B.2 (cont.)**

NaCl : \_\_\_\_\_ g, Na<sub>2</sub>SiF<sub>6</sub> : \_\_\_\_\_ g



GBR-1 C-5 A-5

**A5-5**  
British English (United Kingdom)

**B.3** (8 pt)

$\text{CF}_3\text{Cl}$  : \_\_\_\_\_ g

GBR-1 C-6 C-1

**GBR-1 C-6 C**  
Alexander Thow

**IChO**  
**Problem 6**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## The Solid-State Chemistry of Transition Metals

13 % of the total											
Question	A.1	A.2	A.3	B.1	B.2	B.3	B.4	C.1	C.2	C.3	Total
Points	6	3	3	6	4	4	4	5	5	5	<b>45</b>
Score											



Volcano at Sakurajima island

### Part A

Japan has many volcanoes. When silicate minerals crystallize from magma, some transition-metal ions ( $M^{n+}$ ) in the magma are incorporated into the silicate minerals.

The  $M^{n+}$  studied in this problem are coordinated by oxide ions ( $O^{2-}$ ). They adopt a four-coordinate tetrahedral ( $T_d$ ) geometry in the magma and a six-coordinate octahedral ( $O_h$ ) geometry in the silicate minerals, both of which exhibit a high-spin electron configuration.

The distribution coefficient of  $M^{n+}$  between the silicate minerals and magma,  $D$ , can be expressed by:

$$D = \frac{[M]_s}{[M]_l}$$

where  $[M]_s$  and  $[M]_l$  are the concentrations of  $M^{n+}$  in the silicate minerals and the magma, respectively.

The table below shows the  $D$  values of  $\text{Cr}^{2+}$  and  $\text{Mn}^{2+}$  as examples.

	$\text{Cr}^{2+}$	$\text{Mn}^{2+}$
$D$	7.2	1.1

In an  $O_h$  field,  $\Delta_o$  is the energy separation of the d-orbitals of  $\text{M}^{n+}$  and  $\text{CFSE}^O$  is the crystal-field stabilization energy.

$\Delta_T$  and  $\text{CFSE}^T$  are the equivalents in a  $T_d$  field.

**A.1**  $\Delta\text{CFSE}$  is defined as:

6pt

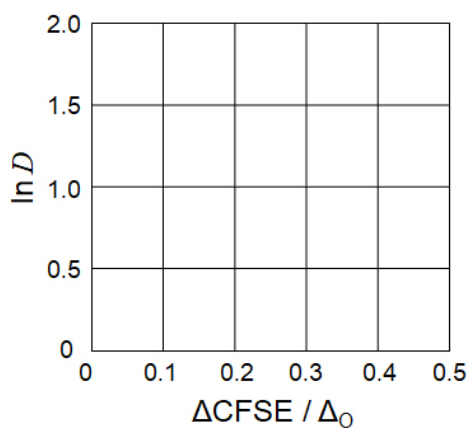
$$\Delta\text{CFSE} = |\text{CFSE}^O - \text{CFSE}^T|$$

**Calculate**  $\Delta\text{CFSE}$  in terms of  $\Delta_o$  for  $\text{Cr}^{2+}$ ,  $\text{Mn}^{2+}$ , and  $\text{Co}^{2+}$ .

Assume  $\Delta_T = 4/9\Delta_o$ .

**A.2** A linear relationship is observed by plotting  $\ln D$  against  $\Delta\text{CFSE} / \Delta_o$  as shown below. 3pt

**Estimate**  $D$  for  $\text{Co}^{2+}$ .





Metal oxides, MO (where M is Ca, Ti, V, Mn, or Co), crystallise in a rock-salt structure wherein the  $M^{n+}$  adopts an  $O_h$  geometry with a high-spin electron configuration.

The lattice enthalpy of these oxides is mainly governed by the Coulomb interactions based on the radius and charge of the ions and some contributions from the CFSE of  $M^{n+}$  in the  $O_h$  field.

**A.3** **Choose** the appropriate set of lattice enthalpies [ $\text{kJ mol}^{-1}$ ] from one of the options (a) to (f). 3pt

	CaO	TiO	VO	MnO	CoO
(a)	3460	3878	3913	3810	3916
(b)	3460	3916	3878	3810	3913
(c)	3460	3913	3916	3810	3878
(d)	3810	3878	3913	3460	3916
(e)	3810	3916	3878	3460	3913
(f)	3810	3913	3916	3460	3878

## Part B

A mixed oxide **A**, which contains  $\text{La}^{3+}$  and  $\text{Cu}^{2+}$ , crystallises in a tetragonal unit cell shown in Fig.1.

The  $[\text{CuO}_6]$  octahedron is distorted from the regular  $O_h$  geometry: the Cu–O length along the  $z$ -axis ( $l_z$ ) is longer than that of the  $x$ -axis ( $l_x$ ).

This distortion removes the degeneracy of the  $e_g$  orbitals ( $d_{x^2-y^2}$  and  $d_{z^2}$ ).

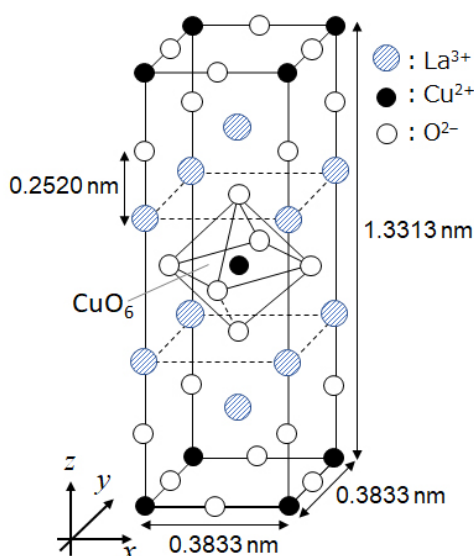


Fig. 1

**A** can be synthesised by thermal decomposition of complex **B**. **B** is formed by mixing metal chlorides in dilute aqueous ammonia solution containing the diacid, squaric acid  $\text{C}_4\text{H}_2\text{O}_4$ .

The thermal decomposition of **B** in dry air shows a weight loss of 29.1% up to 200 °C due to the loss of crystallisation water, followed by another weight loss up to 700 °C due to the release of  $\text{CO}_2$ .

The total weight loss during the formation of **A** from **B** is 63.6%. Only water and  $\text{CO}_2$  are released in the thermal decomposition reaction.

**B.1** Write the chemical formulae for **A** and **B**.

6pt

**B.2** Calculate  $l_x$  and  $l_z$  using Fig. 1.

4pt

**B.3** Complete the diagram in your answer sheet for  $\text{Cu}^{2+}$  in the distorted  $[\text{CuO}_6]$  octahedron in **A** of Fig. 1:

- Write the names of the split  $e_g$  orbitals ( $d_{x^2-y^2}$  and  $d_{z^2}$ ) in (i) and (ii).
- Draw the electron configuration in the dotted box.

**A** is an insulator.

When one  $\text{La}^{3+}$  is substituted with one  $\text{Sr}^{2+}$ , one hole is generated in the crystal lattice that can conduct electricity. As a result, the  $\text{Sr}^{2+}$ -doped **A** shows superconductivity below 38 K.

When a substitution reaction took place for **A**,  $2.05 \times 10^{27}$  holes  $\text{m}^{-3}$  were generated.

**B.4 Calculate** the percentage of  $\text{La}^{3+}$  substituted by  $\text{Sr}^{2+}$  based on the mole ratio 4pt  
in the substitution reaction.  
Note that the valence of the constituent ions and the crystal structure are not altered by the substitution reaction.

### Part C

$\text{Cu}_2(\text{CH}_3\text{CO}_2)_4$  is composed of four  $\text{CH}_3\text{CO}_2^-$  ions coordinated to two  $\text{Cu}^{2+}$  ions (Fig. 2A).

$\text{Cu}_2(\text{CH}_3\text{CO}_2)_4$  exhibits high levels of structural symmetry, with two axes passing through the carbon atoms of the four  $\text{CH}_3\text{CO}_2^-$  and an axis passing through the two  $\text{Cu}^{2+}$ , all of which are oriented orthogonal relative to each other.

A "cage complex" is formed if a dicarboxylate ligand is used instead of  $\text{CH}_3\text{CO}_2^-$ .

The cage complex  $\text{Cu}_4(\text{L1})_4$  is composed of planar dicarboxylate **L1** (Fig. 2B) and  $\text{Cu}^{2+}$  (Fig. 2C). The angle  $\theta$  between the coordination directions of the two carboxylates, indicated by the arrows in Fig. 2B, determines the structure of the cage complex.

For **L1**  $\theta = 0^\circ$ .

Note that hydrogen atoms are not shown in Fig. 2.

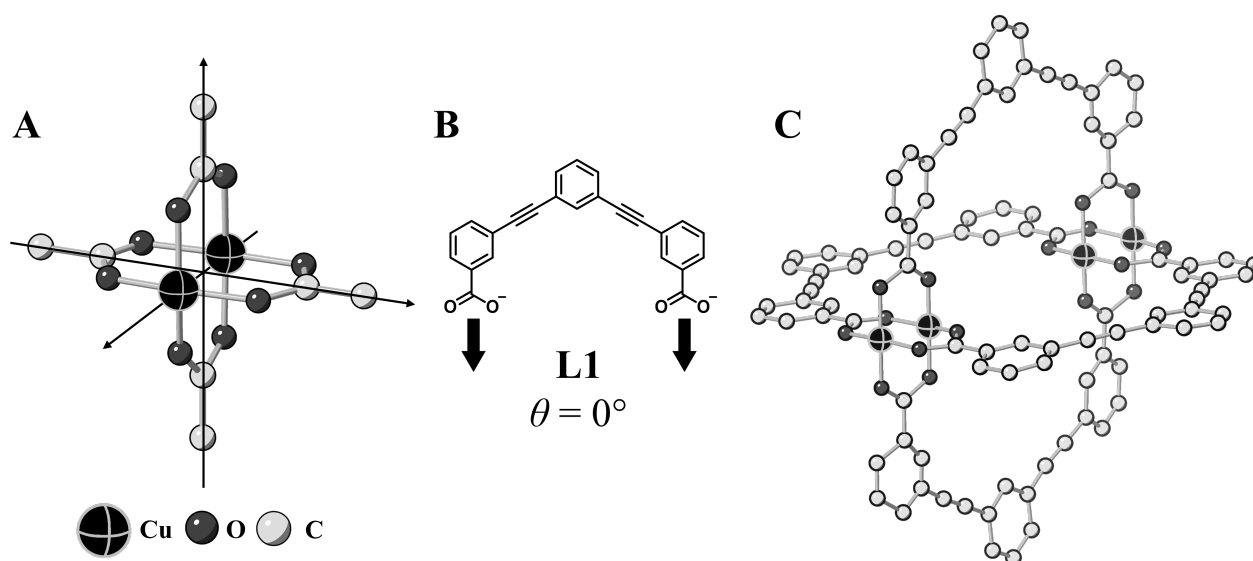


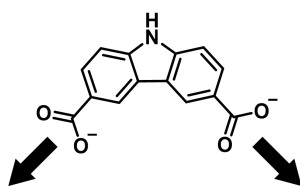
Fig. 2

- C.1 The  $\theta$  of the planar dicarboxylate **L2** below is fixed to  $90^\circ$ .  
 The cage complex formed from **L2** and  $\text{Cu}^{2+}$  is  $\text{Cu}_n(\text{L2})_m$ .

5pt

**Give** the smallest integer combination of  $n$  and  $m$ .

Assume that only the  $\text{CO}_2^-$  groups of **L2** form a coordination bond to  $\text{Cu}^{2+}$  ions.



**L2**

$\theta = 90^\circ$

A zinc complex,  $\text{Zn}_4\text{O}(\text{CH}_3\text{CO}_2)_6$ , contains four tetrahedral  $\text{Zn}^{2+}$ , six  $\text{CH}_3\text{CO}_2^-$ , and one  $\text{O}^{2-}$  (Fig. 3A). In  $\text{Zn}_4\text{O}(\text{CH}_3\text{CO}_2)_6$ , the  $\text{O}^{2-}$  is located at the origin, and the three axes passing through the carbon atoms of  $\text{CH}_3\text{CO}_2^-$  are oriented orthogonal relative to each other.

When *p*-benzenedicarboxylate (Fig. 3B, **L3**,  $\theta = 180^\circ$ ) is used instead of  $\text{CH}_3\text{CO}_2^-$ , the  $\text{Zn}^{2+}$  clusters are linked to each other to form a crystalline solid (**X**) that is called a “porous coordination polymer” (Fig. 3C). The composition of **X** is  $[\text{Zn}_4\text{O}(\text{L3})_3]_n$ , and it has a cubic crystal structure with nano-sized pores.

One pore is represented as a sphere in Fig. 3D, and each tetrahedral  $\text{Zn}^{2+}$  cluster is represented as a dark grey polyhedron in Fig. 3C and 3D. *Note that hydrogen atoms are not shown in Fig. 3.*

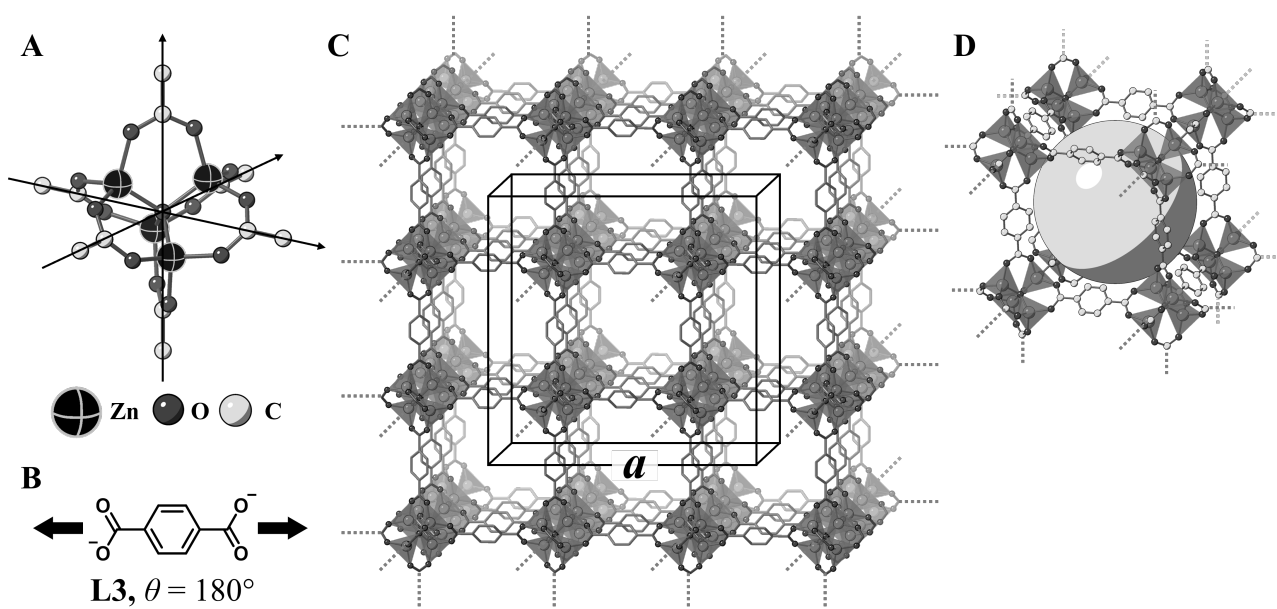


Fig. 3

**C.2** **X** has a cubic unit cell with a side length of  $a$  (see Fig. 3C) and a density of  $0.592 \text{ g cm}^{-3}$ . 5pt

**Calculate**  $a$  in [cm].

**C.3** **X** contains a considerable number of pores. 5pt  
1 g of **X** can accommodate  $3.0 \times 10^2$  mL of  $\text{CO}_2$  gas in the pores at 1 bar and  $25^\circ\text{C}$ .

**Calculate** the average number of  $\text{CO}_2$  molecules per pore.



GBR-1 C-6 A-1

**A6-1**  
British English (United Kingdom)

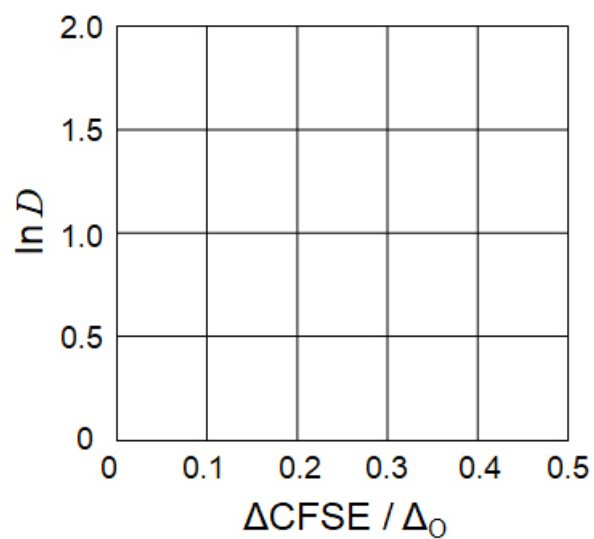
## The Solid-State Chemistry of Transition Metals

### Part A

**A.1** (6 pt)

$\text{Cr}^{2+}$  : \_\_\_\_\_  $\Delta_{\text{O}}$ ,  $\text{Mn}^{2+}$  : \_\_\_\_\_  $\Delta_{\text{O}}$ ,  $\text{Co}^{2+}$  : \_\_\_\_\_  $\Delta_{\text{O}}$

**A.2** (3 pt)



$D$  : \_\_\_\_\_

**A.3** (3 pt)

\_\_\_\_\_



GBR-1 C-6 A-3

**A6-3**  
British English (United Kingdom)

**Part B**

**B.1** (6 pt)

A : \_\_\_\_\_, B : \_\_\_\_\_

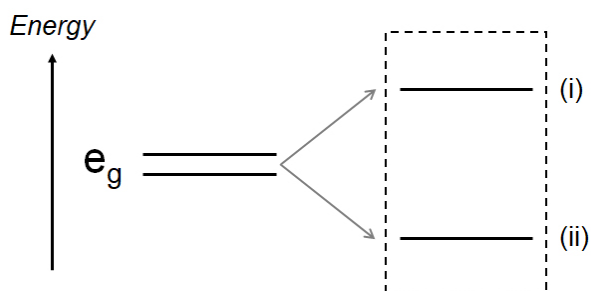
**B.2** (4 pt)

$l_x =$  \_\_\_\_\_ nm,  $l_z =$  \_\_\_\_\_ nm



**B.3** (4 pt)

(i) : \_\_\_\_\_ , (ii) : \_\_\_\_\_



**B.4** (4 pt)

\_\_\_\_\_ %



GBR-1 C-6 A-5

# A6-5

British English (United Kingdom)

## Part C

**C.1** (5 pt)

$n =$  \_\_\_\_\_,  $m =$  \_\_\_\_\_

**C.2** (5 pt)

$a =$  \_\_\_\_\_ cm



GBR-1 C-6 A-6

**A6-6**  
British English (United Kingdom)

**C.3** (5 pt)

---

GBR-1 C-7 C-1

**GBR-1 C-7 C**  
Alexander Thow

**IChO**  
**Problem 7**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

## Playing with Non-benzenoid Aromaticity

13 % of the total					
Question	A.1	A.2	A.3	B.1	Total
Points	5	2	19	10	<b>36</b>
Score					

Prof. Nozoe (1902–1996) founded the research field of non-benzenoid aromatic compounds.



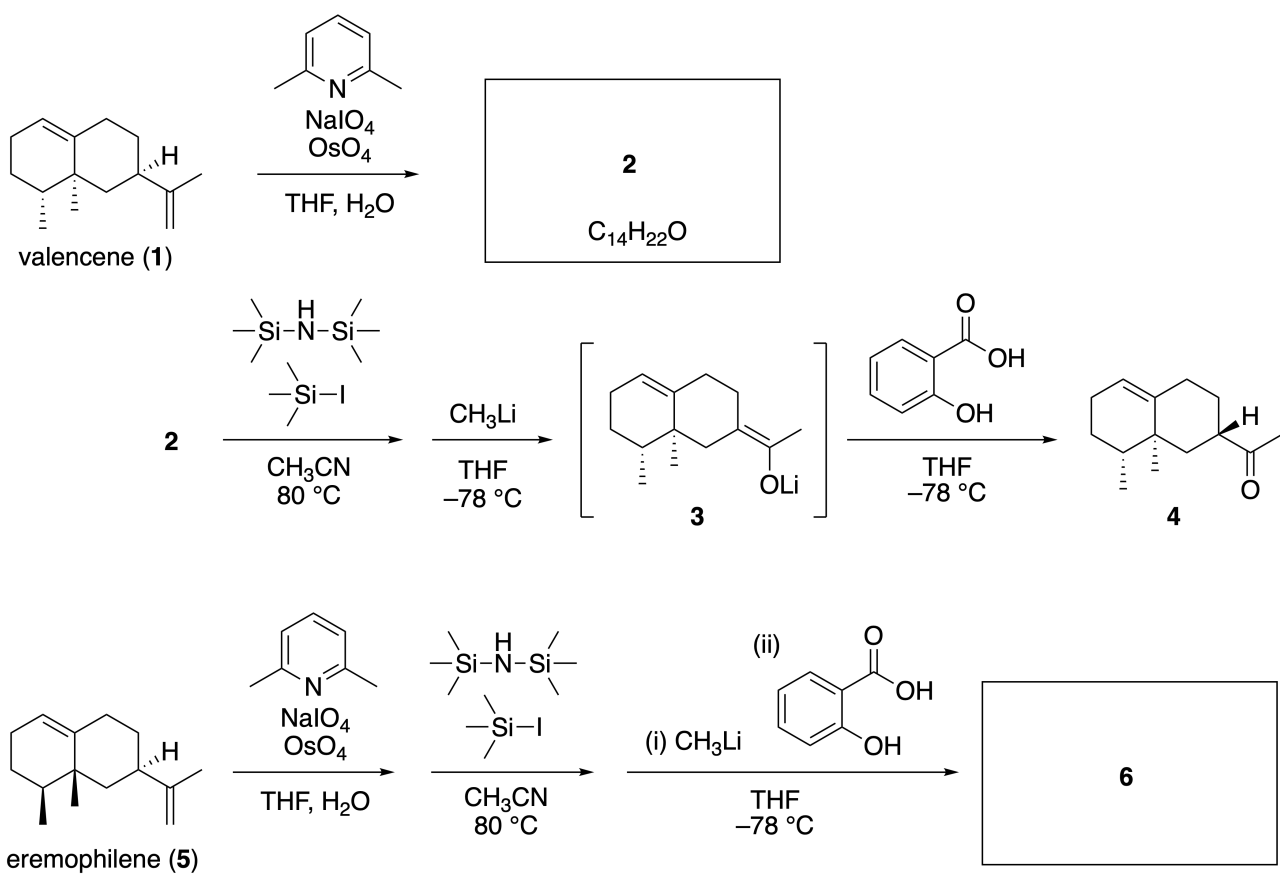
Photo credit: Tohoku Univ.

### Part A

Lineariifolianone is a natural product which was isolated from *Inula linariifolia*. From valencene (**1**), a one-step conversion yields **2**, before a three-step conversion via **3** yields ketone **4**. Eremophilene (**5**) is converted into **6** by performing the same four-step conversion.



*Inula linariifolia*



**A.1** Draw the structures of **2** and **6** and clearly identify the stereochemistry where necessary. 5pt



GBR-1 C-7 Q-3

# Q7-3

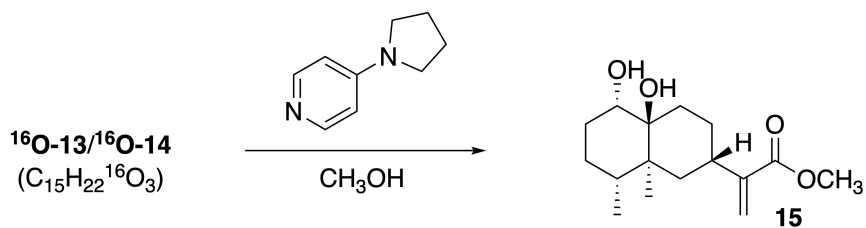
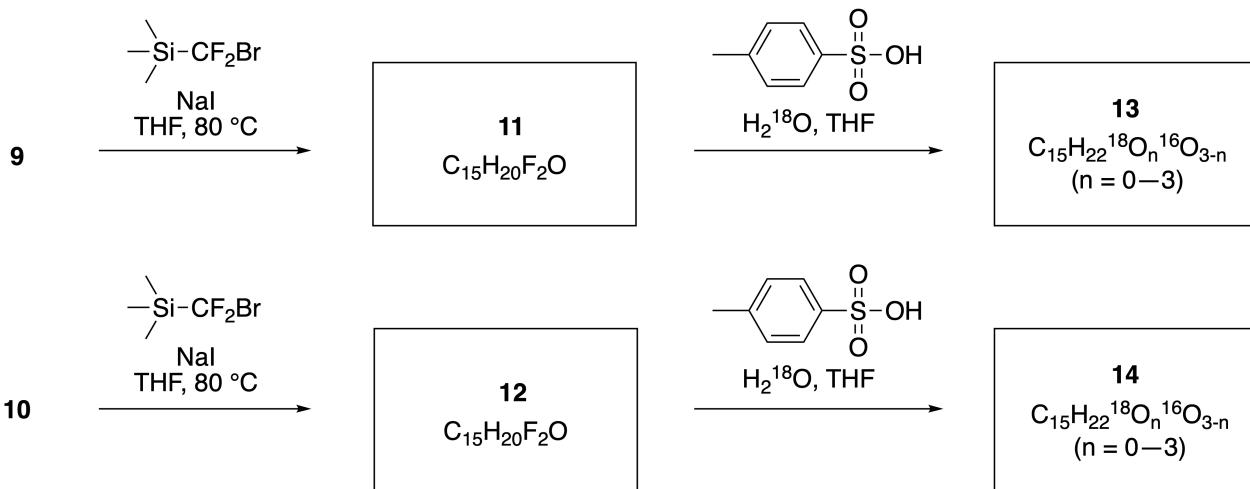
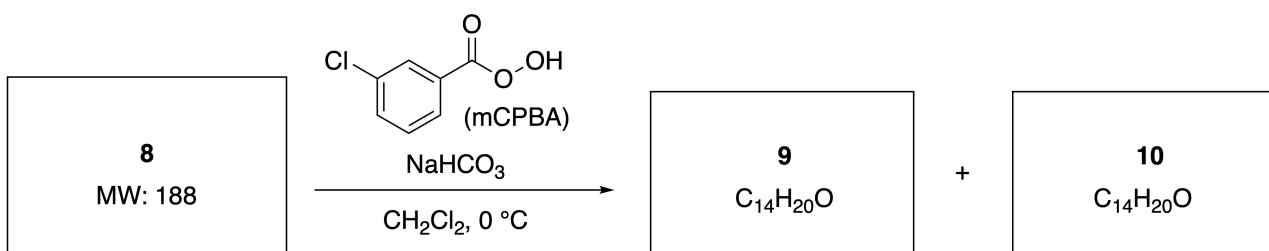
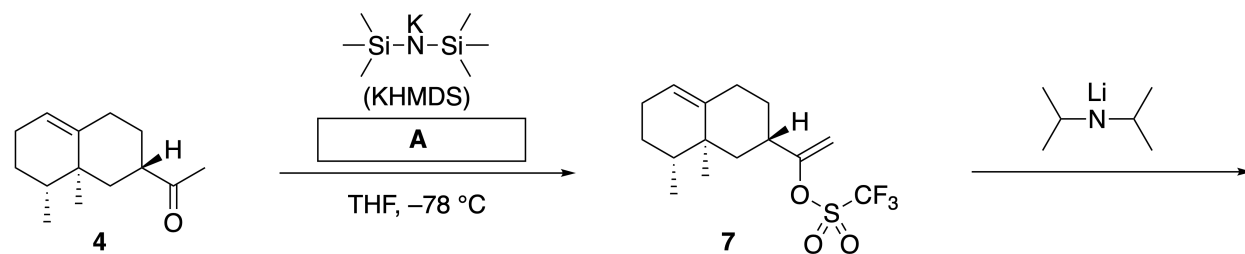
British English (United Kingdom)

---

Then, ketone **4** is converted into ester **15**. Compound **8** (molecular weight: 188) retains all the stereocentres in **7**. Compounds **9** and **10** have five stereocentres and no carbon-carbon double bonds.

Assume that  $\text{H}_2^{18}\text{O}$  is used instead of  $\text{H}_2^{16}\text{O}$  for the synthesis of  $^{18}\text{O}$ -labelled-linearifolianones **13** and **14** from **11** and **12**, respectively.

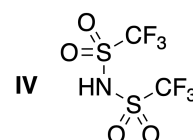
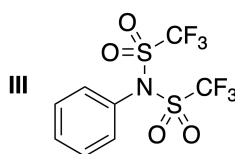
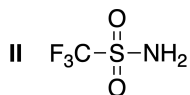
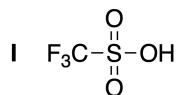
Compounds **13** and **14** are  $^{18}\text{O}$ -labelled isotopomers. Both **13** and **14** provide the same product **15** with identical stereochemistry ignoring isotopic labelling.





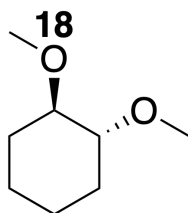
**A.2** Choose the appropriate structure for **A**.

2pt



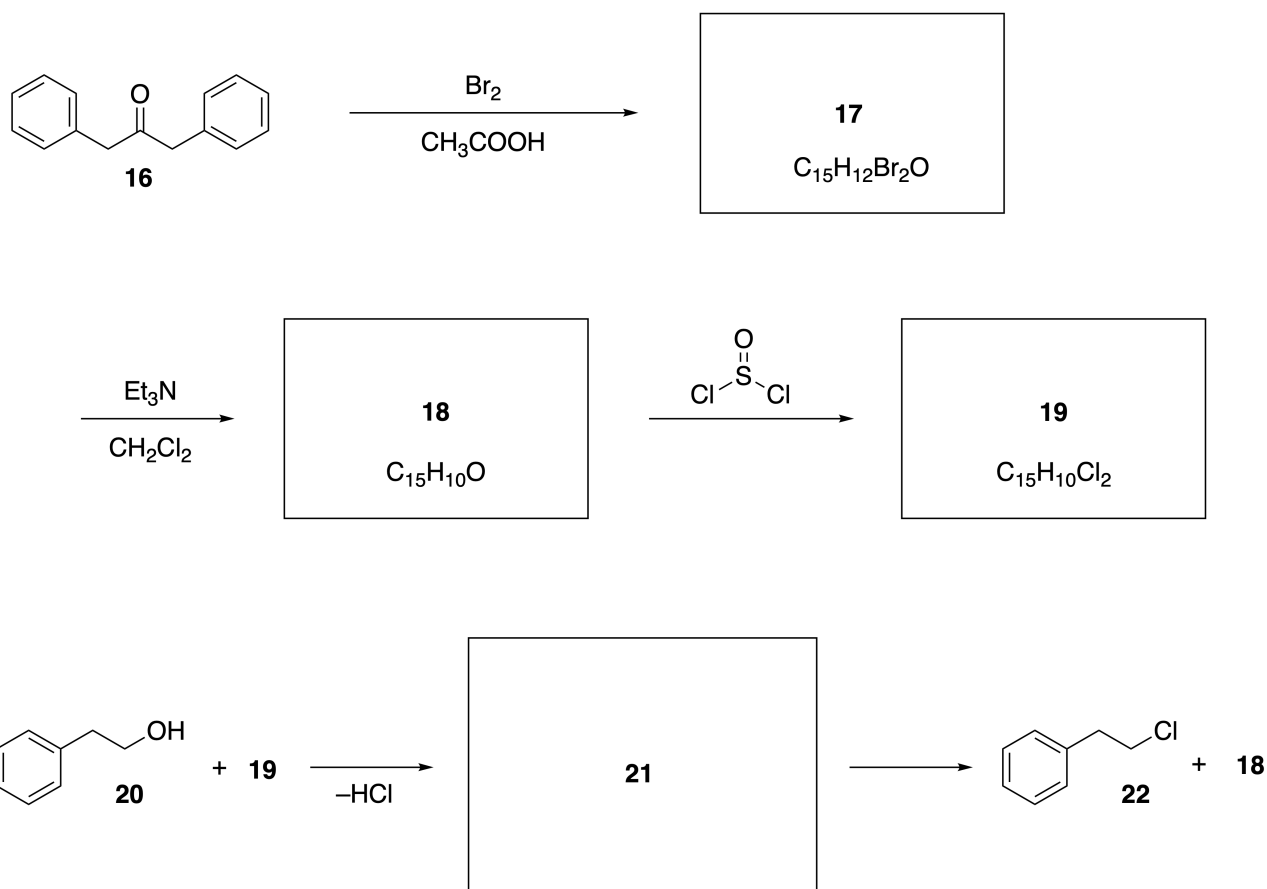
**A.3** Draw the structures of **8–14** and clearly identify the stereochemistry where necessary. Also, indicate the introduced  $^{18}\text{O}$  atoms for **13** and **14** as shown in the example below.

19pt



## Part B

Compound **19** is synthesised as shown below. Linking to non-benzenoid aromaticity, **19** can be used as an activator for alcohols, and **20** was converted to **22** via ion-pair intermediate **21**. Although the formation of **21** was observed by NMR, **21** gradually decomposes to give **18** and **22**.



$^1\text{H}$  NMR ( $\text{CD}_3\text{CN}$ , ppm) **20**:  $\delta$  7.4–7.2 (5H), 3.7 (2H), 2.8 (2H), 2.2 (1H)

**21**:  $\delta$  8.5–7.3 (15H), 5.5 (2H), 3.4 (2H)

**B.1** **Draw** the structures of **17–19** and **21**. Stereochemistry is not required.

10pt



GBR-1 C-7 A-1

**A7-1**  
British English (United Kingdom)

## Playing with Non-benzenoid Aromaticity

### Part A

**A.1** (5 pt)

**2** (2 pt)

**6** (3 pt)

**A.2** (2 pt)



GBR-1 C-7 A-2

**A7-2**  
British English (United Kingdom)

**A.3** (19 pt)

**8** (3 pt)



**9** (2 pt)

**10** (2 pt)



**11** (2 pt)

**12** (2 pt)



**13** (4 pt)

**14** (4 pt)



GBR-1 C-7 A-3

# A7-3

British English (United Kingdom)

## Part B

**B.1** (10 pt)

**17** (2 pt)

**18** (2 pt)

**19** (3 pt)

**21** (3 pt)

GBR-1 C-8 C-1

**GBR-1 C-8 C**  
Alexander Thow

**IChO**  
**Problem 8**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

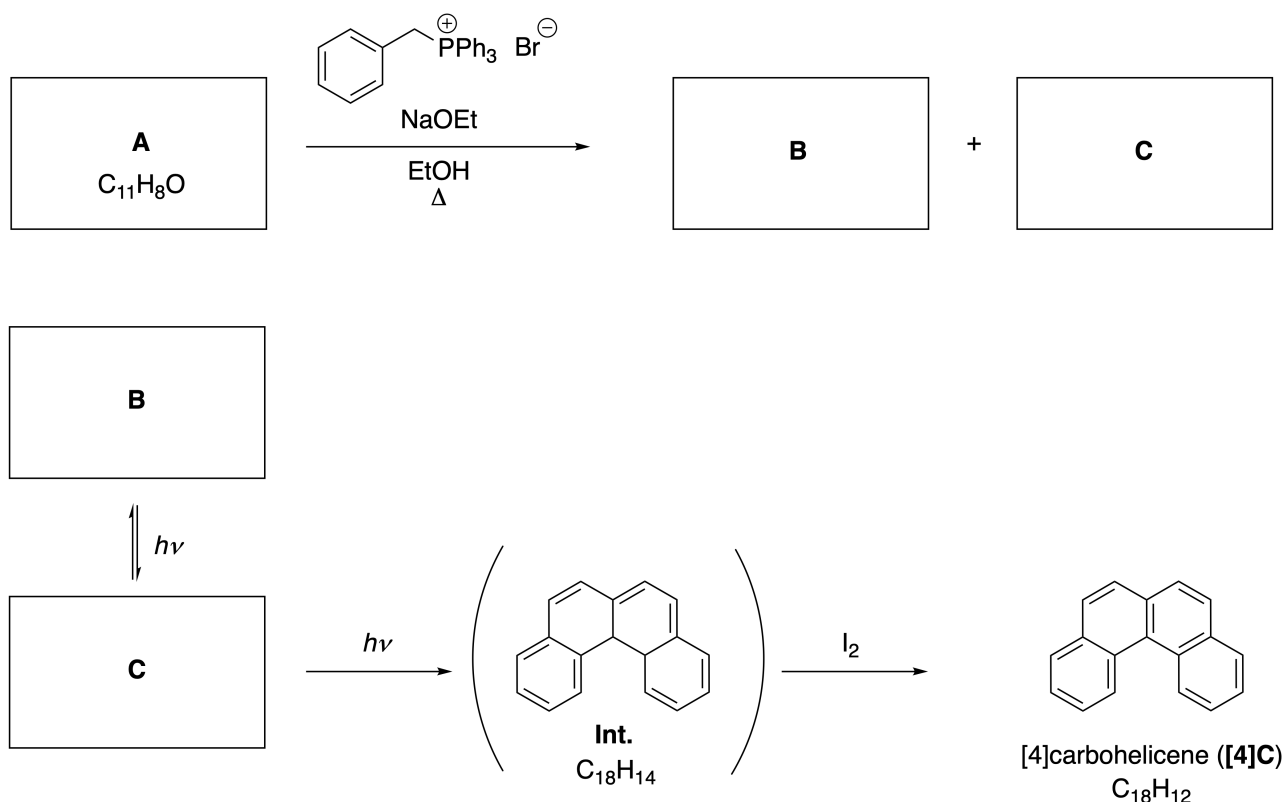
## Dynamic Organic Molecules and Their Chirality

11 % of the total						
Question	A.1	A.2	A.3	B.1	B.2	Total
Points	9	3	7	3	4	<b>26</b>
Score						

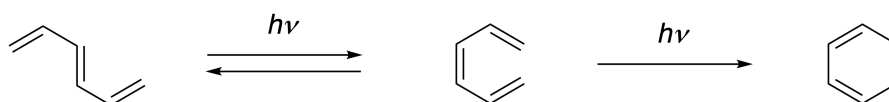
### Part A

Polycyclic aromatic hydrocarbons with successive ortho-connections are called [n]carbohelicenes (here, n represents the number of six-membered rings) (see below).

[4]Carbohelicene (**[4]C**) is efficiently prepared by a route using a photoreaction as shown below, via an intermediate (**Int.**) that is readily oxidized by iodine.



The photoreaction proceeds in a manner similar to the following example.





GBR-1 C-8 Q-2

# Q8-2

British English (United Kingdom)

**Note:** For all of Question 8, please draw alternating single and double bonds in your answers to the problems as depicted in the examples of carbohelicene. Do not use circles for conjugated  $\pi$  systems.

**A.1** Draw the structures of **A-C**. Stereoisomers should be distinguished. 9pt

**A.2** Attempts to synthesize [5]carbohelicene from the same phosphonium salt and an appropriate starting compound resulted in the formation of only a trace amount of [5]carbohelicene, instead affording product **D** whose molecular weight was 2 Da lower than that of [5]carbohelicene. 3pt

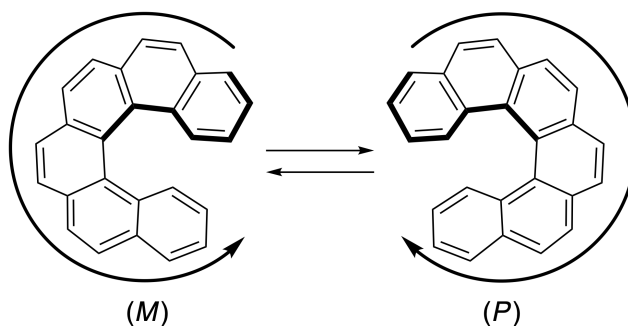
Draw the structure of **D**.

The  $^1\text{H}$  NMR chemical shifts of **D** are listed below.

**[D** ( $\delta$ , ppm in  $\text{CS}_2$ , r.t.), 8.85 (2H), 8.23 (2H), 8.07 (2H), 8.01 (2H), 7.97 (2H), 7.91 (2H)]



[5]- and larger [n]carbohelicenes have helical chirality and interconversion between enantiomers of these helicenes is very slow at room temperature. The chirality of [n]carbohelicenes is defined as (*M*) or (*P*) as shown below.



[n]Carbohelicenes with *n* larger than 4 can be enantiomerically separated by chiral column chromatography, which was developed by Prof. Yoshio Okamoto.

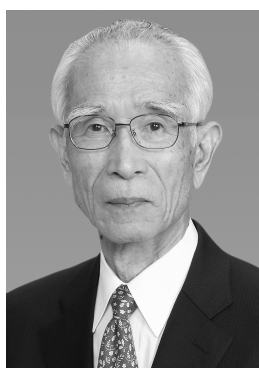
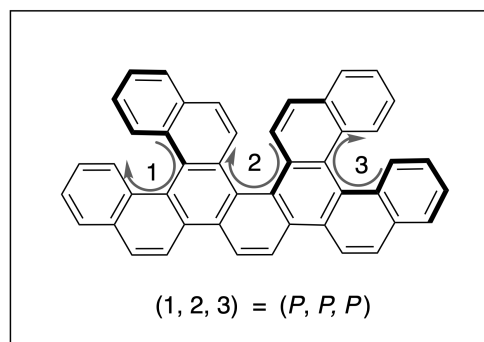
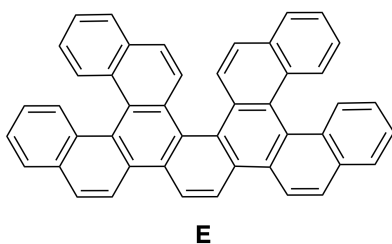


Photo credit: The Japan Prize Foundation

A molecule that contains two or more helicene-like structures is called a multiple-helicene. If helical chirality is considered there can be several stereoisomers of a multiple-helicene.

For example, compound **E** contains three [5]carbohelicene-like moieties in one molecule. One of the stereoisomers is described as  $(P, P, P)$  as shown below.



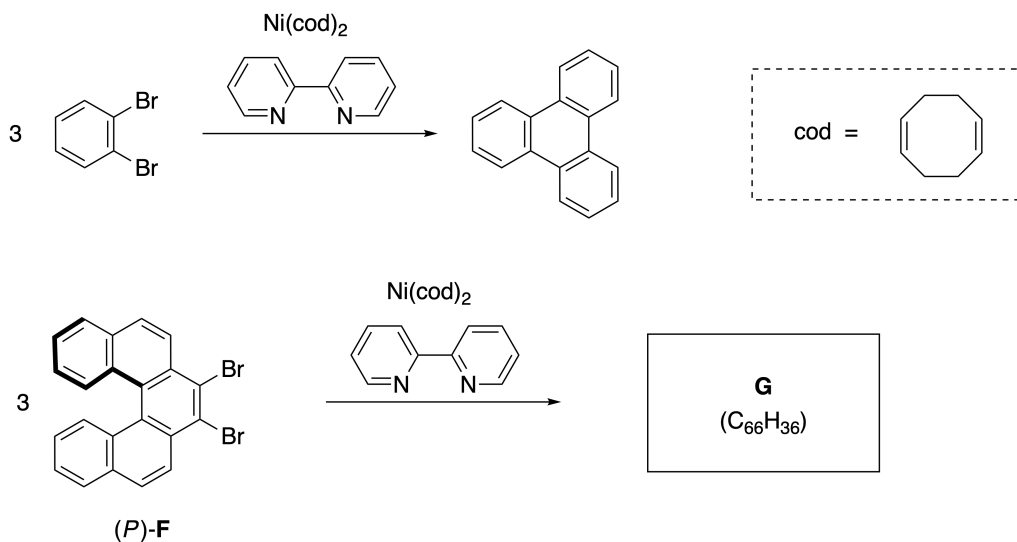
**A.3** The nickel-mediated trimerisation of 1,2-dibromobenzene generates triphenylene. 7pt  
When the same reaction is applied to an enantiomer of **F**, (*P*)-**F**, multiple-helicene **G** (C<sub>66</sub>H<sub>36</sub>) is obtained.

Given that interconversion between stereoisomers does not occur during the reaction, **identify all** the possible stereoisomers of **G** formed in this process, without duplicating stereoisomers.

As a reference, one isomer should be drawn completely with the chirality defined as in the example above, with numerical labels.

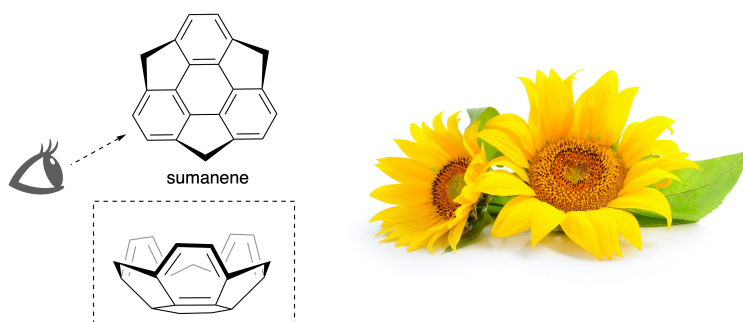
The other stereoisomers should be listed with location numbers and *M* and *P* labels according to the same numbering.

For instance, the other stereoisomers of **E** should be listed as (1, 2, 3) = (*P*, *M*, *P*), (*P*, *M*, *M*), (*P*, *P*, *M*), (*M*, *M*, *M*), (*M*, *M*, *P*), (*M*, *P*, *P*), and (*M*, *P*, *M*).

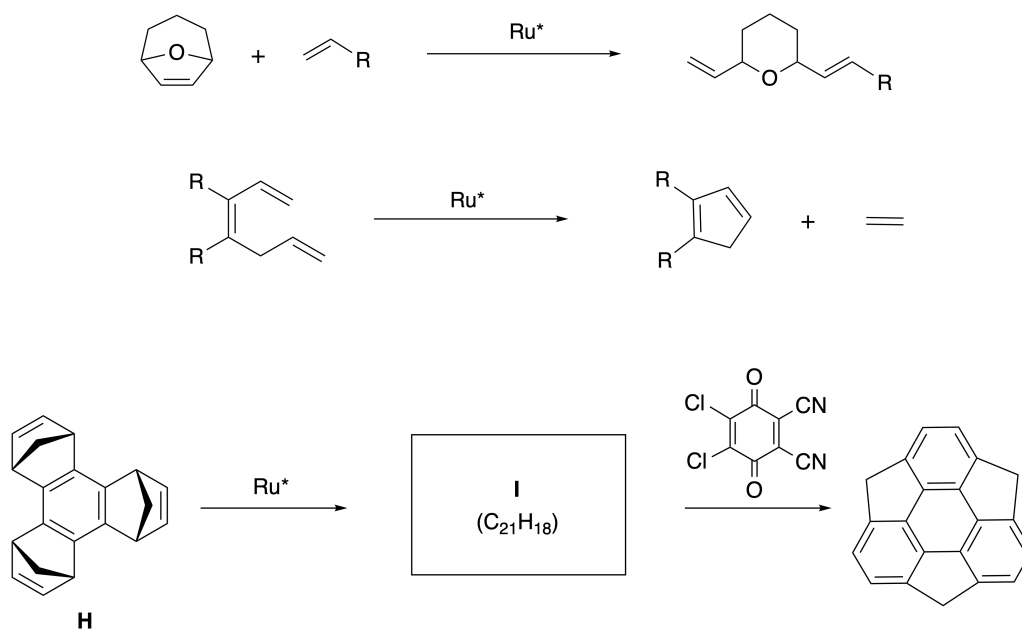


## Part B

Sumanene is a bowl-shaped hydrocarbon that was first reported in Japan in 2003. The name "sumanene" derives from a Sanskrit-Hindi word "suman" that means sunflower. The synthesis of sumanene was achieved by a reaction sequence that consists of a ring-opening and a ring-closing metathesis.

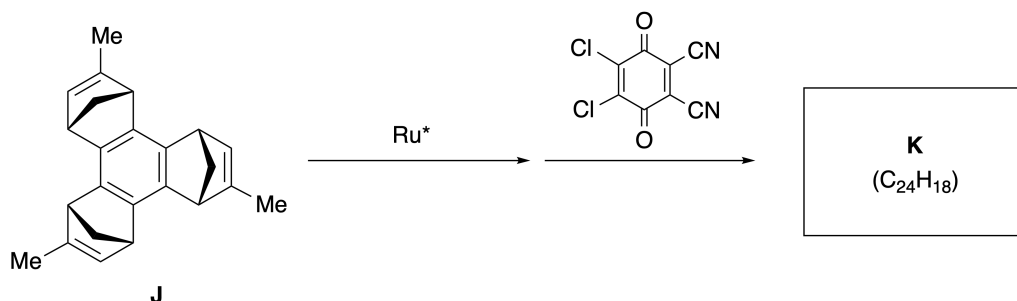


Representative metathesis reactions catalysed by a ruthenium catalyst ( $\text{Ru}^*$ ) are shown below.



**B.1** Draw the structure of intermediate **I** (its stereochemistry is not required).

3pt



- B.2** Starting from the optically active precursor **J**, the same reaction sequence gives the optically active sumanene derivative **K**. The stereocentres in **J** suffer no inversion during the metathesis reaction. **Draw** the structure of **K** with the appropriate stereochemistry. 4pt



GBR-1 C-8 A-1

**A8-1**  
British English (United Kingdom)

## Dynamic Organic Molecules and Their Chirality

### Part A

**A.1** (9 pt)

**A** (3 pt)

**B** (3 pt)

**C** (3 pt)

**A.2** (3 pt)



GBR-1 C-8 A-2

**A8-2**  
British English (United Kingdom)

**A.3** (7 pt)



GBR-1 C-8 A-3

# A8-3

British English (United Kingdom)

## Part B

**B.1** (3 pt)

**B.2** (4 pt)



GBR-1 C-9 C-1

**GBR-1 C-9 C**  
Alexander Thow

**ICHO**  
**Problem 9**  
**Cover sheet**

Please return this cover sheet together with all the related question sheets.

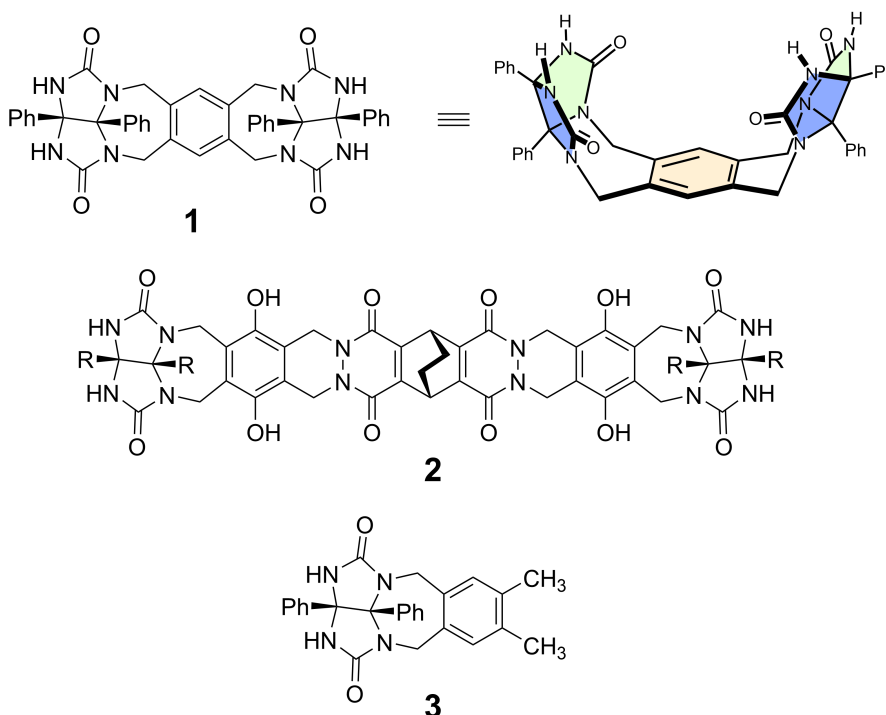
## Likes and Dislikes of Capsule

10 % of the total						
Question	A.1	A.2	A.3	A.4	A.5	Total
Points	13	2	2	3	3	<b>23</b>
Score						

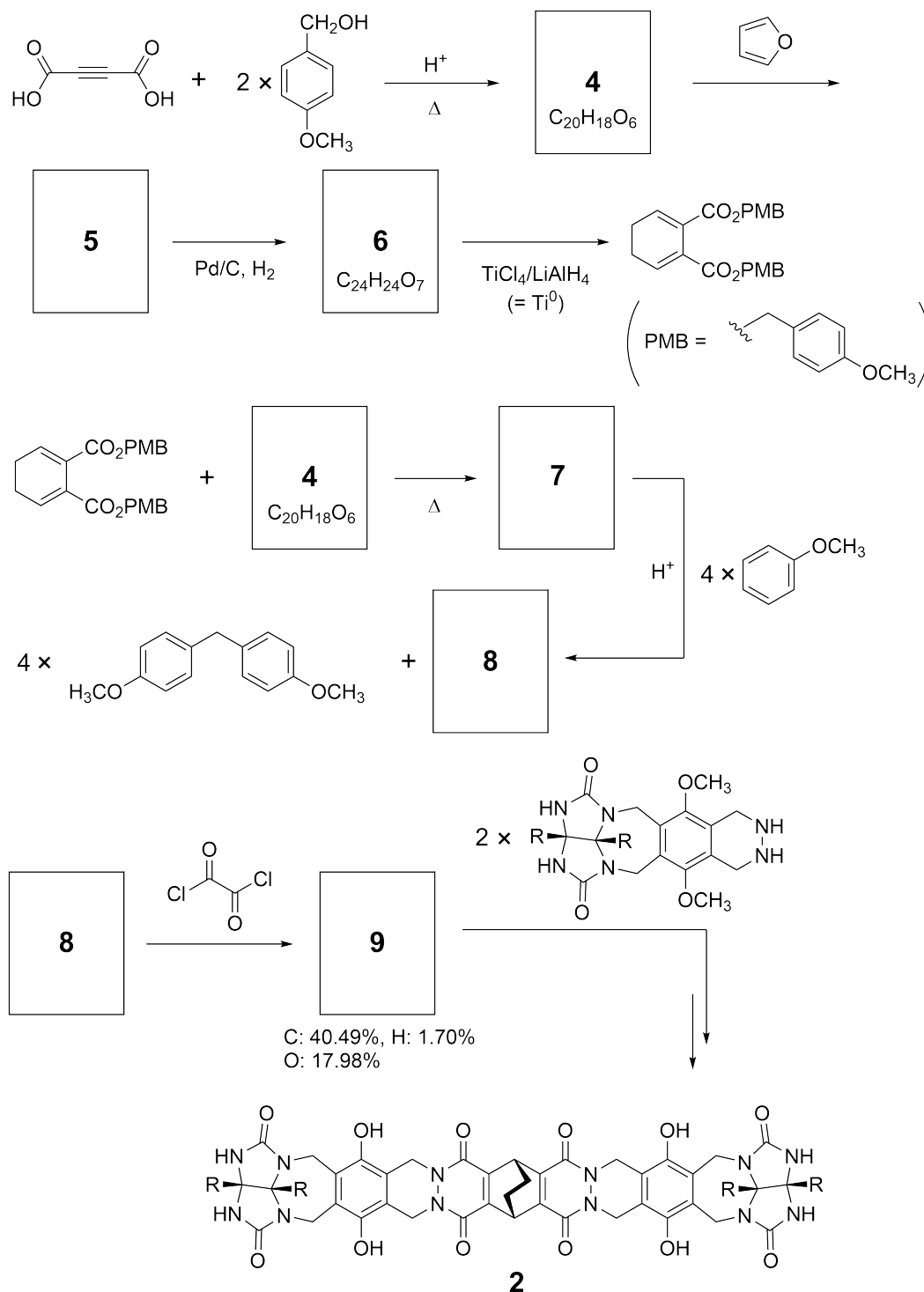
If you cut a tennis ball along the seam, you can disassemble it into two U-shaped pieces.



Compounds **1** and **2** are U-shaped molecules with different sizes, inspired by this idea. Compound **3** was prepared for comparison with **1** and the encapsulation behaviour of these compounds was investigated.



The synthetic route to **2** is shown below. Elemental composition of compound **9**: C 40.49%, H 1.70%, and O 17.98% by mass.





GBR-1 C-9 Q-3

# Q9-3

British English (United Kingdom)

**A.1** **Draw** the structures of **4-9**; stereochemistry can be neglected. Use the "PMB" abbreviation instead of drawing the whole structure of *p*-methoxybenzyl group shown in the scheme above. 13pt

In the mass spectrum of **1**, the ion peak corresponding to its dimer ( $1_2$ ) was clearly observed, whereas an ion peak for  $3_2$  was not observed in the spectrum of **3**.

In the  $^1\text{H}$  NMR spectrum of a solution of  $1_2$ , all the NH protons derived from **1** were observed to be chemically equivalent, and their chemical shift was significantly different from that of the NH protons of **3**. These data indicate that hydrogen bonds are formed between the NH moieties of **1** and atoms **X** of another molecule of **1** to form the dimeric capsule.

**A.2** **Circle** all the appropriate atom(s) **X** in **1**. 2pt

**A.3** **Give** the number of the hydrogen bonds in the dimeric capsule ( $1_2$ ). 2pt

The dimeric capsule of **1** ( $1_2$ ) has an internal space wherein an appropriate small molecule Z can be encapsulated. This phenomenon is expressed by the following equation:

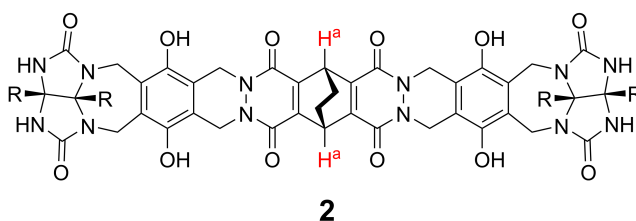


The equilibrium constant of the encapsulation of Z into  $1_2$  is shown below:

$$K_a = \frac{[Z@1_2]}{[Z][1_2]} \quad (2)$$

Encapsulation of a molecule into a capsule can be monitored by NMR spectroscopy. For example,  $1_2$  in  $C_6D_6$  gave different signals in the  $^1H$  NMR spectra before and after addition of  $CH_4$ .

Compound **2** also forms a larger, rigid dimeric capsule ( $2_2$ ). The  $^1H$  NMR spectrum of  $2_2$  was measured in  $C_6D_6$ ,  $C_6D_5F$ , and a  $C_6D_6/C_6D_5F$  solvent mixture, with all other conditions being kept constant. The chemical shifts for the  $H^a$  proton of **2** in the above solvents are summarized below, and no other signals from  $H^a$  in **2**, except for those listed, were observed. Assume that the interior of the capsule is always filled with the largest possible number of solvent molecules and that each signal corresponds to one species of filled capsule.



solvent	$\delta$ (ppm) of $H^a$
$C_6D_6$	4.60
$C_6D_5F$	4.71
$C_6D_6 / C_6D_5F$	4.60, 4.71, 4.82

**A.4 Determine** the number of  $C_6D_6$  and  $C_6D_5F$  molecules encapsulated in  $2_2$  giving 3pt each  $H^a$  signal.

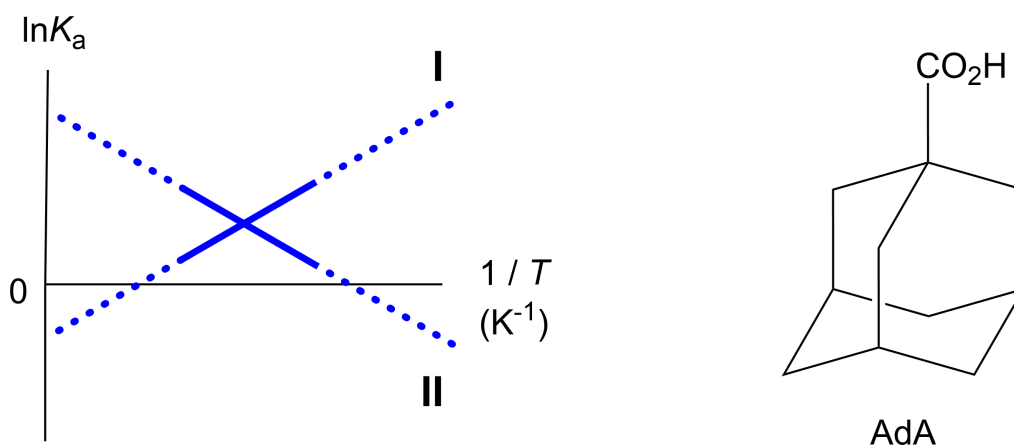
$^1\text{H}$  NMR measurements in  $\text{C}_6\text{D}_6$  revealed that  $2_2$  can incorporate one molecule of 1-adamantanecarboxylic acid (AdA).

The association constants ( $K_a$ ) which are expressed below were determined for various temperatures.

$[\text{solvent}@2_2]$  denotes a species containing one or more solvent molecules.

$$K_a = \frac{[\text{Z}@2_2]}{[\text{Z}][\text{solvent}@2_2]} \quad (3)$$

Similarly, the  $K_a$  values of  $\text{CH}_4$  and  $1_2$  shown in eq (2) were also determined by  $^1\text{H}$  NMR measurements at various temperatures in  $\text{C}_6\text{D}_6$ . The plots of the two association constants (as  $\ln K_a$  vs  $1/T$ ) are shown below.



No  $\text{C}_6\text{D}_6$  is encapsulated in  $1_2$ .

**A.5** **Choose** the correct options in gaps (1)–(5) in the following paragraph from A 3pt and B.

In line **II**, the entropy change ( $\Delta S$ ) is ( 1 ) and enthalpy change ( $\Delta H$ ) is ( 2 ), indicating that the driving force for the encapsulation in line **II** is ( 3 ). Therefore, line **I** corresponds to ( 4 ), and line **II** corresponds to ( 5 ).

	A	B
(1)	positive	negative
(2)	positive	negative
(3)	$\Delta S$	$\Delta H$
(4)	$1_2$ and $\text{CH}_4$	$2_2$ and AdA
(5)	$1_2$ and $\text{CH}_4$	$2_2$ and AdA



GBR-1 C-9 A-1

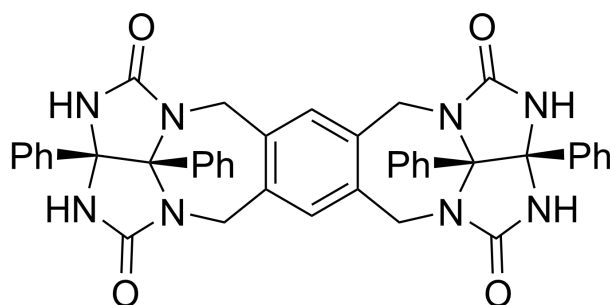
**A9-1**  
British English (United Kingdom)

## Likes and Dislikes of Capsule

**A.1** (13 pt)

<b>4</b> (2 pt)	<b>5</b> (3 pt)
<b>6</b> (2 pt)	<b>7</b> (2 pt)
<b>8</b> (2 pt)	<b>9</b> (2 pt)

**A.2** (2 pt)



**A.3** (2 pt)

**A.4** (3 pt)

$\delta$ (ppm) of H <sup>a</sup>	number of C <sub>6</sub> D <sub>6</sub>	number of C <sub>6</sub> D <sub>5</sub> F
4.60 ppm		
4.71 ppm		
4.82 ppm		

**A.5** (3 pt)

(1) : \_\_\_\_\_ (2) : \_\_\_\_\_ (3) : \_\_\_\_\_

(4) : \_\_\_\_\_ (5) : \_\_\_\_\_